

Halide-Assisted Macrocyclic Ring Formation in Cyclometalated Carbosilane Dendrimers with 1-[C₆H₃(CH₂NMe₂)₄-(PdCl)-3] Peripheral Groups: Application as Aldol Condensation Catalysts

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A series of cyclopalladated carbosilane dendrimers functionalized with monoanionic [C₆H₄(CH₂NMe₂)₄][−] (= CN) C,N-chelating ligands via a multiple and selective C–H bond activation process have been prepared. The structure of the support, i.e., the branching degree as well as the nature of the alkanediyl linkers present in the carbosilane dendrimer, was varied to investigate its influence on the dimer formation of the peripheral [CN–PdCl] complexes and on the catalytic performance of these Pd(II)-functionalized dendrimers. Me₂-Si{(CH₂)₃SiMe₂–CN–PdCl}₂ (**4**) and Si{(CH₂)₂SiMe₂–CN–PdCl}₄ (**7a**) were characterized by a single X-ray crystal structure determination. In both cases a unique intramolecular halide-assisted macrocyclic ring structure was present. Variable-temperature NMR spectroscopic investigations indicated that in solutions of **4** and **7a** the *cis*-bridging mode of the peripheral dimers is predominant. For the two asymmetrically branched [G1] palladodendrimers Si{(CH₂)₂Si(Me){(CH₂)₂SiMe₂–CN–PdCl}₂}₄ (**10a**) and Si{(CH₂)₃Si(Me){(CH₂)₃SiMe₂–CN–PdCl}₂}₄ (**12a**) isomeric forms of the [CN–PdCl] dimeric units were observed. In addition, for the [G2] dendrimer Si{(CH₂)₃Si{(CH₂)₃Si{(CH₂)₃SiMe₂–CN–PdCl(pyr)}₃}₃}₄, **15b**, appended with a maximum of 36 [CN–PdCl(pyr)] complexes, a *cis/trans* isomerization of the peripheral [CN–PdCl(pyr)] groups is noted. The applicability of the cationic cyclopalladated carbosilane dendrimers, obtained by halide abstraction with AgBF₄, as Lewis acid catalysts in a model aldol condensation reaction was demonstrated. Their catalytic performance was compared to that of the parent mononuclear derivative [Pd(C₆H₃{CH₂NMe₂}-2-(SiMe₃)-5)(pyr)(H₂O)]BF₄ (**16**). The obtained catalytic data showed that up to generation [G2], the selectivity of the reaction is not affected. However, the rate of the reaction decreased with increasing steric congestion at the dendrimer periphery.

Introduction

The chemistry of C,N-chelated cyclometalated compounds has received much attention,¹ and recently the preparation of multi-cyclometalated systems has been reported.² These include systems such as (chiral) ferrocenylamine,³ ferrocenylimine,⁴ and primary and/or tertiary benzylamine ligands.⁵ We are interested in the design of macromolecular (dendritic)⁶ organometallics

for application in catalysis⁷ and sensor-type materials.⁸ The application of dendritic metalodendrimers in membrane reactors has recently been attempted, and the feasibility of recycling these dendrimer catalysts by ultrafiltration techniques was shown.⁹ We have also reported the selective conversion of dendritic carbosilanes functionalized with the monoanionic arylamine

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(1) For reviews on this subject see: (a) Ryabov, A. D. *Chem. Rev.* **1990**, *90*, 403. (b) Pfeffer, M. *Pure Appl. Chem.* **1992**, *64*, 335–342. (c) Steenwinkel, P.; Gossage, R. A.; van Koten, G. *Chem. Eur. J.* **1998**, *4*, 159.

(2) See for instance: (a) Burckhardt, U.; Baumann, M.; Trabesinger, G.; Gramlich, V.; Togni, A. *Organometallics* **1997**, *16*, 5252–5259. (b) Fuchita, Y.; Yoshinaga, K.; Hanaki, T.; Kawano, H.; Kinoshita-Nagaoka, J. *J. Organomet. Chem.* **1999**, *580*, 273–281.

(3) For recent examples see: (a) Ranatunge-Bandarage, P. R. R.; Robinson, B. H.; Simpson, J. *Organometallics* **1994**, *13*, 500–510. (b) Ranatunge-Bandarage, P. R. R.; Duffy, N. W.; Johnston, S. M.; Robinson, B. H.; Simpson, J. *Organometallics* **1994**, *13*, 511–521.

(4) For recent examples see: (a) López, C.; Bosque, R. *J. Organomet. Chem.* **1996**, *524*, 247–252. (b) Zhao, G.; Xue, F.; Zhang, Z. Y.; Mak, T. C. W. *Organometallics* **1997**, *16*, 4023–4026. (c) Bosque, R.; López, C.; Solans, X.; Font-Bardía, M. *Organometallics* **1999**, *18*, 1267–1274.

(5) (a) Cope, A. C.; Friedrich, E. C. *J. Am. Chem. Soc.* **1968**, *90*, 909–913. (b) Vicente, J.; Saura-Llamas, I.; Jones, P. G. *J. Chem. Soc., Dalton Trans.* **1993**, 3619–3624. (c) Albert, J.; Granell, J.; Minguez, J.; Muller, G.; Sainz, D.; Valerga, P. *Organometallics* **1997**, *16*, 3561–3564.

(6) For the most recent reviews on dendritic molecules see: (a) Bosman, A. W.; Janssen, H. M.; Meijer, E. W. *Chem. Rev.* **1999**, *99*, 1665. (b) Newkome, G.; He, E.; Moorefield, C. *Chem. Rev.* **1999**, *99*, 1689. (c) Majoral, J. P.; Caminade, A. M. *Chem. Rev.* **1999**, *99*, 845–880.

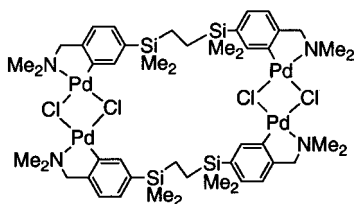


Figure 1. Schematic representation of a bis-palladated, dimeric carbosilane derivative.

ligands $[\text{C}_6\text{H}_3(\text{CH}_2\text{NMe}_2)_2\text{-}3,5]^-$ (NCN) and $[\text{C}_6\text{H}_4(\text{CH}_2\text{NMe}_2)_4]^-$ (CN) into synthetically useful, polyolithiated ligand arrays¹⁰ and their subsequent use for the preparation of a variety of group 8 metallodendritic compounds.^{7d,11} As these polyolithium reagents are sensitive toward hydrolysis, incomplete transmetalation is difficult to avoid. This prompted us to investigate the possibility of alternative and more direct methods for the introduction of metal sites into arylamine-functionalized carbosilane supports. In principle, this could overcome the problem of incomplete metalation.

Our earlier work had led to the preparation of a stable, macrocyclic cyclopalladated carbosilane which was structurally characterized by X-ray crystallography (Figure 1).¹¹ This palladated carbosilane exists as a dimeric aggregate in the solid state with a 26-membered-ring structure. Noteworthy is the fact that in this particular palladation reaction, unlike in previous reported work,^{1c,12} the connected carbosilane backbone was not attacked by the metal reagent, showing that regioselective cyclopalladation of this type of dendritic carbosilanes with peripheral CN ligands is feasible.

Here, we present a general strategy for the incorporation of Pd(II) metal sites into the periphery of arylamine-functionalized carbosilane dendrimers by a multiple C–H bond activation. As an extension of our earlier efforts,¹¹ we have also studied the halide-assisted dimer formation at the periphery of these novel metallodendritic species by variation of the branching degree as well as the alkanediyl linking fragments present in the carbosilane dendrimer support (Figure 2). Study of these

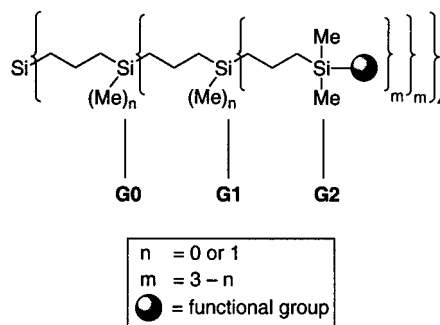


Figure 2. Schematic representation of the variation of the structure of the dendritic carbosilane support.

latter aspects is relevant because of our recent finding, which pointed to the effect of the number of organometallic catalytic sites at the periphery on the catalyst performance.^{7d,e} We have employed the cyclopalladated carbosilane dendrimers as catalyst precursors in the aldol condensation reaction. The obtained catalytic data are compared with the parent model species and discussed.

Results and Discussion

Synthesis of Cyclopalladated Compounds. The model dendritic compounds (viz., **4** and **5a**) were prepared to study the cyclopalladation and halide-assisted dimer formation in more detail. Treatment of dialyldimethylsilane with an excess of chlorodimethylsilane in the presence of a platinum catalyst (i.e., $(\text{NBu}_4)_2\text{-PtCl}_6$) gave the hydrosilylated product **1** in quantitative yield (Scheme 1). Compound **1**, which is moisture sensitive, was converted to the moisture stable silicon hydride **2** in 74% yield by reduction of **1** in the presence of an excess of LiAlH_4 in Et_2O at 0 °C. Not only did the selective formation of **2** confirm the stoichiometry of its precursor **1**, but at the same time **2** itself is an interesting starting material for the synthesis of hyperbranched carbosilane polymers.¹³ The CN-functionalized carbosilane $[\text{G}0]^{1/2}\text{-(SiMe}_2\text{-CN)}_2$ (**3**) was obtained in 76% yield by treatment of the silicon chloride **1** with an excess of $[\text{Li}(\text{C}_6\text{H}_4\{\text{CH}_2\text{NMe}_2\}_4)]$ (abbreviated as Li-[CN]) in Et_2O at –78 °C.¹⁰ Cyclopalladation of **3** in the presence of $\text{Pd}(\text{OAc})_2$ in MeOH at ambient temperature yielded the bis palladium chloride derivative $[\text{G}0]^{1/2}\text{-(SiMe}_2\text{-CN-PdCl)}_2$ (**4**) as a yellow-to-orange solid in 83% yield after addition of LiCl (Scheme 1). Crystals suitable for X-ray diffraction (vide infra) were obtained by layering a solution of **4** in benzene with pentane.

The $[\text{G}0]$ dendrimers **5a** and **7a**, which differ with respect to the length of the alkanediyl spacer (viz., C_2 vs C_3), were prepared as reported for **4** (Scheme 2). Thus, addition of $\text{Pd}(\text{OAc})_2$ to a stirred solution of $[\text{G}0]\text{-(SiMe}_2\text{-CN)}_4$ in MeOH followed by an appropriate workup yielded orange-colored $[\text{G}0]\text{-(SiMe}_2\text{-CN-PdCl)}_4$ (**5a**) in 51% yield (Scheme 2). The dendritic palladium complex **7a** was obtained by a two-step sequence. First, known $\text{Si}(\text{CH}_2\text{CH}_2\text{SiMe}_2\text{Cl})_4$ ¹⁴ (abbreviated as $[\text{G}0]^\#$ - $(\text{SiMe}_2\text{Cl})_4$, the symbol # denotes the presence of C_2 instead of C_3 branches/spacers in this compound) was treated with Li-[CN] (as described for **3**) to give $[\text{G}0]^\#$ -

(7) (a) Knapen, J. W. J.; van der Made, A. W.; De Wilde, J. C.; van Leeuwen, P. W. M. N.; Wijkens, P.; Grove, D. M.; van Koten, G. *Nature* **1994**, *372*, 659. (b) Dijkstra, H. P.; Steenwinkel, P.; Grove, D. M.; Lutz, M.; Spek, A. L.; van Koten, G. *Angew. Chem., Int. Ed.* **1999**, *38*, 2186. (c) Gossage, R. A.; Jastrzebski, J. T. B. H.; van Ameijde, J.; Mulders, S. J. E.; Brouwer, A. J.; Liskamp, R. M. J.; van Koten, G. *Tetrahedron Lett.* **1999**, *40*, 1461–1464. (d) Kleij, A. W.; Gossage, R. A.; Jastrzebski, J. T. B. H.; Boersma, J.; van Koten, G. *Angew. Chem., Int. Ed.* **2000**, *39*, 176. (e) Kleij, A. W.; Gossage, R. A.; Klein Gebbink, R. J. M.; Brinkmann, N.; Kragl, U.; Reyerse, E. J.; Lutz, M.; Spek, A. L.; van Koten, G. *J. Am. Chem. Soc.* **2000**, *122*, 12112–12124.

(8) Albrecht, M.; Gossage, R. A.; Lutz, M.; Spek, A. L.; van Koten, G. *Chem. Eur. J.* **2000**, *6*, 1431.

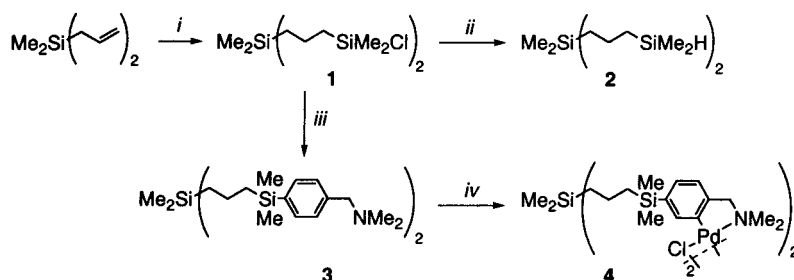
(9) (a) Hovestad, N. J.; Eggeling, E. B.; Heidebüchel, H. J.; Jastrzebski, J. T. B. H.; Kragl, U.; Keim, W.; Vogt, D.; van Koten, G. *Angew. Chem.* **1999**, *111*, 1763; *Angew. Chem., Int. Ed.* **1999**, *38*, 1655. (b) Brinkmann, N.; Giebel, D.; Lohmer, G.; Reetz, M. T.; Kragl, U. *J. Catal.* **1999**, *183*, 163–168. (c) de Groot, D.; Eggeling, E. B.; de Wilde, J. C.; Kooijman H.; van Haaren, R. J.; van der Made, A. W.; Spek, A. L.; Vogt, D.; Reek, J. N. H.; Kamer, P. J. C.; van Leeuwen, P. W. N. M. *Chem. Commun.* **1999**, 1623–1624.

(10) Kleij, A. W.; Kleijn, H.; Jastrzebski, J. T. B. H.; Smeets, W. J. J.; Spek, A. L.; van Koten, G. *Organometallics* **1999**, *18*, 268–276.

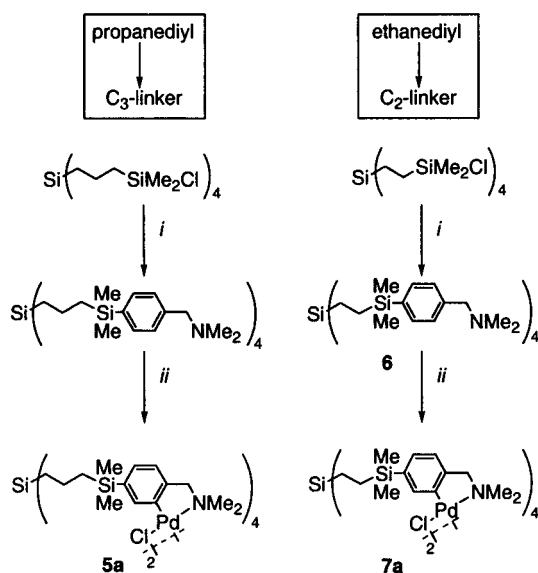
(11) Kleij, A. W.; Kleijn, H.; Jastrzebski, J. T. B. H.; Spek, A. L.; van Koten, G. *Organometallics* **1999**, *18*, 277–285. The assignment of *cis* and *trans* isomers is arbitrary but follows literature reports. See, for example, ref 18.

(12) (a) Steenwinkel, P.; James, S. L.; Grove, D. M.; Kooijman, H.; Spek, A. L.; van Koten, G. *Organometallics* **1997**, *16*, 513. (b) Dani, P.; Karlen, T.; Gossage, R. A.; Smeets, W. J. J.; Spek, A. L.; van Koten, G. *J. Am. Chem. Soc.* **1997**, *119*, 11317–11318.

(13) See for instance: Lach, C.; Müller, P.; Frey, H.; Mülhaupt, R. *Macromol. Rapid Commun.* **1997**, *18*, 253.

Scheme 1. Synthesis of the Model Derivative **4**^a

^a Reagents and conditions: (i) neat HSiMe₂Cl, RT, (NBu₄)₂PtCl₆ as catalyst; (ii) excess LiAlH₄, Et₂O, 0 °C; (iii) Li-[CN], Et₂O, -78 °C → RT; (iv) Pd(OAc)₂, MeOH; then LiCl, MeOH.

Scheme 2. Synthesis of the [G0] Dendrimers **5a** and **7a**^a

^a Reagents and conditions: (i) Li-[CN], Et₂O, -78 °C → RT; (ii) Pd(OAc)₂, MeOH; then LiCl, MeOH.

(SiMe₂-CN)₄, **6**, as a waxy white solid in 72% yield. Subsequent palladium incorporation as reported for **4** and **5** (Scheme 2) afforded [G0][#]-(SiMe₂-CN-PdCl)₄, **7**, in good yield (91%). Unlike their [G1] generation analogues (vide infra), the [G0] generation dendrimers **5a** and **7a** as well as “dendritic” **4** are soluble in aromatic solvents such as benzene and toluene.

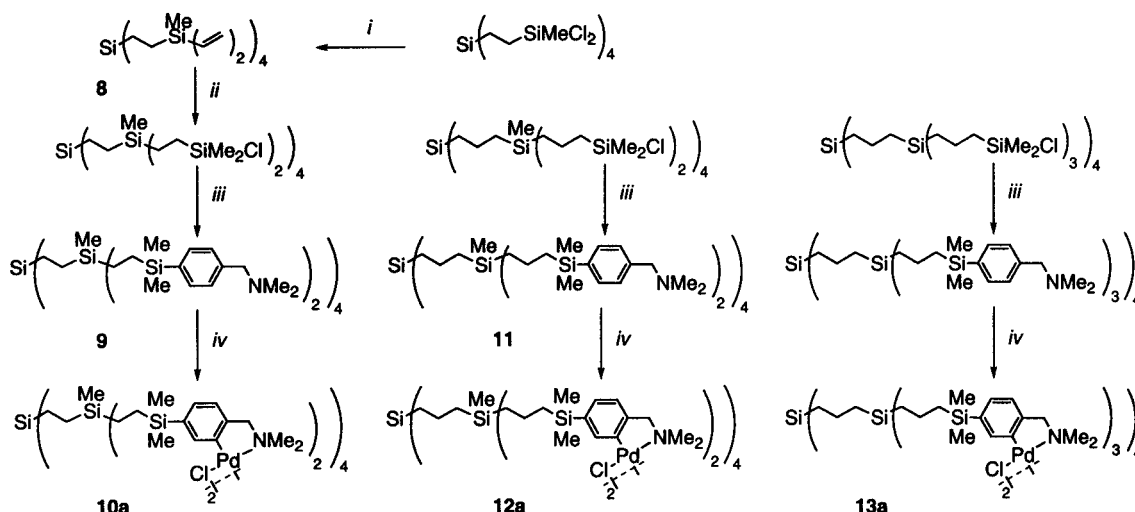
The synthesis of the [G1] generation dendrimers is outlined in Scheme 3. The derivative [G1][#]-(CH=CH₂)₈ (**8**) was obtained in 63% yield by the vinylation of the known silicon chloride [G0][#]-(SiMeCl₂)₄.¹⁴ The olefinic derivative **8** was converted into the chlorosilane dendrimer [G1][#]-(SiMe₂Cl)₈¹⁴ by a hydrosilylation reaction with HSiMe₂Cl using a platinum-based catalyst. The dendrimer ligand [G1][#]-(SiMe₂-CN)₈ (**9**) was obtained in 84% yield by treatment of [G1][#]-(SiMe₂Cl)₈ with an excess of the lithium reagent Li-[CN]. Likewise, [G1]-(SiMe₂-CN)₈ (**11**) was obtained as a yellow viscous oil in 78% yield by using an approach similar to that described for **9** (Scheme 3).¹⁰ The [G1] generation palladodendrimers were prepared as described for the

[G0] generation dendrimers **5a** and **7a** by a multiple C-H activation using Pd(OAc)₂ in MeOH. The [G1] derivatives [G1][#]-(SiMe₂-CN-PdCl)₈ (**10a**), [G1]-(SiMe₂-CN-PdCl)₈ (**12a**), and [G1]-(SiMe₂-CN-PdCl)₁₂ (**13a**) were isolated as pale yellow-to-orange solids in 40–80% yield (Scheme 3).

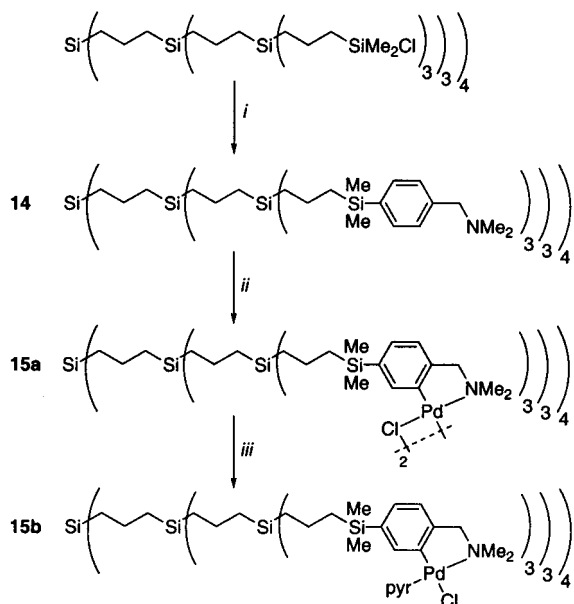
The CN-functionalized dendrimer [G2]-(SiMe₂-CN)₃₆ (**14**) was obtained in a moderate yield (38%) by a procedure similar to the one used for the preparation of **9** and **11** starting from [G2]-(SiMe₂Cl)₃₆.¹⁴ Preliminary investigations showed that the incorporation of Pd(II) sites on the periphery of [G2]-(SiMe₂-CN)₃₆ (**14**) by a multiple C-H activation using Pd(OAc)₂ as described for the [G0] and [G1] analogues afforded a yellow-to-brown, insoluble solid. The insolubility of the presumed product (i.e., [G2]-(SiMe₂-CN-PdCl)₃₆ (**15a**), Scheme 4) could be explained by a large increase in aggregational character in this intermediate species. The presumed product [G2]-(SiMe₂-CN-PdCl)₃₆ (**15a**) was converted into the corresponding pyridine adduct [G2]-(SiMe₂-CN-PdCl{pyr})₃₆ (**15b**) by an in situ preparation in CH₂Cl₂. Indeed, **15b** appeared to be more soluble and was isolated as a light yellow solid in 38% yield after appropriate workup (Scheme 4). The ¹H NMR spectrum of **15b** in CDCl₃ unambiguously showed that all the peripheral CN ligands had been cyclometalated, as indicated by the complete disappearance of the resonances corresponding to the CH₂N and NMe₂ fragments of the parent CN ligand at 3.40 and 2.22 ppm, respectively. Interestingly, two new resonance patterns were observed for several fragments of the [G2] pyridine adduct, of which the relative intensity ratio was temperature dependent (vide infra). The proposed stoichiometry of [G2]-(SiMe₂-CN-PdCl{pyr})₃₆ (**15b**) was further supported by elemental (including Pd and Cl) analyses.

Addition of an excess of pyridine to the dendritic palladium chloride compounds [G0]-(SiMe₂-CN-PdCl)₄ (**5a**), [G0][#]-(SiMe₂-CN-PdCl)₄ (**7a**), [G1][#]-(SiMe₂-CN-PdCl)₈ (**10a**), [G1]-(SiMe₂-CN-PdCl)₈ (**12a**), and [G1]-(SiMe₂-CN-PdCl)₁₂ (**13a**) in CH₂Cl₂ led to a rapid formation of their pyridine adducts [G0]-(SiMe₂-CN-PdCl{pyr})₄ (**5b**), [G0][#]-(SiMe₂-CN-PdCl{pyr})₄ (**7b**), [G1][#]-(SiMe₂-CN-PdCl{pyr})₈ (**10b**), [G1]-(SiMe₂-CN-PdCl{pyr})₈ (**12b**), and [G1]-(SiMe₂-CN-PdCl{pyr})₁₂ (**13b**), respectively. These pyridine adducts (Scheme 5) were isolated as off-white solids in excellent yields (84% to quantitative). The selective formation of the pyridine adducts became directly evident from the ¹H NMR spectra, in which the Ar-H_{ortho} of the peripheral CN ligands shifts to higher field upon coordination of the

(14) (a) Seyferth, D.; Son, D. Y.; Rheingold, A. L.; Ostrander, R. L. *Organometallics* **1994**, *13*, 2682. (b) van der Made, A. W.; van Leeuwen, P. W. M. N.; De Wilde, J. C.; Brandes, A. C. *Adv. Mater.* **1993**, *5*, 466. (c) van der Made, A. W.; van Leeuwen, P. W. M. N. *J. Chem. Soc., Chem. Commun.* **1992**, 1400. (d) Zhou, L.-L.; Roovers, J. *Macromolecules* **1993**, *26*, 963.

Scheme 3. Synthesis of the [G1] Dendrimers **10**, **12a**, and **13a**^a

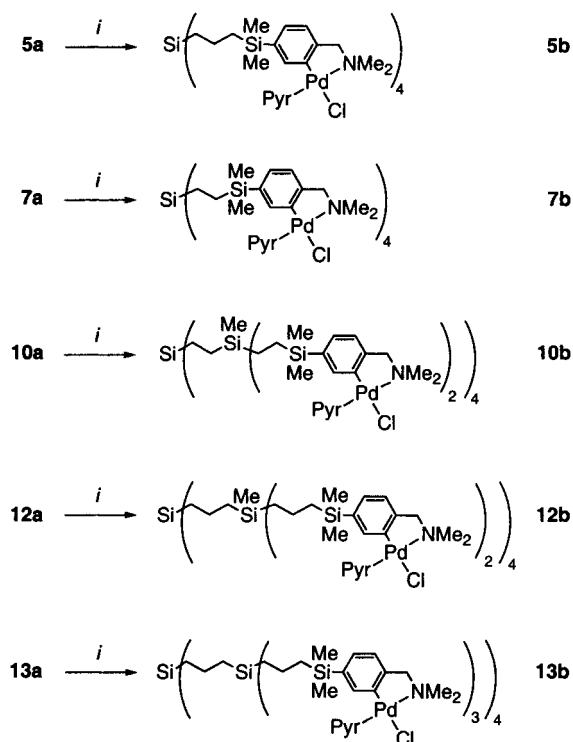
^aReagents and conditions: (i) excess VinylMgBr, THF, RT; (ii) neat HSiMe₂Cl, RT, (NBu₄)₂PtCl₆ as catalyst; (iii) Li-[CN], Et₂O, -78 °C → RT; (iv) Pd(OAc)₂, MeOH; then LiCl, MeOH.

Scheme 4. Synthesis of the [G2] Dendrimer **15b**^a

^aReagents and conditions: (i) Li-[CN], Et₂O, -78 °C → RT; (ii) Pd(OAc)₂, MeOH; then LiCl, MeOH; (iii) pyridine, CH₂Cl₂.

pyridine ligand to the metal center (i.e., from ca. 7.3–7.4 ppm to 6.00–6.05 ppm). In addition, for the connecting SiMe₂ groups, a significant ¹H NMR shift from ca. 0.2 to -0.04 ppm was observed, pointing to a proximity effect between the coordinated pyridine ligand and the (*meta*) SiMe₂ groups. Moreover, the line widths of the resonances sharply decreased in the case of the pyridine adducts when compared to their chloride-bridged pallado-dendrimers. Finally, the combined analytical data of the pyridine adducts supported the selective and *quantitative* (cf., the nonquantitative lithiation/transmetalation protocol^{7(d,e)}) palladation of the dendrimer ligands presented in Schemes 3 and 4. The pyridine adducts were used as starting materials for the preparation of homogeneous Lewis acid catalysts for the aldol condensation reaction (vide infra).

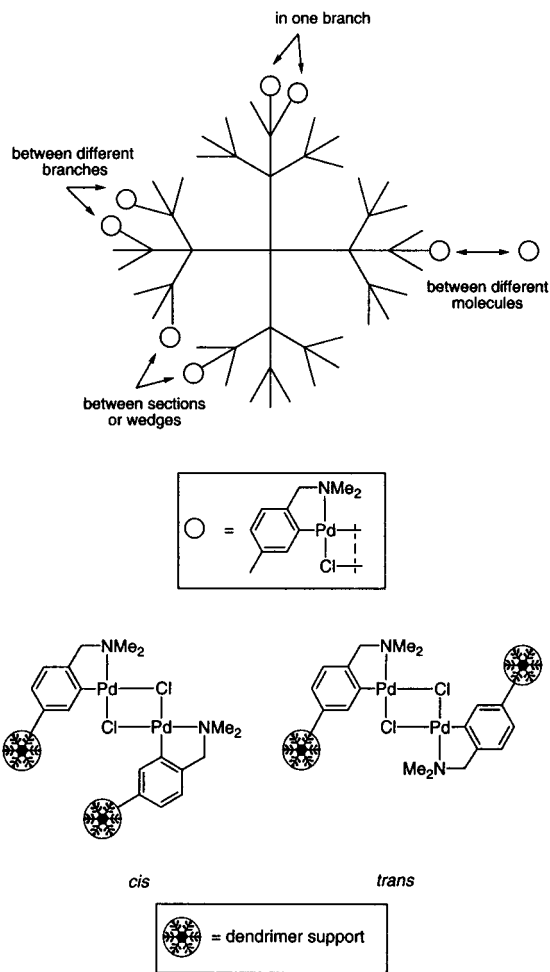
NMR Spectroscopic Analysis of the Metallodendrimer Species. In contrast with the sharp resonances

Scheme 5. Synthesis of the Dendrimer (Poly)pyridine Adducts^a

^aReagents and conditions: (i) pyridine, CH₂Cl₂.

observed in the ¹H NMR spectra of nearly all dendrimer ligands, spectra with relatively broad resonances were collected for the palladodendrimers [G0]-(SiMe₂-CN-PdCl)₄ (**5a**), [G0][#]-(SiMe₂-CN-PdCl) (**7a**), [G1][#]-(SiMe₂-CN-PdCl)₈ (**10a**), [G1]-(SiMe₂-CN-PdCl)₈ (**12a**), and [G1]-(SiMe₂-CN-PdCl)₁₂ (**13a**). Most probably this broadening arises from the halide-assisted dimer formation of the attached Pd(II)Cl complexes, which was clearly lost when a suitable Lewis base (i.e., pyridine) was added, resulting in spectra with relatively sharp resonances. It should be mentioned that this halide-assisted dimer formation leads in the case of the [G1] and [G2] species **10a**, **12a**, **13a**, and **15a** to the formation of several stereoisomers as schematically shown in

Scheme 6. Schematic Representation of the Stereoisomer Formation in the Cyclopalladated CN-Derivatized Carbosilane Dendrimers



Scheme 6. These stereoisomers can originate from dimer formation between [CN–PdCl] complexes of different branches, sections, or wedges. As illustrated in Scheme 6, the formation of *intermolecular*, large aggregates is not ruled out and the relatively low isolated yields of the higher generation species **13a** and **15a** may serve as indicative examples thereof. Due to the applied workup procedure(s), the insoluble, presumably polymeric fractions of these chloride-bridged Pd(II) dendrimers were not isolated. We have intentionally only focused on the soluble, isolated materials that can be used in homogeneous catalytic applications and have undertaken no studies toward elucidation of the structures of the insoluble materials.

The occurrence of *cis* and *trans* isomeric forms¹⁶ of the dimer has been well recognized for, for example, the parent species (μ -Cl-[PdCl(C₆H₃{CH₂NMe₂}-2-SiMe₃-5))₂ by means of ¹H and ¹³C{¹H} NMR spectroscopy (isomer ratio \approx 1.4, δ (¹H, CDCl₃, Ar–H_{ortho}) = 7.43 (*cis*), 7.33 (*trans*)).¹¹ Molecular modeling of both isomeric

(15) The halide-assisted dimer formation is reversible due to the presence of a number of different nucleophiles (see for instance: Hockless, D. C. R.; Gugger, P. A.; Leung, P. H.; Mayadunne, R. C.; Pabel, M.; Wild, S. B. *Tetrahedron* **1997**, *53*, 4083). This would allow (quantitative) dimer formation throughout the periphery of these dendrimers.

(16) The prefixes *cis* and *trans* refer to the relative positions of the NMe₂ groupings with respect to the Pd₂Cl₂ “square” unit, cf., Scheme 6.

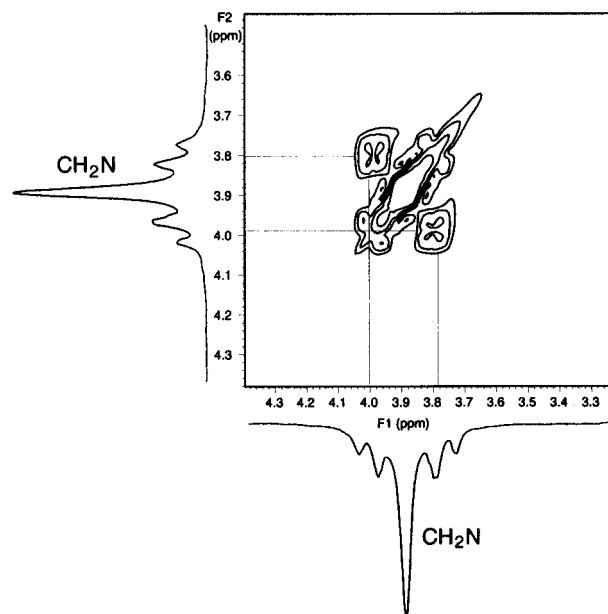


Figure 3. COSY (¹H, ¹H) spectrum recorded for **12a** in CDCl₃ at RT.

forms suggested a very small energy difference (i.e., 0.15 kcal/mol) between the *cis* and *trans* isomers of this particular compound in the gas phase.¹⁷ Recently, Gül et al. described analogous *cis/trans* isomerization processes involving similar dimeric Pd(II)X complexes (X = halide) derived from simple (chiral) arylamines. The ratio between the *cis* and *trans* isomeric forms (ca. 1:1, calculated from the ¹H NMR spectra) of these complexes was temperature independent between 20 °C and the coalescence temperature (50–80 °C). From these data a ΔG^\ddagger_{298} of 15–16 (± 0.5) kcal/mol was calculated for this isomerization process and the Gibbs activation energy was found to be practically solvent insensitive.¹⁸

The ¹H NMR spectra of the dendrimers [G1][#]-(SiMe₂-CN-PdCl)₈ (**10a**) and [G1]-(SiMe₂-CN-PdCl)₈ (**12a**), which have lower branching multiplicity as compared with **5a**, **7a**, and **13a**, displayed two resonance patterns for the CH₂N groups of the attached [CN–PdCl] complexes (Figures 3 and 4). One pattern is represented by a singlet resonance, while the other is an AB type pattern (²J(H_a–H_b) \approx 13 Hz) for these protons. The assignment of this AB pattern was supported by ¹H decoupling experiments giving rise to *singlet* resonances for both doublets after selective irradiation of the presumed coupled spins. In addition, the COSY (¹H, ¹H) NMR spectrum of **12a** (Figure 3) showed unambiguously that these doublets are correlated. Furthermore, the HETCOR (¹H, ¹³C) NMR spectrum of **12a** (Figure 4) clearly displayed a cross-peak between the CH₂ protons and the benzylic carbon atom.

Together, from these data we suggest the presence of more than one type of C,N-chelated Pd(II) complex for both **10a** and **12a** in solution, which is most probably due to different bridging modes between the peripheral Pd(II) complexes (Scheme 6). To our knowledge, these are the first characterized examples of isomeric forms

(17) CAChe Molecular Mechanics using the implemented MM2 force field parameters, Cambridge Molecular Group.

(18) Gül, N.; Nelson, J. H. *Organometallics* **2000**, *19*, 91.

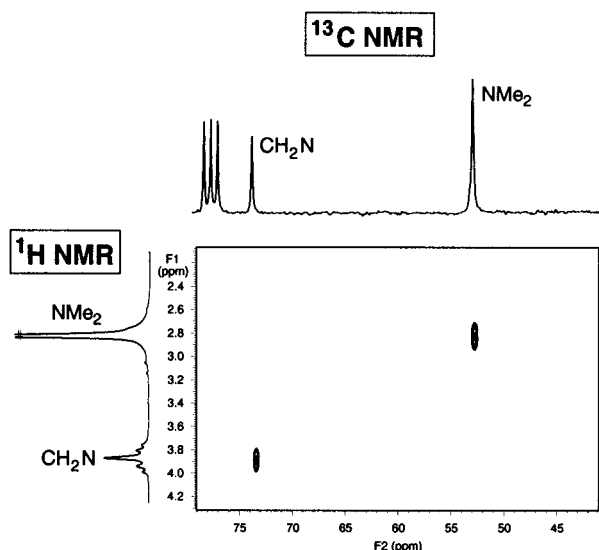


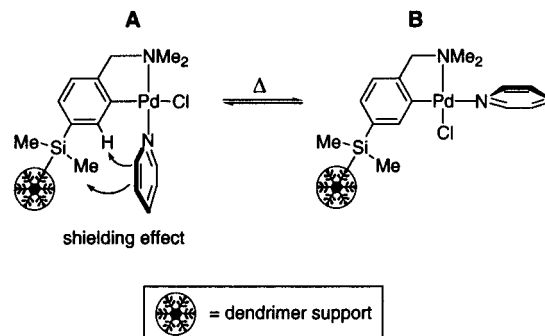
Figure 4. HETCOR (^1H , ^{13}C) spectrum recorded for **12a** in CDCl_3 at RT.

of dinuclear, d^8 -metal complexes at the periphery of a dendrimer. Moreover, it seems likely that the asymmetric branching of the dendrimer support as well as the fact that in **10a** and **12a** the branches and wedges have more mobility is of importance for these observations. It should be noted that the mutual points of attachment of the CN ligands in the chloride-bridged Pd(II) carbosilane dendrimers (see for examples the structures of **4** and **7a**, *vide infra*) are responsible for the lower degree of symmetry in these [G1] generation molecules.

We started to investigate the isomerism behavior of the dendritic species **4**, **7a**, and **12a** as the results of earlier studies^{11,18} suggested that *cis/trans* isomerism of these dimeric [CN–PdCl] complexes is a common feature. Therefore, variable-temperature (VT) NMR spectroscopy in CD_2Cl_2 was conducted for **12a**. In the temperature range -80 to $+40$ °C, no significant changes in relative intensities were observed for the AB pattern and the singlet resonance of the CH_2N groups. Interestingly, two distinct signals (partly overlapping) were noted for the Ar-H_{ortho} and SiMe_2 group of **12a** in CD_2Cl_2 in the same temperature range, whereas for the latter fragment, one coalesced signal was found in the temperature range 0 to $+40$ °C. It is important to emphasize that upon treatment of **12a** with pyridine, only *one* single resonance pattern was observed for all groups. This confirmed the complete conversion of the chloride-bridged, dinuclear into mononuclear, pyridine-ligated Pd(II) centers and unambiguously proved the proposed stoichiometry of **12a**.

To study the *temperature*-dependence of the isomer ratio of **4** and **7a**, VT ^1H NMR spectroscopy in toluene- d_8 was conducted. Interestingly, for both complexes only *one* (predominant) resonance pattern was observed in the temperature range -80 to $+100$ °C. For the CH_2N , $\text{N}(\text{CH}_3)_2$, and $\text{Si}(\text{CH}_3)_2$ groups, singlet resonances were found. From the VT ^1H NMR spectroscopic studies carried out for **4**, **7a**, and **12a**, we conclude that under these conditions the peripheral chloride-bridged Pd(II) dimers are stereochemically stable.¹⁵ In addition, the NMR data suggested that in **4** and **7a** there is a higher degree of symmetry as compared with **10a** and **12a**.

Scheme 7. Schematic Representation of the Proposed Binding Mode of the Pyridine Ligand in **15b**



For the [G2] dendrimer **15b**, two main, clearly different resonance patterns (partly overlapping) were observed for the SiMe_2 , CH_2N , and NMe_2 groups, whereas for the pyridine ligand three broad signals were noted. For the Ar-H_{ortho} a broad, singlet resonance is found at δ 6.00 ppm, which supports the presence of peripheral Pd(II) complexes which are ligated by pyridine ligands as described for the [G0] and [G1] analogues (Scheme 7, **A**). The two signals for the SiMe_2 groups were found at δ 0.16 and -0.07 ppm, respectively (cf., ca. 0.20 ppm for the SiMe_2 groups in the chloride-bridged [CN–PdCl] complexes in the [G0] and [G1] Pd(II) dendrimers). The ^1H NMR signal integration for the two peaks observed for the SiMe_2 groups corresponds well with that of the NMe_2 groups. The resonance at -0.07 ppm further supported the formation of the isomer **A**. We ascribe the second resonance (i.e., at 0.16 ppm) to the presence of CN-chelated Pd(II) complexes, in which the proximity between the pyridine ligand and the SiMe_2 dendrimer-to-ligand connecting groups in [G2]-(SiMe_2 -CN-PdCl{pyr})₃₆ (**15b**) is less as compared to isomer **A** (Scheme 7, **B**). VT ^1H NMR spectroscopy carried out for **15b** showed that at $+60$ °C the **A/B** ratio slightly changes in favor of **B** (**A/B** = 0.48). After cooling to ambient temperature, the original isomer ratio (**A/B** = 0.53) was again obtained. Similar ^1H NMR spectroscopic changes were noted for the NMe_2 and CH_2N groups.

Mass Spectrometric Analysis of the Dendrimer Species. Compounds **3**, **4**, **5a**, **6**, and **7a** (calcd molecular weight 526–1473) were analyzed using fast atom bombardment (FAB) mass spectrometry, while the metallodendrimers **8**, **9**, **10a**, **11**, **12a**, **13a**, and **14** (calcd molecular weight 636–9590) were studied by MALDI-TOF-MS. These results afforded unique additional proof for the monodisperse nature of the cyclometalated, dendritic species. For all cyclopalladated dendrimers, characteristic fragments ions $(M - \text{Cl})^+$ were unambiguously identified (e.g., Figure 5).

MALDI-TOF-MS studies performed for the pyridine adducts **7b**, **10b**, **12b**, and **13b** were less conclusive, as no molecular ions for this type of species could be generated. Most likely this was due to a fast *in situ* loss of the pyridine ligands under the experimental conditions. For example, the MALDI-TOF mass spectrum of **7b** displayed the most intensive peak at m/z 1439.4, which was attributed to the fragment ion $[\mathbf{7b} - \text{Cl} - 4\text{pyr}]^+$.

Solid State Structures of **4 and **7a**.** The influence of the presence of different linkers (i.e., propanediyl vs

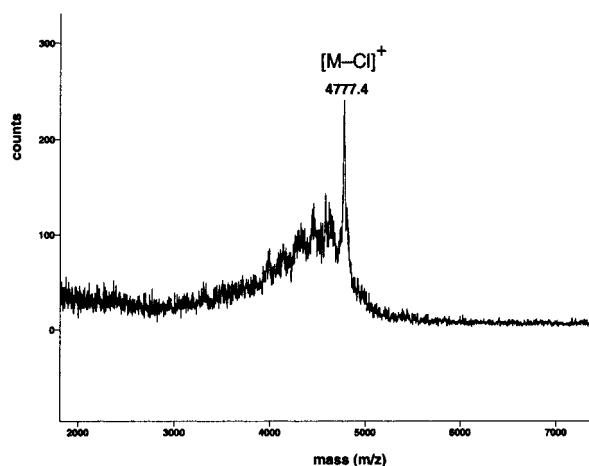


Figure 5. MALDI-TOF-MS spectrum of **13a**.

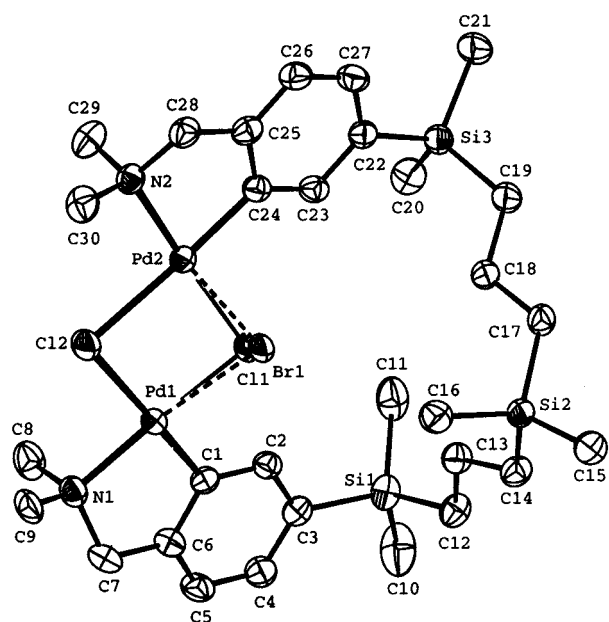


Figure 6. Displacement ellipsoid plot (50% probability) of the molecular structure of **4** showing a small halide disorder. Hydrogen atoms and cocrystallized benzene have been omitted for clarity.

ethanedyl, Scheme 2 and Figure 2) in the carbosilane dendrimer support on the halide-assisted dimer formation between peripheral $[CN-PdCl]$ units was studied by determination of the structures of **4** (C_3 -linker) and **7a** (C_2 -linker) in the solid state. Moreover, the structures of these two molecules can serve as models for sections or wedges present in their corresponding higher generation dendritic analogues. The molecular structures of **4** and **7a** are depicted in Figures 6 and 7, respectively. Relevant bond lengths and distances are summarized in Table 1, while the crystallographic data for **4** and **7a** are given in Table 2.

The structures of **4**¹⁹ and **7a** show the *monomeric*, i.e., *intramolecular*, halide-bridged, *cis*-isomers¹⁶ unlike the dimeric, i.e., *intermolecular*, halide-bridged structure

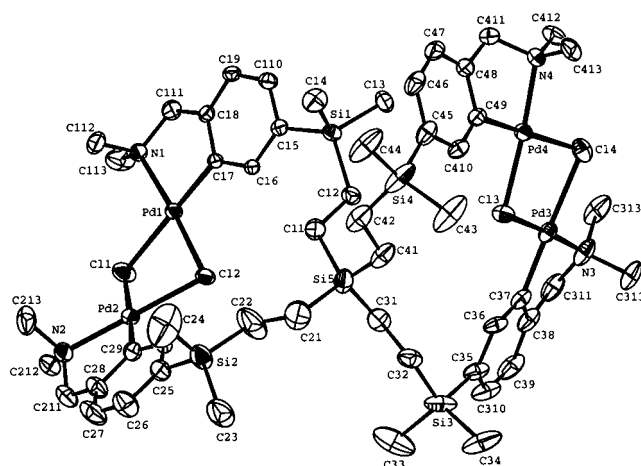


Figure 7. Displacement ellipsoid plot (50% probability) of the molecular structure of **7a**. Hydrogen atoms and cocrystallized pentane have been omitted for clarity.

found for $[CH_2SiMe_2(C_6H_3\{CH_2NMe_2\}-4-(PdCl)-3)]_2$ (abbreviated as $CS-[CN-PdCl]_2$, Figure 1).¹¹ The fact that in both **4** and **7a** *intra* rather than *inter* halide-assisted dimer formation takes place can be explained by the presence of a structurally different carbosilane backbone present in **4** and **7a** when compared with the carbosilane support in $CS-[CN-PdCl]_2$. Apparently, the conformations along the carbosilane support in **4** and **7a** enabled the pendant cyclopalladated complexes to come into each other's proximity to form an *intramolecular* bridge via the chloride anions. It is interesting to note that this *intramolecular* halide-bridging mode does not significantly affect the bond distances and angles found in both derivatives. The interatomic distances Pd–Cl(1), Pd–Cl(2), Pd–C, and Pd–N fall in the same range as was found for $CS-[CN-PdCl]_2$ (vide supra), although some bond lengths show (minor) deviations. The same applies for the bond angles, although it should be noted that in **7a** there were some marked differences found around the palladium metal center, e.g., expressed in the angle Cl(2)–Pd(1)–N(1) of $171.40(19)^\circ$ (cf., $175.76(10)^\circ$ for $CS-[CN-PdCl]_2$).¹¹ Additionally, in both structures some ring strain in the carbosilane interior is present. In **4**, the dihedral angles C(14)–Si(2)–C(17)–C(18) and C(17)–Si(2)–C(14)–C(13) are $73.2(3)^\circ$ and $-66.6(3)^\circ$, respectively, while in **7a** additional distortions are found in the torsion angles Si(5)–C(11)–C(12)–Si(1) and Si(5)–C(41)–C(42)–Si(4), which amounted to $-157.6(5)^\circ$ and $-162.3(5)^\circ$. This increase in conformational distortion in the dendritic backbone of **4** and **7a** as compared with the close to ideal Si–C–C–Si dihedral angles in $CS-[CN-PdCl]_2$ of ca. 179° most probably has to be ascribed to the formation of *intra*- rather than *intermolecular* aggregates.

Molecular mechanics calculations¹⁷ carried out for **4** (Scheme 1) and the [G0] derivatives **5a** and **7a** (Scheme 2) clearly showed that the *cis*-bridging mode is more favored than the *trans*-bridging mode. The latter halide bridging mode leads to more conformational restraints in the Pd_2Cl_2 "square" units. These calculations were in line with those carried out for $CS-[CN-PdCl]_2$,¹¹ and the calculated structures of **4** and **7a** matched well with the results of the crystallographic studies. As compared with the molecular modeling carried out for the model species $(\mu-Cl-[PdCl(C_6H_3\{CH_2NMe_2\}-2-SiMe_3-5)]_2)$ ¹¹ (vide

(19) This compound showed a halide disorder in the solid state. We ascribe this to the presence of LiBr during the crystallization of **4** (see for refinement procedures the Experimental Section). In a separate experiment we prepared bromide-free **4**, the stoichiometry of which was spectroscopically and analytically supported (Experimental Section).

Table 1. Selected Interatomic Distances (Å), Bond Angles (deg), and Dihedral Angles (deg) for the Cyclopalladated Carbosilanes 4 and 7a

4		7a	
Bond Distances			
Pd(1)–Cl(1)	2.305(6)	Pd(1)–Cl(1)	2.461(2)
Pd(1)–Cl(2)	2.462(2)	Pd(1)–Cl(2)	2.334(2)
Pd(1)–C(1)	1.978(4)	Pd(1)–C(17)	1.977(6)
Pd(1)–N(1)	2.068(3)	Pd(1)–N(1)	2.078(6)
Pd(2)–Cl(1)	2.314(6)	Pd(2)–Cl(1)	2.465(2)
Pd(2)–Cl(2)	2.462(2)	Pd(2)–Cl(2)	2.356(2)
Pd(2)–C(24)	1.961(4)	Pd(2)–C(29)	1.955(8)
Pd(2)–N(2)	2.073(3)	Pd(2)–N(2)	2.094(6)
Bond Angles			
Cl(1)–Pd(1)–Cl(2)	86.20(14)	Cl(1)–Pd(1)–Cl(2)	85.72(7)
Cl(1)–Pd(1)–N(1)	177.32(17)	Cl(2)–Pd(1)–N(1)	171.40(19)
Cl(1)–Pd(1)–C(1)	95.05(18)	Cl(2)–Pd(1)–C(17)	96.62(19)
Cl(2)–Pd(1)–N(1)	95.62(9)	Cl(1)–Pd(1)–N(1)	94.76(16)
Cl(2)–Pd(1)–C(1)	175.94(11)	Cl(1)–Pd(1)–C(17)	176.09(19)
N(1)–Pd(1)–C(1)	83.26(14)	N(1)–Pd(1)–C(17)	83.4(2)
Dihedral Angles			
C(14)–Si(2)–C(17)–C(18)	73.2(3)	Si(5)–C(12)–C(13)–Si(1)	–157.6(5)
C(17)–Si(2)–C(14)–C(13)	–6.6(3)	Si(5)–C(21)–C(22)–Si(2)	179.8(6)
C(17)–C(18)–C(19)–Si(3)	–177.4(3)	Si(5)–C(31)–C(32)–Si(3)	172.6(4)
Si(1)–C(12)–C(13)–C(14)	–177.4(2)	Si(5)–C(41)–C(42)–Si(4)	–162.3(4)

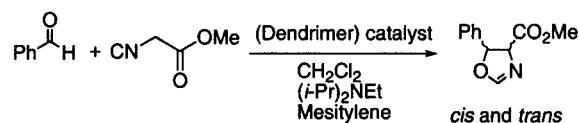
Table 2. Crystal Data and Details of the Structure Determination for 4 and 7a

	4	7
empirical formula	C ₃₀ H ₅₂ N ₂ Si ₃ Pd ₂ Br _{0.15} Cl _{1.85} ·0.5C ₆ H ₆	C ₅₂ H ₈₄ N ₄ Si ₄ Pd ₄ Cl ₄
fw	854.48	1473.21
cryst syst	monoclinic	orthorhombic
space group	<i>P2</i> ₁ / <i>c</i> (no. 14)	<i>Pbca</i> (no. 61)
<i>a</i> /Å	12.746(9)	13.3571(2)
<i>b</i> /Å	14.627(4)	26.9874(4)
<i>c</i> /Å	23.890(16)	39.6304(5)
β /deg	109.50(5)	90
<i>V</i> /Å ³	4198(4)	14285.7(4)
<i>Z</i>	4	8
<i>D</i> _(calc) /[g/cm ³]	1.352	1.401
<i>F</i> (000)	1751	6102
μ (Mo K α)/mm ⁻¹	1.23	1.3
cryst size/mm	0.9 × 0.2 × 0.1	0.06 × 0.31 × 0.38
<i>T</i> /K	150	150
radiation/Å	Mo K α , 0.71073	Mo K α , 0.71073
θ min, max/deg	1.66, 27.47	1.5, 25.0
scan (type and range)/deg	0.65 + 0.35 tan(θ)	
dataset	0.91:27.50; –16:16; –18:0; –20:30	0:15; 0:32; 0:47
no. of tot, unique data, <i>R</i> (int)	11 405, 9601, 0.0366	103609, 12559, 0.0628
no. obsd data [<i>I</i> > 2.0 σ (<i>I</i>)]	7266	7941
<i>N</i> _{ref} , <i>N</i> _{par}	25, 393	0.0506, 0.1862, 0.995
<i>R</i> , <i>wR</i> ² , <i>S</i>	0.039, 0.091, 0.998	0.056, 0.1862, 1.0
<i>w</i> ⁻¹	$\sigma^2(F_o^2) + (0.044P^2)$	$\sigma^2(F_o^2) + (0.100P^2)$ and $P = (F_o^2 + 2F_c^2)/3$
max and av shift/error	0.002, 0.00	0.02, 0.00
min/max resd dens/[e/Å ³]	–0.61, 0.55	–0.90, 1.73

supra), these latter calculations, together with the crystallographic studies, indicate that, on the basis of geometrical considerations, the *cis* bridging mode in these peripheral halide-assisted Pd(II) dimers is more likely to occur. Furthermore, interconversion to the *trans* isomer seems unlikely, as the carbosilane support and the Pd₂Cl₂ “square” units then have to adopt rather strained conformations (cf., the conformation of the carbosilane support in the solid state structures of **4** and **7a**).

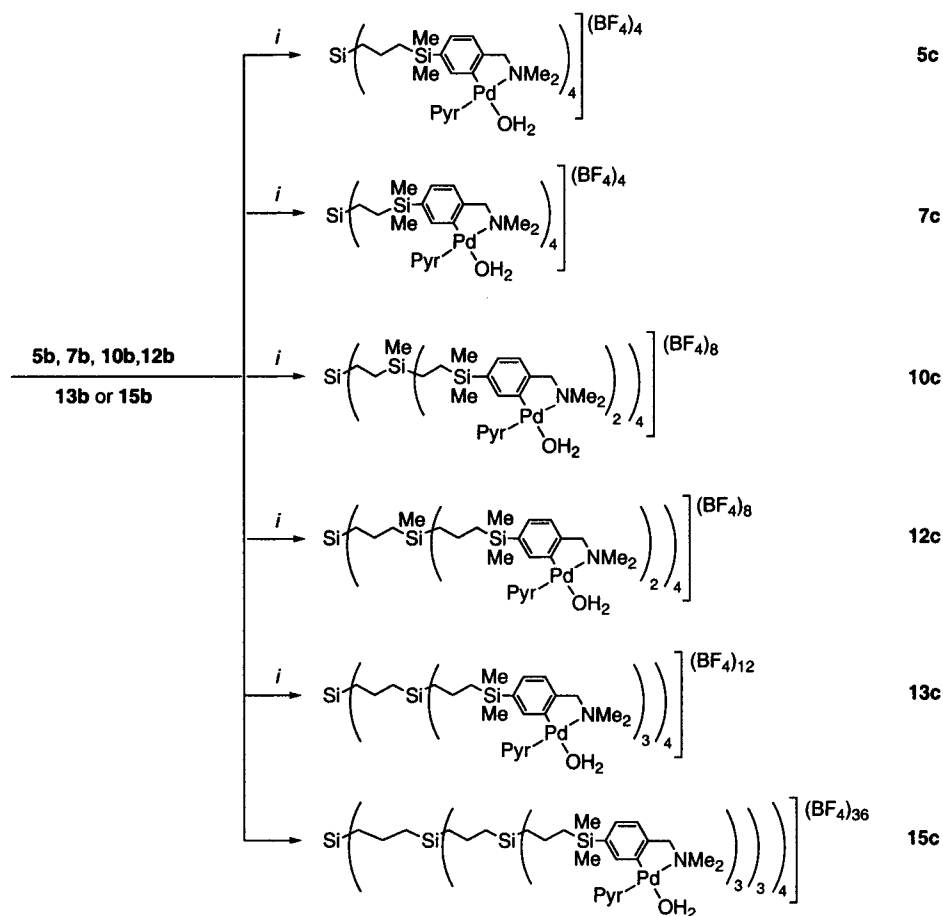
Cyclopalladated Dendrimers as Catalysts for the Aldol Condensation Reaction. The pyridine adducts **5b**, **7b**, **10b**, **12b**, **13b**, and **15b** were used as catalytic precursors in the aldol condensation reaction²⁰ of benzaldehyde with methyl isocyanacetate (Figure 8).

Consequently, **5b**, **7b**, **10b**, **12b**, **13b**, and **15b** were converted in their corresponding (poly)cationic analogues **5c**, **7c**, **10c**, **12c**, **13c**, and **15c** by abstracting the

**Figure 8.** Catalyzed aldol condensation of benzaldehyde and methyl isocyanacetate yielding oxazolines.

halide anion using AgBF₄ in wet acetone (Scheme 8). After appropriate workup involving a thorough filtration of the solutions containing the polycations through a path of Celite to remove the insoluble silver halide salt,²¹

(20) For recent relevant examples concerning aldol condensation catalysis see: (a) Nesper, R.; Pregosin, P. S.; Püntener, K.; Wörle, M. *Helv. Chim. Acta* **1993**, *76*, 2239–2248. (b) Hayashi, T.; Sawamura, M.; Ito, Y. *Tetrahedron* **1999**, *48*, 1999–2012. (c) Longmire, J. M.; Zhang, X.; Shang, M. *Organometallics* **1998**, *17*, 4374–4379. (d) Gorla, F.; Togni, A.; Venanzi, L. M. *Organometallics* **1994**, *13*, 1607–1616. (e) Stark, M. A.; Jones, G.; Richards, C. J. *Organometallics* **2000**, *19*, 1282.

Scheme 8. Synthesis of the Polycationic Palladium(II) Dendrimers^a

^a Reagents and conditions: (i) AgBF₄, wet acetone.

Table 3. Aldol Condensation of Benzaldehyde and Methyl Isocyanoacetate Catalyzed by the Cationic Dendrimer Species and the Parent Model Complex 16^a

catalyst	mol % Pd	time, h	conversion % ^b	% <i>trans</i> ^b	% <i>cis</i> ^b	TOF/Pd ^c	TON/Pd
5c , [G0]-Pd ₄	0.99	24	>99	64	36	26	101
7c , [G0] [#] -Pd ₄	1.04	24	97	64	36	20	93
10c , [G1] [#] -Pd ₈	0.93	24	55	57	43	9	59
12c , [G1]-Pd ₈	0.93	24	96	61	39	22	104
13c , [G1]-Pd ₁₂	0.89	24	97	58	42	27	109
15c , [G2]-Pd ₃₆	1.19	24	99	63	37	11	83
16 , Pd ₁	0.94	24	97	64	36	26	103
no catalyst		24	5	(n.d.) ^d	(n.d.) ^d		

^a Reaction carried out in CH₂Cl₂ (10 mL) at RT with ca. 10 mol % Hünigs base (=NEt(*i*-Pr)₂). Selectivity for these reactions is ≥98% determined with mesitylene as an internal standard. ^b Calculated by specific ¹H NMR spectroscopic signal integration using the internal standard. ^c Per hour during the first hour. ^d Not determined.

these polycationic materials were isolated as rather insoluble and very hygroscopic, yellow to brown powders (isolated yields: 80% to quantitative). Identification of these cations was carried out by means of ¹H NMR spectroscopy in acetone-*d*₆.

Despite the low solubility of the polycationic dendritic species **5c**, **7c**, **10c**, **12c**, and **13c** in solvents such as CHCl₃, CH₂Cl₂, and benzene, these species dissolved readily in CH₂Cl₂ in the presence of methyl isocyanoacetate,²² i.e., under the catalytic conditions. During the catalytic runs with the dendrimer catalysts no formation of palladium black was observed, which points to the increased stability of the cationic palladium complexes under the catalytic conditions.²³

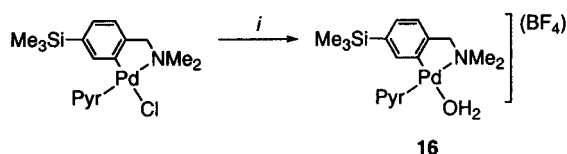
(21) This is important to note, as the precipitated silver salt can also act as a (heterogeneous) Lewis acid catalyst.

The catalytic results are presented in Table 3 and are compared with the activity of the corresponding mononuclear, model compound [Pd(C₆H₃{CH₂NMe₂}-2-(SiMe₃)-5)(pyr)(H₂O)]BF₄ (**16**) (Scheme 9) as well as the uncatalyzed reaction. The parent model compound **16** was produced similarly to the dendrimer compounds starting from the known pyridine adduct [PdCl(C₆H₃{CH₂NMe₂}-2-(SiMe₃)-5)(pyr)].¹¹

(22) It is generally assumed that the initial step in this particular aldol condensation reaction is the replacement of the labile H₂O ligand by the isocyanoacetate substrate. For general mechanistic aspects concerning this kind of catalysis see: (a) Richard, J. P.; Nagorski, R. *W. J. Am. Chem. Soc.* **1999**, *121*, 4763–4770. (b) Saito, S.; Yamamoto, H. *Chem. Eur. J.* **1999**, *5*, 1959–1962.

(23) Though during the catalytic runs no palladium black formation was observed, the polycationic dendrimers are susceptible toward decomposition in (acetone) solution as well as in the solid state, as illustrated by darkening of the samples under these conditions.

Scheme 9. Synthesis of the Model Cationic Species 16^a



^a Reagents and conditions: (i) AgBF₄, wet acetone.

The smaller dendritic species and the model compound showed only minor differences when both the initial reaction rate (TOF) and total turnover number (TTN) after 24 h per Pd(II) site are compared. For the higher generation Pd-dendrimer catalysts **10c**, **12c**, and **13c** an almost negligible decrease in the *trans/cis* ratio of the oxazoline product formed in this particular aldol condensation reaction was observed. An interesting difference in reaction rate was noted between the [G1] generation species **10c** and **12c**. Apparently the more compact structure of **10c** due to the C₂-linkers instead of C₃-linkers present in **12c** has a distinct effect on the catalytic activity of the individual Pd(II) sites. These data are in line with the fact that also for the more crowded [G2] generation species **15c** a decrease in catalytic activity per Pd(II) site has been observed.

Conclusions

In summary, this work demonstrates that different generations of cyclometalated dendrimers can be readily prepared using a multiple C–H bond activation procedure. For the palladated dendrimers, a unique halide-assisted dimer formation between the peripheral Pd(II) complexes was encountered. For the [G1] dendrimers **10a** and **12a**, an unexpected formation of a temperature-independent ratio of isomeric forms of these dimeric Pd(II) complexes was observed, which was supported by VT and (2D) NMR spectroscopic studies. An unexpected isomerization of the peripheral pyridine ligands was proposed for the [G2] generation dendrimer **15b**, which is most likely due to steric congestion at the periphery of the dendrimer.

The Pd(II) metallodendrimers were successfully applied as homogeneous catalysts in the aldol condensation between benzaldehyde and methyl isocyanacetate. The catalytic data indicated that more rigid/crowded dendritic catalysts become slower accompanied by a slight selectivity change in the *trans/cis* ratio of the oxazoline product. These novel Pd(II) dendrimers hold promise for a wide range of C–C or cross-coupling reactions.²⁴ Therefore, we are currently focused on the use of these Pd(II) dendrimers in a range of other (catalytic) applications.

Experimental Section

General Comments. All air-sensitive manipulations were carried out in an inert (nitrogen) atmosphere using standard Schlenk techniques. All solvents were carefully dried and

distilled prior to use. Standard chemicals were purchased from Acros Chimica or Aldrich and used without further purification. Compounds [G0]-(SiMe₂-CN)₄,¹⁰ [G1]-(SiMe₂-CN)₁₂,¹⁰ [G0]⁺-(SiMeCl)₄,^{14a,c} [G1]⁺-(SiMe₂Cl)₈,^{14a} and [G2]-(SiMe₂Cl)₃₆,^{14b} were prepared according to previously reported procedures. ¹H, ¹³C{¹H}, and ²⁹Si{¹H} NMR spectroscopic measurements were performed on a Varian Inova/Mercury 200 or 300 MHz spectrometer at RT, and chemical shifts (δ) are given in ppm with TMS as an external standard unless otherwise noted. The MALDI-TOF mass spectra were acquired using a Voyager-DE BioSpectrometry workstation (PerSeptive Biosystems Inc., Framingham, MA) mass spectrometer. Sample solutions with an approximate concentration of 20–30 mg/mL in THF or CH₂Cl₂ were prepared. The matrix was 3,5-dihydroxybenzoic acid (DHB) or 9-nitroanthracene (9-NA) with approximate concentrations of 20–30 mg/mL. A 0.2 μL portion of the sample solution and 0.2 μL of the matrix solution were combined and placed on a golden MALDI target and analyzed after evaporation of the solvents. Elemental analyses were performed by Dornis und Kolbe, Mülheim a/d Ruhr, Germany.

[G0]^{1/2}-(SiMe₂Cl)₂ (1). To a mixture of dimethylallylsilane (2.60 g, 18.5 mmol) and chlorodimethylsilane (ca. 8 mL, 6.8 g, 72 mmol) were added some drops of a concentrated solution of [Bu₄N]₂PtCl₆ in EtOH. The reaction mixture was stirred and after ¹H NMR spectroscopy showed that the conversion into **1** was complete (≤17 h). Excess HSiMe₂Cl was removed under reduced pressure to give a clear, colorless liquid (6.08 g, 18.5 mmol, >99%). ¹H NMR (C₆D₆): δ 1.47–1.35 (m, 4H, (CH₃)SiCH₂CH₂), 0.79–0.71 (m, 4H, CH₂Si(CH₃)₂Cl), 0.60–0.49 (m, 4H, (CH₃)SiCH₂CH₂), 0.25 (s, 12H, CH₂Si(CH₃)₂Cl), –0.05 (s, 6H, (CH₃)SiCH₂CH₂). ¹³C{¹H} NMR (C₆D₆): δ 23.6, 19.5, 18.0 (3 × CH₂), 1.8 (Si(CH₃)₂Cl), –3.2 ((CH₃)₂Si). This compound was converted to the moisture stable hydride **2**.

[G0]^{1/2}-(SiMe₂H)₂ (2). A solution of **1** (1.77 g, 5.37 mmol) in Et₂O (30 mL) was slowly added to a suspension of LiAlH₄ (0.73 g, 19.2 mmol) in Et₂O (20 mL) at 0 °C. The reaction mixture was stirred at RT for 20 h and then cautiously treated with H₂O and 4 M HCl in order to dissolve most of the precipitated salts. After filtration, the H₂O layer was extracted with pentane (2 × 50 mL). The combined organic layers were dried on MgSO₄ and concentrated in vacuo. Flash distillation at reduced pressure gave a colorless liquid (1.03 g, 3.95 mmol, 74%). ¹H NMR (C₆D₆): δ 4.15–4.04 (m, 2H, *J* = 3.5 Hz, CH₂-Si(CH₃)₂H), 1.53–1.37 (m, 4H, (CH₃)SiCH₂CH₂), 0.70–0.57 (m, 8H, (CH₃)SiCH₂CH₂H and CH₂Si(CH₃)₂H), 0.06 (d, 12H, *J* = 3.7 Hz, CH₂Si(CH₃)₂H), –0.01 (s, 6H, (CH₃)SiCH₂CH₂–). ¹³C{¹H} NMR (C₆D₆): δ 19.8, 19.4, 19.0 (3 × CH₂), –3.1 ((CH₃)₂Si), –4.3 (Si(CH₃)₂H). GC–MS (EI): *m/z* 159 (M – (CH₂)₃Si(CH₃)₂H)⁺. Anal. Calcd for C₁₂H₃₂Si₃: C 55.30, H 12.37, Si 32.33. Found: C 55.45, H 12.30, Si 32.46.

[G0]^{1/2}-(SiMe₂-CN)₂ (3). This compound was prepared using a previously reported method.¹⁰ To a solution of 1-bromo-4-[(dimethylamino)methyl]benzene (1.99 g, 9.29 mmol) in Et₂O (25 mL) was added *t*-BuLi (ca. 10 mL of a 1.5 M solution in pentane, 15 mmol) at –78 °C. After stirring the white suspension for 15 min, a solution of **1** (1.13 g, 3.43 mmol) in Et₂O (30 mL) was added, and the resulting mixture was allowed to reach RT. After purification by Kugelrohr distillation (0.2 mmHg, 170–190 °C) a clear, yellow viscous oil was obtained (1.37 g, 2.60 mmol, 76%). ¹H NMR (C₆D₆): δ 7.54 (d, 4H, *J* = 7.8 Hz, Ar–H), 7.43 (d, 4H, *J* = 7.6 Hz, Ar–H), 3.30 (s, 4H, CH₂N), 2.11 (s, 12H, N(CH₃)₂), 1.52–1.40 (m, 4H, (CH₃)₂SiCH₂CH₂), 0.91–0.83 (m, 4H, CH₂Si(CH₃)₂Ph), 0.64–0.56 (m, 4H, (CH₃)₂SiCH₂CH₂), 0.29 (s, 12H, CH₂Si(CH₃)₂Ph), –0.03 (s, 6H, (CH₃)₂SiCH₂CH₂). ¹³C{¹H} NMR (C₆D₆): δ 140.8, 138.0, 134.0, 128.7 (4 × Ar–C), 64.6 (CH₂N), 45.5 (N(CH₃)₂), 20.8, 20.3, 19.0 (3 × CH₂), –2.6 (Si(CH₃)₂Ph), –3.0 ((CH₃)₂Si). FAB–MS: *m/z* 527 (M + H)⁺, 525 (M – H)⁺. Anal. Calcd for C₃₀H₅₄N₂Si₃: C 68.37, H 10.33, N 5.32, Si 15.99. Found: C 68.52, H 10.40, N 5.28, Si 15.83.

(24) (a) Motoyama, Y.; Narusawa, H.; Nishiyama, H. *Chem. Commun.* **1999**, 131–132. (b) Motoyama, Y.; Mikami, Y.; Kawakami, H.; Aoki, K.; Nishiyama, H. *Organometallics* **1999**, *18*, 3584–3588. (c) Weissmann, H.; Milstein, D. *Chem. Commun.* **1999**, 1901–1902. (d) Leung, P. H.; Ng, K. H.; Li, Y.; White, A. J. P.; Williams, D. J. *Chem. Commun.* **1999**, 2435–2436. (e) Stark, M. A.; Richards, C. *Tetrahedron Lett.* **1997**, *38*, 5881–5884.

[G0]^{1/2}-(SiMe₂-CN-PdCl)₂ (4). Compound **3** (1.57 g, 2.98 mmol) was dissolved in MeOH (80 mL), and to this solution was added solid Pd(OAc)₂ (1.34 g, 5.96 mmol) and NaOAc (0.50 g, 6.10 mmol). The resulting mixture was stirred for 20 min, and during this time the color of the suspension changed from red-brown to black. The mixture was filtered through a path of Celite, and the resulting clear orange to red solution was treated with LiCl (0.33 g, 7.78 mmol). Immediately a yellow-orange suspension was obtained which was concentrated in vacuo. To the residue was added benzene (100 mL) and H₂O (100 mL). The organic layer was separated and the H₂O layer extracted with benzene (2 × 50 mL). The combined benzene layers were washed with saturated aqueous NaCl (2 × 50 mL) and then dried on K₂CO₃. After removal of the solvent under reduced pressure, the product was washed with pentane (2 × 50 mL) and dried in vacuo to give a yellow to orange solid (2.00 g, 2.47 mmol, 83%). Analytically pure **4** was obtained by slow diffusion of pentane into a concentrated solution of **4** in benzene. These crystals were suitable for an X-ray molecular structure determination. ¹H NMR (C₆D₆): δ 8.03 (s, 2H, Ar-H_{ortho}), 7.30 (d, 2H, J = 7.1 Hz, Ar-H), 6.74 (d, 2H, J = 7.2 Hz, Ar-H), 3.12 (s, 4H, CH₂N), 2.28 (s, 12H, N(CH₃)₂), 1.82–1.72 (m, 4H, (CH₃)₂SiCH₂CH₂), 0.99–0.91 (m, 4H, CH₂Si(CH₃)₂-Ph), 0.79–0.71 (m, 4H, (CH₃)₂SiCH₂CH₂), 0.38 (s, 12H, CH₂-Si(CH₃)₂Ph), 0.03 (s, 6H, (CH₃)₂SiCH₂CH₂). ¹³C{¹H} NMR (C₆D₆): δ 148.0, 143.9, 139.6, 137.0, 130.0, 121.3 (6 × Ar-C), 72.9 (CH₂N), 52.0 (N(CH₃)₂), 21.3, 20.8, 19.4 (3 × CH₂), -2.4 (Si(CH₃)₂Ph), -2.5 ((CH₃)₂Si). FAB-MS: *m/z* 808 (M)⁺, 773 (M - Cl)⁺. Anal. Calcd for crystalline C₃₀H₅₂N₂Si₃Pd₂Cl₂·C₆H₆: C 48.75, H 6.59, N 3.16, Si 9.50. Found: C 48.79, H 6.65, N 3.22, Si 9.63.

[G0]-(SiMe₂-CN-PdCl)₄ (5a). This compound was prepared in a way similar to that described for **4**. To a solution of [G0]-SiMe₂-CN (0.59 g, 0.61 mmol) in MeOH (40 mL) was added solid Pd(OAc)₂ (0.69 g, 3.07 mmol). The resultant red-colored mixture was stirred for 6 h, upon which LiCl (0.69 g, 16.3 mmol) was added. After workup, a yellow solid was isolated (0.47 g, 0.31 mmol, 51%). ¹H NMR (C₆D₆): δ 8.00 (s, 4H, Ar-H_{ortho}), 7.28 (d, 4H, J = 7.0 Hz, Ar-H), 6.71 (d, 4H, J = 7.2 Hz, Ar-H), 3.17 (s, 8H, CH₂N), 2.31 (s, 24H, N(CH₃)₂), 1.78–1.62 (m, 8H, SiCH₂CH₂CH₂), 1.01–0.93 (m, 8H, SiCH₂-CH₂CH₂), 0.84–0.76 (m, 8H, SiCH₂CH₂CH₂), 0.37 (s, 12H, Si(CH₃)₂). ¹³C{¹H} NMR (C₆D₆): δ 147.66, 143.77, 139.60, 137.39, 130.00, 121.17 (6 × Ar-C), 72.97 (CH₂N), 52.03 (N(CH₃)₂), 21.06, 19.38, 18.01 (3 × CH₂), -2.23 (Si(CH₃)₂). FAB-MS: *m/z* 1529.4 (M)⁺, 1493.4 (M - Cl)⁺. Anal. Calcd for C₅₆H₉₂N₄Si₅Pd₄Cl₄: C 43.98, H 6.06, N 3.66, Si 9.18. Found: C 43.82, H 6.15, N 3.83, Si 9.32.

[G0]-(SiMe₂-CN-PdCl{pyr})₄ (5b). To a solution of **5a** (62.1 mg, 0.0406 mmol) in CH₂Cl₂ (20 mL) was added an excess of pyridine (ca. 3 mL). The clear yellowish reaction mixture was stirred for 16 h. Then the volatiles were removed in vacuo, and the product was precipitated from CH₂Cl₂/pentane, washed with pentane (10 mL), and dried to give **5b** as an off-white solid (66.3 mg, 0.0359 mmol, 88%). ¹H NMR (CDCl₃): δ 8.87 (d, J = 3.6 Hz, 8H, pyr-H_{ortho}), 7.79 (br, 4H, pyr-H), 7.33 (br, 8H, pyr-H), 7.08 (d, J = 7.0 Hz, 4H, Ar-H), 6.94 (d, J = 7.0 Hz, 4H, Ar-H), 6.04 (s, 4H, Ar-H_{ortho}), 3.95 (s, 8H, CH₂N), 2.93 (s, 24H, N(CH₃)₂), 1.15 (br m, 8H, dendrimer CH₂), 0.55 (br m, 8H, dendrimer CH₂), 0.38 (br m, 8H, dendrimer CH₂), -0.04 (s, 24H, Si(CH₃)₂). ¹³C{¹H} NMR (CDCl₃): δ 153.90, 148.85, 148.45, 137.86, 137.29, 136.74, 129.83, 125.13, 121.30 (9 × Ar-C), 74.15 (CH₂N), 53.04 (N(CH₃)₂), 20.79, 18.66, 17.63 (3 × dendrimer CH₂), -2.75 (Si(CH₃)₂). Anal. Calcd for C₇₆H₁₁₂N₈Si₅Pd₄Cl₄·4CH₂Cl₂: C 43.97, H 5.53, N 5.13. Found: C 43.95, H 5.86, N 4.75.

[G0]^{1/2}-(SiMe₂-CN)₄ (6). This compound was prepared using a previously reported method.¹⁰ To a solution of 1-bromo-4-[(dimethylamino)methyl]benzene (4.61 g, 21.5 mmol) in Et₂O (40 mL) was added *t*-BuLi (ca. 26 mL of a 1.5 M solution in pentane, 39 mmol) at -78 °C. After stirring the white

suspension for 15 min, a solution of [G0]^{1/2}-(SiMe₂Cl)₄ (2.20 g, 4.28 mmol) in Et₂O (30 mL) was added, and the resulting mixture was allowed to reach room temperature and then stirred for 18 h. After purification by Kugelrohr distillation (0.2 mmHg, 130 °C) a clear, yellow viscous oil was obtained (2.79 g, 3.07 mmol, 72%), which solidified at RT. ¹H NMR (C₆D₆): δ 7.53 (d, 8H, J = 8.4 Hz, Ar-H), 7.41 (d, 8H, J = 8.1 Hz, Ar-H), 3.30 (s, 8H, CH₂N), 2.10 (s, 24H, N(CH₃)₂), 0.64 (m, 16H, SiCH₂CH₂Si, AA'BB' pattern), 0.29 (s, 24H, Si(CH₃)₂). ¹³C{¹H} NMR (C₆D₆): δ 140.70, 137.71, 133.97, 128.69 (4 × Ar-C), 64.56 (CH₂N), 45.48 (N(CH₃)₂), 8.22, 3.01 (2 × CH₂), -3.3 (Si(CH₃)₂). FAB-MS: *m/z* 909.5 (M + H)⁺. Anal. Calcd for C₅₂H₈₈N₄Si₅: C 68.65, H 9.75, N 6.16, Si 15.44. Found: C 68.70, H 9.84, N 6.10, Si 15.53.

[G0]^{1/2}-(SiMe₂-CN-PdCl)₄ (7a). This compound was prepared analogously to **4**. To solution of **6** (0.50 g, 0.55 mmol) in MeOH (30 mL) was added Pd(OAc)₂ (0.55 g, 2.45 mmol), and the resultant red to orange colored mixture was stirred for 16 h followed by addition of LiCl (0.47 g, 11.1 mmol). After appropriate workup, a slightly yellow solid was obtained (0.73 g, 0.50 mmol, 91%). Crystals suitable for an X-ray molecular structure determination were obtained by slow diffusion of pentane into a solution of **6** in benzene. ¹H NMR (C₆D₆): δ 8.09 (s, 4H, Ar-H_{ortho}), 7.26 (d, 4H, J = 7.4 Hz, Ar-H), 6.71 (d, 4H, J = 7.2 Hz, Ar-H), 3.13 (s, 8H, CH₂N), 2.24 (s, 24H, N(CH₃)₂), 1.04–0.87 (m, 16H, SiCH₂CH₂Si, AA'BB' pattern), 0.30 (s, 24H, Si(CH₃)₂). ¹³C{¹H} NMR (C₆D₆): δ 147.71, 143.99, 139.44, 137.38, 130.24, 121.22 (6 × Ar-C), 72.84 (CH₂N), 52.01 (N(CH₃)₂), 8.00, 3.71 (2 × CH₂), -2.55 (Si(CH₃)₂). FAB-MS: *m/z* 1472.7 (M)⁺, 1437.0 (M - Cl)⁺. Anal. Calcd for C₅₂H₈₄N₄-Si₅Pd₄Cl₄: C 42.40, H 5.75, N 3.80, Si 9.53. Found: C 42.58, H 5.66, N 3.72, Si 9.64.

[G0]^{1/2}-(SiMe₂-CN-PdCl{pyr})₄ (7b). A mixture of **7a** (63.3 mg, 0.0430 mmol) and an excess of pyridine (3 mL) in CH₂Cl₂ (25 mL) was stirred for 19 h. Then the solvent was removed in vacuo and the product washed with pentane (50 mL) to give **7b** as a white solid. Yield: 66.0 mg (0.0369 mmol, 86%). ¹H NMR (CDCl₃): δ 8.79 (d, J = 5.2 Hz, 8H, pyr-H_{ortho}), 7.65 (t, J = 7.6 Hz, 4H, pyr-H), 7.15–7.06 (m, 12H, pyr-H + Ar-H), 6.95 (d, J = 7.0 Hz, 4H, Ar-H), 6.02 (s, 4H, Ar-H_{ortho}), 3.98 (s, 8H, CH₂N), 2.96 (s, 24H, N(CH₃)₂), 0.16 (m, 16H, dendrimer CH₂), -0.04 (s, 24H, Si(CH₃)₂). ¹³C{¹H} NMR (CDCl₃): δ 153.75, 148.97, 137.69, 137.44, 136.49, 129.80, 125.10, 121.40 (8 × Ar-C, one signal most probably coincident with another one), 74.10 (CH₂N), 53.11 (N(CH₃)₂), 8.02, 2.81 (2 × dendrimer CH₂), -3.48 (Si(CH₃)₂). Anal. Calcd for C₇₂H₁₀₄N₈Si₅Pd₄Cl₄: C 48.32, H 5.86, N 6.26, Cl 7.92. Found: C 48.39, H 5.93, N 6.17, Cl 8.06.

G1^{1/2}-(CH=CH₂)₈ (8). To a solution of vinylmagnesium bromide (100 mL of a 1.0 M solution in THF, 100 mmol) was added dropwise [G0]^{1/2}-(SiMeCl₂)₄ (2.72 g, 4.56 mmol) dissolved in THF (200 mL). The orange-colored mixture was stirred for 1.5 h, and ¹H NMR spectroscopy showed at this stage that the reaction was complete. The mixture was carefully quenched with H₂O, after which most of the THF was removed in vacuo. The product was extracted with pentane (250 mL) and the organic layer was washed with H₂O (2 × 250 mL), dried over Na₂SO₄, and filtered. After concentration under reduced pressure a clear yellow, viscous oil was obtained (1.53 g, 2.89 mmol, 63%). ¹H NMR (C₆D₆): δ 6.20 (dd, ³J = 9.8 Hz, 8H, CH=CH₂), 5.99 (d, ³J = 9.8 Hz, ²J = 2.8 Hz, 8H, CH=CH₂), 5.76 (d, ³J = 13.4 Hz, ²J = 2.8 Hz, 8H, CH=CH₂), 0.67 (m, 16H, -CH₂CH₂-), 0.16 (s, 12H, SiCH₃). ¹³C{¹H} NMR (C₆D₆): δ 137.01, 133.30 (CH=CH₂), 6.71, 3.11 (-CH₂CH₂-), -5.53 (SiCH₃). ²⁹Si{¹H} NMR (C₆D₆): δ 10.52 (Si_{core}), -10.82 (SiCH₃). MALDI-TOF-MS: *m/z* 635.08 and 637.00 (M + ¹⁰⁷Ag/¹⁰⁹Ag)⁺. Calcd for (M + ¹⁰⁷Ag/¹⁰⁹Ag)⁺: *m/z* 635.20 and 637.19. Anal. Calcd for C₂₈H₅₂Si₅: C 63.55, H 9.91, Si 26.54. Found: C 63.66, H 9.98, Si 26.36.

[G1]^{1/2}-(SiMe₂-CN)₈ (9). This compound was prepared using a previously reported procedure.¹⁰ To a solution of 1-bromo-4-[(dimethylamino)methyl]benzene (2.45 g, 11.4 mmol) in Et₂O

(25 mL) was added *t*-BuLi (ca. 14.5 mL of a 1.5 M solution in pentane, 21.8 mmol) at -78 °C. After stirring the white suspension for 15 min, a solution of $[G1]^\#-(SiMe_2Cl)_8$ (1.14 g, 0.89 mmol) in Et₂O (30 mL) was added, and the resulting mixture was allowed to reach room temperature. After purification by Kugelrohr distillation (0.2 mmHg, 110–120 °C) a clear yellow, viscous oil was obtained (1.55 g, 0.75 mmol, 84%). ¹H NMR (C₆D₆): δ 7.57 (d, $J = 7.8$ Hz, 16H, Ar–H), 7.44 (d, $J = 7.5$ Hz, 16H, Ar–H), 3.32 (s, 16H, CH₂N), 2.11 (s, 48H, N(CH₃)₂), 0.73 (m, 48H, dendrimer –CH₂–), 0.32 (s, 48H, Si(CH₃)₂), 0.07 (s, 12H, SiCH₃). ¹³C{¹H} NMR (C₆D₆): δ 140.81, 137.81, 134.11, 128.81 (4 × Ar–C), 64.69 (CH₂N), 45.59 (N(CH₃)₂), 8.51, 5.46, 5.17, 3.21 (4 × dendrimer CH₂), –3.08 (Si(CH₃)₂), –6.14 (SiCH₃). ²⁹Si{¹H} NMR (C₆D₆): δ 10.17 (Si_{core}), 8.56 (SiCH₃), –1.51 (Si(CH₃)₂). MALDI-TOF-MS: m/z 2076.21 (M⁺). Calcd for M⁺: m/z 2076.13. Anal. Calcd for C₁₁₆H₂₀₄N₈Si₁₃: C 67.11, H 9.90, N 5.40, Si 17.59. Found: C 67.27, H 9.96, N 5.54, Si 17.47.

[G1][#]-(SiMe₂-CN-PdCl)₈ (10a). To a solution of **9** (0.34 g, 0.16 mmol) in MeOH (80 mL) was added solid Pd(OAc)₂ (0.36 g, 1.60 mmol). The resultant suspension was stirred for 23 h, after which LiCl (0.24 g, 5.80 mmol) was added. This mixture was then stirred for 0.5 h, and subsequently the solvent was removed in vacuo. The solid residue was extracted with CH₂-Cl₂ (2 × 100 mL), washed with H₂O (75 mL), and dried on Na₂SO₄. After filtration and drying under reduced pressure, a light orange-colored solid was isolated (0.42 g, 0.13 mmol, 80%). ¹H NMR (CDCl₃): δ 7.41 (s, 8H, Ar–H_{ortho}), 7.09 (d, $J = 6.8$ Hz, 8H, Ar–H), 6.84 (d, $J = 7.2$ Hz, 8H, Ar–H), 3.91 (s, 16H, CH₂N), 2.84 (s, 48H, N(CH₃)₂), 0.88–0.33 (br m, 48H, dendrimer –CH₂–), 0.18 (s, 48H, Si(CH₃)₂), –0.14 (s, 12H, SiCH₃). ¹³C{¹H} NMR (CDCl₃): δ 147.44, 143.03, 138.81, 137.60, 130.27, 121.37 (6 × Ar–C), 73.45 (CH₂N), 52.97 (N(CH₃)₂), 7.94, 5.43, 2.87 (3 × dendrimer CH₂), –2.19, –2.89 (1 × dendrimer CH₂ + Si(CH₃)₂), –6.15 (SiCH₃). MALDI-TOF-MS: m/z 3168.12 (M – Cl)⁺. Calcd for (M – Cl)⁺: m/z 3167.54. Anal. Calcd for C₁₁₆H₁₉₆N₈Si₁₃Pd₈Cl₈·4CH₂Cl₂: C 40.68, H 5.80, N 3.16. Found: C 40.80, H 5.84, N 3.09.

[G1][#]-(SiMe₂-CN-PdCl{pyr})₈ (10b). To a solution of **10a** (0.19 g, 0.059 mmol) in CH₂Cl₂ (10 mL) was added an excess of pyridine (0.5 mL). The mixture was stirred for 10 min and then evaporated to dryness. The solid residue was washed with pentane (20 mL) and dried in vacuo to yield a slightly salmon colored powder (>99%). ¹H NMR (CDCl₃): δ 8.83 (br, 16H, pyr-H_{ortho}), 7.76 (br m, 8H, pyr-H), 7.34 (br m, 16H, pyr-H), 6.04 (s, 8H, Ar–H_{ortho}), 3.98 (br s, 16H, CH₂N), 2.95, 2.85 (2 × s, 48H, N(CH₃)₂), 0.25 (br, dendrimer –CH₂–), –0.04 (br s, 48H, Si(CH₃)₂), –0.16, –0.22 (br, 12H, SiCH₃). ¹³C{¹H} NMR (CDCl₃): δ 147.44, 143.03, 138.81, 137.60, 130.27, 121.37 (6 × Ar–C), 73.45 (CH₂N), 52.97 (N(CH₃)₂), 7.93, 5.43, 2.87 (3 × dendrimer CH₂), –2.19, –2.88 (SiCH₃ + Si(CH₃)₂), –6.15 (dendrimer CH₂). Anal. Calcd for C₁₅₆H₂₃₆N₁₆Si₁₃Pd₈Cl₈: C 48.85, H 6.20, N 5.84, Cl 7.39. Found: C 48.72, H 6.25, N 5.78, Cl 7.31.

[G1]-(SiMe₂-CN)₈ (11). This compound was prepared using a method described in the literature.¹⁰ To a solution of 4-bromo-[(dimethylamino)methyl]benzene (2.33 g, 10.9 mmol) in Et₂O was added *t*-BuLi (14 mL of a 1.5 M solution in pentane, 21 mmol) at -78 °C. The resultant suspension was stirred for 10 min, whereupon $[G1]-(SiMe_2Cl)_8$ (1.42 g, 0.98 mmol) was added as a solution in Et₂O (30 mL). After appropriate workup and further purification by Kugelrohr distillation (130–140 °C, 2–3 h) a bright yellow, viscous oil was obtained (1.70 g, 0.76 mmol, 78%). ¹H NMR (C₆D₆): δ 7.55 (d, 16H, $J = 8.0$ Hz, Ar–H), 7.42 (d, 16H, $J = 7.6$ Hz, Ar–H), 3.31 (s, 16H, CH₂N), 2.11 (s, 48H, N(CH₃)₂), 1.55 (m, –CH₂–, dendrimer support), 0.92 (m, –CH₂–, dendrimer support), 0.72 (m, –CH₂–, dendrimer support), 0.31 (s, 48H, Si(CH₃)₂), 0.06 (s, 12H, SiCH₃). ¹³C{¹H} NMR (C₆D₆): δ 141.08, 138.36, 134.29, 129.08 (4 × Ar–C), 65.02 (CH₂N), 45.93 (N(CH₃)₂), 21.38, 19.57, 19.52 (3 × outer –CH₂–, dendrimer support), 20.14, 19.75, 18.74 (3 × inner

–CH₂–, dendrimer support), –2.07 (Si(CH₃)₂), –4.17 (SiCH₃). ²⁹Si{¹H} NMR (C₆D₆): δ 1.44 (Si(CH₃)₂), 1.04 (SiCH₃), –3.86 (Si-aryl). MALDI-TOF-MS: m/z 2245.81 (M + H)⁺. Calcd: 2245.47 (M + H)⁺. Anal. Calcd for C₁₂₈H₂₂₈N₈Si₁₃: C 68.50, H 10.24, N 4.99, Si 16.27. Found: C 68.50, H 10.29, N 5.06, Si 16.20.

[G1]-(SiMe₂-CN-PdCl)₈ (12a). This compound was prepared as described for **10**. To a solution of **11** (0.46 g, 0.21 mmol) in MeOH (60 mL) was added Pd(OAc)₂ (0.50 g, 2.21 mmol). Within 10 min a suspension was observed, which was stirred for 2 h, and subsequently LiCl (0.23 g, 5.56 mmol) was added. After a workup similar to that described for **11**, a yellow to orange solid was isolated (0.45 g, 0.13 mmol, 65%). ¹H NMR (CDCl₃, mixture of isomers): δ 7.41 (br s, 8H, Ar–H_{ortho}), 7.08 (br d, $J = 5.0$ Hz, 8H, Ar–H), 6.83 (br d, $J = 6.8$ Hz, 8H, Ar–H), 3.99 (d, $J(A-B) = 14.0$ Hz, CH₂N), 3.90 (br s, CH₂N), 3.80 (d, $J(A-B) = 13.8$ Hz, CH₂N), 2.83 (br s, 48H, N(CH₃)₂), 1.31 (br, –CH₂–, dendrimer backbone), 0.76 (br, –CH₂–, dendrimer backbone), 0.55 (br, –CH₂–, dendrimer backbone), 0.22 (br s, 48H, Si(CH₃)₂), –0.13 (br s, 12H, SiCH₃). ¹³C{¹H} NMR (CDCl₃): δ 147.44, 142.93, 138.92, 137.76, 129.94, 121.29 (6 × Ar–C), 73.43 (CH₂N), 52.87 (N(CH₃)₂), 20.93, 20.69, 19.89, 18.95, 18.76, 17.89 (6 × –CH₂–, dendrimer backbone), –2.38 (Si(CH₃)₂), –4.32 (SiCH₃). MALDI-TOF-MS: m/z 3335.45 (M – Cl)⁺. Calcd: m/z 3335.92 (M – Cl)⁺. Anal. Calcd for C₁₂₈H₂₂₀N₈Si₁₃Pd₈Cl₈: C 45.60, H 6.58, N 3.32, Cl 8.41. Found: C 45.46, H 6.47, N 3.24, Cl 8.52.

[G1]-(SiMe₂-CN-PdCl{pyr})₈ (12b). To a solution of **12a** (0.17 g, 0.050 mmol) in CH₂Cl₂ (10 mL) was added an excess of pyridine (0.5 mL). The mixture was stirred for 20 min and then evaporated to dryness. The solid residue was washed with pentane (20 mL) and dried in vacuo to yield a slightly yellow powder (>99%). ¹H NMR (CDCl₃): δ 8.88 (br, 16H, pyr-H_{ortho}), 7.78 (br, 8H, pyr-H), 7.32 (br, 16H, pyr-H), 7.07 (br, 8H, Ar–H), 6.95 (br, 8H, Ar–H), 6.04 (s, 8H, Ar–H_{ortho}), 3.94 (s, 16H, CH₂N), 2.92 (s, 48H, N(CH₃)₂), 1.80 (br m, 8H, dendrimer CH₂), 1.20 (br m, 16H, dendrimer CH₂), 0.82 (br m, 8H, dendrimer CH₂), 0.55 (br m, 16H, dendrimer CH₂), 0.44 (br m, 16H, dendrimer CH₂), 0.21 (br m, 8H, dendrimer CH₂), –0.04 (s, 48H, Si(CH₃)₂), –0.17 (s, 12H, SiCH₃). ¹³C{¹H} NMR (CDCl₃): δ 153.87, 148.85, 148.52, 137.81, 137.30, 136.82, 129.81, 125.13, 121.30 (9 × Ar–C), 74.14 (CH₂N), 53.02 (N(CH₃)₂), 20.64, 19.51, 18.91, 18.76, 18.63, 17.93 (6 × dendrimer CH₂), –2.69 (Si(CH₃)₂), –4.68 (SiCH₃). Anal. Calcd for C₁₆₈H₂₆₀N₁₆Si₁₃Pd₈Cl₈: C 50.39, H 6.55, N 5.60, Cl 7.08. Found: C 50.47, H 6.61, N 5.52, Cl 7.07.

[G1]-(SiMe₂-CN-PdCl)₁₂ (13a). This compound was prepared as described for **10**. To a stirred solution of $[G1]-(SiMe_2-CN)_{12}$ (0.66 g, 0.21 mmol) in MeOH was added Pd(OAc)₂ (0.65 g, 2.89 mmol). Gradually an orange-brown colored suspension was observed, which was stirred for 21 h. Then LiCl (0.53 g, 12.5 mmol) was added and this mixture stirred for ca. 2 h. After a similar workup as described for **10**, a light yellow to orange powder was isolated (0.40 g, 0.083 mmol, 40%). ¹H NMR (CDCl₃): δ 7.40 (br, 12H, Ar–H_{ortho}), 7.08 (br, 12H, Ar–H), 6.83 (br, 12H, Ar–H), 3.88 (br, 24H, CH₂N), 2.83 (br s, 72H, N(CH₃)₂), 1.29 (br m, –CH₂–, dendrimer backbone), 0.76 (br m, –CH₂–, dendrimer backbone), 0.53 (br m, –CH₂–, dendrimer backbone), 0.21 (br s, 72H, Si(CH₃)₂). ¹³C{¹H} NMR (CDCl₃): δ 147.47, 143.02, 138.82, 137.60, 129.99, 121.39 (6 × Ar–C), 73.45 (CH₂N), 52.91 (N(CH₃)₂), 21.22, 20.82, 18.91, 17.89 (–CH₂– dendrimer backbone, several signals coincide with each other), –2.33 (Si(CH₃)₂). MALDI-TOF-MS: m/z 4777.4 (M – Cl)⁺. Calcd: m/z 4777.02 (M – Cl)⁺. Anal. Calcd for C₁₈₀H₃₀₀N₁₂Si₁₃Pd₁₂Cl₁₂: C 44.92, H 6.28, N 3.49, Cl 8.84. Found: C 44.86, H 6.15, N 3.41, Cl 8.68.

[G1]-(SiMe₂-CN-PdCl{pyr})₁₂ (13b). To a solution of **13a** (0.16 g, 0.033 mmol) in CH₂Cl₂ (10 mL) was added an excess of pyridine (0.5 mL). The yellow-colored mixture was stirred for 5 min and then evaporated to dryness. The solid residue was washed with pentane (20 mL) and dried in vacuo to

yield a slightly yellow powder (0.16 g, 0.028 mmol, 84%). ^1H NMR (CDCl_3): δ 8.88 (br, 16H, pyr- H_{ortho}), 7.78 (br, 8H, pyr-H), 7.32 (br, 16H, pyr-H), 7.07 (br, 8H, Ar-H), 6.95 (br, 8H, Ar-H), 6.04 (s, 8H, Ar- H_{ortho}), 3.94 (s, 16H, CH_2N), 2.92 (s, 48H, $\text{N}(\text{CH}_3)_2$), 1.80 (br m, 8H, dendrimer CH_2), 1.20 (br m, 16H, dendrimer CH_2), 0.82 (br m, 8H, dendrimer CH_2), 0.55 (br m, 16H, dendrimer CH_2), 0.44 (br m, 16H, dendrimer CH_2), 0.21 (br m, 8H, dendrimer CH_2), -0.04 (s, 48H, $\text{Si}(\text{CH}_3)_2$), -0.17 (s, 12H, SiCH_3). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 153.83, 148.86, 148.59, 137.84, 137.25, 136.74, 129.77, 125.13, 121.31 (9 \times Ar-C), 74.10 (CH_2N), 53.01 ($\text{N}(\text{CH}_3)_2$), 20.82, 18.73, 17.93, 17.63 (4 \times dendrimer CH_2 , 2 signals are coincident), -2.66 ($\text{Si}(\text{CH}_3)_2$). Anal. Calcd for $\text{C}_{240}\text{H}_{360}\text{N}_{24}\text{Si}_{17}\text{Pd}_{12}\text{Cl}_{12}$: C 50.03, H 6.30, N 5.83, Cl 7.38. Found: C 49.92, H 6.21, N 5.77, Cl 7.49.

[G2]-(SiMe₂-CN)₃₆ (14). This compound was prepared using a literature procedure¹⁰ by treatment of [G2]-SiMe₂Cl (1.42 g, 0.235 mmol) with an excess of the in situ prepared lithium reagent [Li(C₆H₄{CH₂NMe₂}-4)] (ca. 16 mmol). After appropriate workup using Kugelrohr distillation (180–200 °C), the product was isolated as an extremely viscous brown oil (0.86 g, 0.090 mmol, 38%). ^1H NMR (C_6D_6): δ 7.54 (br, 72H, Ar-H), 7.44 (br, 72H, Ar-H), 3.30 (br s, 72H, CH_2N), 2.10 (br s, 216H, $\text{N}(\text{CH}_3)_2$), 1.59 (br m, - CH_2 -, dendrimer support), 0.86 (br m, - CH_2 -, dendrimer support), 0.32 (br s, 216H, $\text{Si}(\text{CH}_3)_2$). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 140.97, 138.36, 134.29, 129.13 (4 \times Ar-C), 64.96 (CH_2N), 45.87 ($\text{N}(\text{CH}_3)_2$), 21.50, 19.60, 18.85, 18.30 (4 \times - CH_2 -, dendrimer support with the other signals most probably coincident), -2.07 ($\text{Si}(\text{CH}_3)_2$). $^{29}\text{Si}\{^1\text{H}\}$ NMR (C_6D_6): δ 1.37, 1.08, 0.90 (inner *S*), -3.92 (outer *S*). MALDI-TOF-MS: *m/z* ca. 7200–9930 (br, M^+ region), ca. 4250–6250 (br, M^{2+} region), ca. 2000–3500 (br, $\text{M}^{3/4+}$ region). Anal. Calcd for $\text{C}_{552}\text{H}_{960}\text{N}_{36}\text{Si}_{53}$: C 69.13, H 10.09, N 5.26, Si 15.52. Found: C 69.07, H 10.02, N 5.15, Si 15.43.

[G2]-(SiMe₂-CN-PdCl)₃₆ (15a). To a solution of **14** (0.30 g, 0.0313 mmol) in MeOH (40 mL) was added Pd(OAc)₂ (0.25 g, 1.13 mmol). The resultant yellow-brown suspension was stirred for 16 h. The solvent was then decanted and the solid residue washed with one portion of MeOH (60 mL). The chloride-bridged intermediate species proved to be insoluble in the common organic solvents. Therefore, compound **15a** was converted into its soluble pyridine adduct in situ without further purification.

[G2]-(SiMe₂-CN-PdCl{pyr})₃₆ (15b). To an unpurified sample of **15a** (vide supra) was added an excess of pyridine (ca. 5 mL), and immediately solubilization of the solid material occurred. The dark solution was stirred for 16 h and then filtered through a path of Celite to remove traces of palladium black and then concentrated to a volume of ca. 10 mL. The product was precipitated by adding pentane (50 mL), collected by centrifugation, and washed again with pentane (50 mL) and acetone (50 mL). Drying in vacuo gave a white solid (0.21 g, 0.0120 mmol, 38% overall). ^1H NMR (CDCl_3 , mixture of isomers): δ 8.83 (br m, 72H, pyr- H_{ortho}), 7.70 (br m, 36H, pyr-H), 7.40 (br m, 72H, pyr-H), 7.05 (br m, 72H, Ar-H), 6.90–6.86 (br m, 72H, Ar-H), 6.01 (br s, 36H, Ar- H_{ortho}), 3.93 and 3.88 (2 \times br s, 72H, CH_2N), 2.91 and 2.81 (2 \times br s, 216H, $\text{N}(\text{CH}_3)_2$), 1.27, 0.84, 0.82, 0.53 (br m, all dendrimer CH_2), 0.16 and -0.07 (2 \times s, 216H, $\text{Si}(\text{CH}_3)_2$). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , mixture of isomers): δ 153.81, 153.53, 148.86, 148.74, 148.55, 148.43, 138.75, 138.40, 137.94, 137.27, 136.74, 129.87, 125.21, 121.42 (all Ar-C + pyr-C), 74.15, 73.44 (CH_2N), 53.05 ($\text{N}(\text{CH}_3)_2$), 20.85, 18.08, 17.66, 16.14 (partly overlapping dendrimer CH_2), -2.20, -2.54 ($\text{Si}(\text{CH}_3)_2$). Anal. Calcd for $\text{C}_{732}\text{H}_{1104}\text{N}_{72}\text{Si}_{53}\text{Pd}_{36}\text{Cl}_{36}$: C 50.21, H 6.36, N 5.76, Cl 7.29, Pd 21.88. Found: C 50.18, H 6.28, N 5.71, Cl 7.34, Pd 21.76.

Preparation of Aldol Condensation Catalysts. General procedure: A mixture of [PdCl(C₆H₃{CH₂NMe₂}-2-(SiMe₃)-4)-(Pyr)]¹¹ (60.0 mg, 0.14 mmol) and a slight excess of AgBF₄ (32.3 mg, 0.17 mmol) in acetone (25 mL) was stirred for 16 h at RT. Then the precipitated silver salt was thoroughly removed by

filtration through a path of Celite to give a clear, yellow-colored filtrate. The solvent was removed in vacuo and the product precipitated from an acetone/pentane mixture to yield [Pd-(C₆H₃{CH₂NMe₂}-2-(SiMe₃)-4)(H₂O)(Pyr)] (**16**), as a solid material (75.1 mg, 0.15 mmol). The dendrimer species **5c**, **7c**, **10c**, **12c**, **13c**, and **15c** were prepared analogously. It should be mentioned that the dendrimer precursor compounds **5b**, **7b**, **10b**, **12b**, **13b**, and **15b** are rather insoluble in acetone, but dissolve upon treatment with AgBF₄. All cationic species are extremely hygroscopic and were used immediately after isolation and ^1H spectroscopic analysis in the catalytic experiments. For **16** (off-white solid, yield >99%): ^1H (CDCl_3): δ 8.84 (d, J = 5.2 Hz, 2H, pyr- H_{ortho}), 7.92 (t, J = 7.8 Hz, 1H, pyr-H), 7.48 (t, J = 7.2 Hz, 2H, pyr-H), 7.15 (d, J = 7.0 Hz, 1H, Ar-H), 6.93 (d, J = 7.4 Hz, 1H, Ar-H), 5.99 (s, 1H, Ar- H_{ortho}), 3.94 (s, 2H, CH_2N), 3.13 (br s, 2H, H₂O), 2.84 (s, 6H, $\text{N}(\text{CH}_3)_2$), 0.02 (s, 9H, $\text{Si}(\text{CH}_3)_3$).

For **5c** (off-white solid, yield >99%): ^1H (acetone-*d*₆): δ 8.95 (d, J = 5.2 Hz, 8H, pyr- H_{ortho}), 8.15 (t, J = 8.0 Hz, 4H, pyr-H), 7.73 (t, J = 7.6 Hz, 8H, pyr-H), 7.14 (d, J = 6.6 Hz, 4H, Ar-H), 6.97 (d, J = 7.8 Hz, 4H, Ar-H), 6.03 (s, 4H, Ar- H_{ortho}), 4.11 (s, 8H, CH_2N), 2.96 (s, 24H, $\text{N}(\text{CH}_3)_2$), 1.45, 0.75, 0.58 (3 \times m, 8H, dendrimer CH_2), -0.04 (s, 24H, $\text{Si}(\text{CH}_3)_2$).

For **7c** (light yellow solid, yield >99%): ^1H (acetone-*d*₆): δ 8.90 (br, 8H, pyr- H_{ortho}), 8.04 (br, 4H, pyr-H), 7.64 (br, 8H, pyr-H), 7.14 (d, J = 7.2 Hz, 4H, Ar-H), 7.04 (d, J = 6.8 Hz, 4H, Ar-H), 6.06 (s, 4H, Ar- H_{ortho}), 4.14 (s, 8H, CH_2N), 3.42 (br s, 8H, H₂O), 2.82 (s, 24H, $\text{N}(\text{CH}_3)_2$), 0.27 (br m, 16H, dendrimer CH_2), -0.01 (s, 24H, $\text{Si}(\text{CH}_3)_2$).

For **10c** (light yellow solid, yield 56%): ^1H (acetone-*d*₆): δ 8.90 (br, 16H, pyr- H_{ortho}), 8.08 (br, 8H, pyr-H), 7.64 (br, 16H, pyr-H), 7.12 (d, J not resolved, 8H, Ar-H), 7.01 (d, J = 7.2 Hz, 8H, Ar-H), 6.06 (s, 8H, Ar- H_{ortho}), 4.11 (s, 16H, CH_2N), 3.36 (br s, 16H, H₂O), 2.83 (s, 48H, $\text{N}(\text{CH}_3)_2$), 0.38 (br m, dendrimer CH_2), 0.18 (br m, dendrimer CH_2), -0.04 (s, 48H, $\text{Si}(\text{CH}_3)_2$), -0.14 (s, 12H, SiCH_3).

For **12c** (off-white solid, yield 91%): ^1H (acetone-*d*₆): δ 8.93 (d, J = 4.8 Hz, 16H, pyr- H_{ortho}), 8.12 (t, 8H, pyr- H_{ortho}), 7.69 (t, J = 8 Hz, 16H, pyr-H), 7.16 (d, J = 7 Hz, 8H, pyr-H), 7.14 (d, J = 7 Hz, 8H, Ar-H), 7.02 (d, J = 7 Hz, 8H, Ar-H), 6.03 (s, 8H, Ar- H_{ortho}), 4.11 (s, 16H, CH_2N), 3.22 (br s, 16H, H₂O), 2.84 (s, 48H, $\text{N}(\text{CH}_3)_2$), 1.30 (br m, dendrimer CH_2), 0.65–0.55 (br m, dendrimer CH_2), -0.04 (s, 48H, $\text{Si}(\text{CH}_3)_2$), -0.11 (s, 12H, SiCH_3).

For **13c** (yellow to brown solid, yield 80%): ^1H (acetone-*d*₆): δ 8.92 (br, 24H, pyr- H_{ortho}), 8.12 (br, 12H, pyr-H), 7.64 (br, 24H, pyr-H), 7.14 (br, 12H, Ar-H), 7.02 (br, 12H, Ar-H), 6.06 (br s, 12H, Ar- H_{ortho}), 4.11 (br s, 24H, CH_2N), 3.28 (br s, 24H, H₂O), 2.84 (br s, 72H, $\text{N}(\text{CH}_3)_2$), 1.42–1.23 (br m, dendrimer CH_2), 0.85 (br m, dendrimer CH_2), 0.62 (br m, dendrimer CH_2), 0.27 (br m, dendrimer CH_2), 0.22 (br m, dendrimer CH_2), -0.01 (br s, 72H, $\text{Si}(\text{CH}_3)_2$).

For **15c** (light brown, sticky solid, yield 78%): ^1H (acetone-*d*₆): δ 8.88 (br, 72H, pyr- H_{ortho}), 8.06 (br, 36H, pyr-H), 7.63 (br, 72H, pyr-H), 7.12 (br, 36H, Ar-H), 6.98 (br, 36H, Ar-H), 6.06 (s, 36H, Ar- H_{ortho}), 4.09 (br s, 144H, CH_2N), 3.16 (br, H₂O), 2.81 (br s, 432H, $\text{N}(\text{CH}_3)_2$), 1.40–0.13 (br, dendrimer CH_2), -0.02 (br s, 216H, $\text{Si}(\text{CH}_3)_2$).

Protocol for the Catalytic Aldol Condensation Reaction. A mixture of benzaldehyde (0.26 g, 2.4 mmol), methyl isocyanoacetate (0.16 g, 1.6 mmol), ethyl-*N,N*-diisopropylamine (cocatalyst, 0.03 g), mesitylene (internal standard, 0.19 g, 1.6 mmol), and CH_2Cl_2 (10 mL) was prepared and added to 1 mol % of catalyst species (vide supra). From the clear and completely homogeneous, stirred reaction mixture samples were taken after regular time intervals and analyzed with ^1H spectroscopy in CDCl_3 after careful removal of the solvent. ^1H NMR (CDCl_3 , oxazoline product): δ 7.38–7.23 (m, Ar-H), 7.10 (d, 4J = 2.2 Hz, (*H*)C=N, *cis* or *trans* isomer, other signal not visible due to overlap), 5.73 (d, 3J = 7.4 Hz, Ar-C(*H*)-O, *cis* isomer), 5.68 (d, 3J = 5.2 Hz, Ar-C(*H*)-O, *trans* isomer), 5.08

(d, $^3J = 11.0$ Hz, $^4J = 1.8$ Hz, $-C(H)CO_2CH_3$, *cis* isomer), 4.62 (d, $^3J = 7.8$ Hz, $^4J = 2.2$ Hz, $-C(H)CO_2CH_3$, *trans* isomer), 3.83 (s, $-C(H)CO_2CH_3$, *trans* isomer), 3.19 (s, $-C(H)CO_2CH_3$, *cis* isomer).

Crystal Structure Determination for 4 and 7a. For **4**: $C_{30}H_{52}Br_{0.15}Cl_{1.85}N_2Pd_2Si_3 \cdot 1/2C_6H_6$, $M_r = 854.48$, yellow, needle-shaped crystals ($0.1 \times 0.2 \times 0.9$ mm), cut from a larger aggregate, monoclinic, space group $P2_1/c$ (no. 14) with $a = 12.746(9)$, $b = 14.627(4)$, $c = 23.890(16)$ Å, $\beta = 109.50(5)^\circ$, $V = 4198(4)$, $Z = 4$, $D_x = 1.352$ g/cm 3 , $F(000) = 1751$, $\mu(Mo K\alpha) = 1.23$ mm $^{-1}$. All data, where relevant, are given without disordered solvent contribution (vide infra); 11 405 reflections measured, 9601 independent reflections, $R_{int} = 0.0366$ ($1.66^\circ < \theta < 27.47^\circ$, ω scan, $T = 150$ K, Mo K α radiation, graphite monochromator, $\lambda = 0.71073$ Å) on an Enraf-Nonius CAD4 Turbo diffractometer on a rotating anode. The structure was solved by direct methods (SHELXS86)²⁵ and refined on F^2 using SHELXL-96.²⁶ One of the bridging Cl atoms is substitutionally disordered with Br. A disorder model was introduced; the occupancy of the Cl component was refined to 0.850(3). Besides the disordered benzene solvate area, the unit cell contains two symmetry-related cavities (220 Å 3 each), filled with disordered solvent. No satisfactory model could be refined; the disordered density was therefore taken into account with the PLATON/SQUEEZE procedure.²⁷ A total electron count of 74 per unit cell was found and corrected for. Hydrogen atoms were included in the refinement on calculated positions riding on their carrier atoms. All non-hydrogen atoms were refined with anisotropic displacement parameters. The displacement parameters of the Br atom were set equal to those of the corresponding Cl atom. Hydrogen atoms were refined with a fixed isotropic displacement parameter related to the value of the equivalent isotropic displacement parameter of their carrier atoms. Refinement of 393 parameters resulted in a final wR2 value of 0.091, $w = 1/[\sigma^2(F^2) + (0.0440P)^2]$, where $P = (\max(F_o^2, 0) + 2F_c^2)/3$, $R1 = 0.0393$ (for 7266 $I > 2\sigma(I)$), $S = 0.998$, $-0.61 < \Delta\rho < 0.55$ e/Å 3 .

(25) Sheldrick, G. M. *SHELXS86*, program for crystal structure solution; University of Göttingen: Germany, 1986.

(26) Sheldrick, G. M. *SHELXL96*, program for crystal structure refinement; University of Göttingen: Germany, 1996.

(27) Spek, A. L. *PLATON*, a multipurpose crystallographic tool; Utrecht University: The Netherlands, 2000.

For **7a**: $C_{52}H_{84}C_{14}N_4Pd_4Si_5 \cdot 0.4CH_2Cl_2$, fw = 1507.05, yellow plate, $0.38 \times 0.31 \times 0.06$ mm 3 , orthorhombic, *Pbca* (no. 61), $a = 13.3571(2)$, $b = 26.9874(4)$, $c = 39.6304(5)$ Å, $V = 14285.7(4)$ Å 3 , $Z = 8$, $D_x = 1.401$ g/cm 3 , $\mu = 1.29$ mm $^{-1}$; 103 609 reflections were measured on a Nonius KappaCCD diffractometer with a rotating anode ($\lambda = 0.71073$ Å) at a temperature of 150(2) K up to a resolution of $(\sin \theta/\lambda)_{max} = 0.59$ Å $^{-1}$. The data were merged using the program SORTAV;²⁸ 12 559 reflections were unique ($R_{int} = 0.063$). An absorption correction was not considered necessary. The structure was solved with direct methods (SHELXS97)²⁹ and refined with SHELXL97³⁰ against F^2 of all reflections. Non-hydrogen atoms were refined freely with anisotropic displacement parameters. Hydrogen atoms were refined as rigid groups. The dichloromethane molecules were refined on two positions with a total occupancy of 0.4 with respect to the main molecule: 666 refined parameters, 109 restraints. R values [$I > 2\sigma(I)$]: $R1 = 0.0506$, $wR2 = 0.1659$. R values [all reflns]: $R1 = 0.0762$, $wR2 = 0.1862$. GoF = 0.995. The remaining electron density was between -0.90 and 1.73 e/Å 3 . Molecular illustration, structure checking, and calculations were performed with the PLATON package.²⁷

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Supporting Information Available: Crystallographic details of the structure determination of **4** and **7a** including thermal parameters. This information is available free of charge via the Internet at <http://pubs.acs.org>.

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(28) Blessing, R. H. *J. Appl. Crystallogr.* **1997**, *30*, 421.

(29) Sheldrick, G. M. *SHELXS97*, program for crystal structure solution; University of Göttingen: Germany, 1997.

(30) Sheldrick, G. M. *SHELXL97*, program for crystal structure refinement; University of Göttingen: Germany, 1997.