

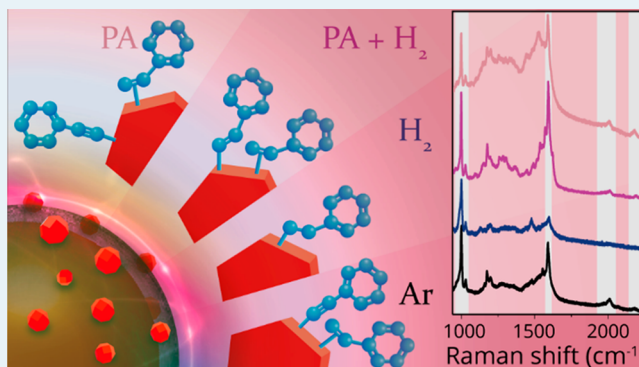
In Situ Shell-Isolated Nanoparticle-Enhanced Raman Spectroscopy to Unravel Sequential Hydrogenation of Phenylacetylene over Platinum Nanoparticles

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Supporting Information

ABSTRACT: Shell-isolated nanoparticle-enhanced Raman spectroscopy (SHINERS) is quickly developing into a powerful characterization tool in heterogeneous catalysis. In this work, we employ Pt catalysts supported on Au@SiO₂ shell-isolated nanoparticles to study hydrogenation reactions. First, we demonstrate the facile preparation of Pt/Au@SiO₂ and its characterization by using adsorption of CO as probe molecule. Next, we use the adsorption and hydrogenation of phenylacetylene as a model reaction for the interaction of triple bonds and aromatic rings with catalytic Pt surfaces. We show that the applicability of SHINERS is not limited to inherently gaseous compounds, thereby expanding the applicability of the technique to more complex systems. Furthermore, by using nonparticipating side groups as labels, we observe the sequential hydrogenation of phenylacetylene into styrene and ultimately ethylbenzene upon reaction with H₂. Upon the absence of H₂, the reverse reaction takes place with those molecules still adsorbed onto the catalyst surface, which allowed a more detailed understanding of the reaction mechanism and the assignment of Raman peaks. This strengthens the position of SHINERS as an easily applicable surface sensitive technique that can be used to study a wide variety of chemical reactions in the field of heterogeneous catalysis.



KEYWORDS: Raman spectroscopy, heterogeneous catalysis, Au nanoparticles, in situ SHINERS, hydrogenation

INTRODUCTION

Understanding metal–adsorbate interactions is of fundamental importance for rational catalyst design.¹ Raman spectroscopy offers the ability to study such interactions, making it a powerful tool in heterogeneous catalysis research by offering fundamental insights into structure–performance relationships. These insights could help in the design of more optimal solid catalysts for both new and already existing chemical processes, increasing conversion, activity, and selectivity and reducing waste and greenhouse gases.^{2,3}

However, the applicability of Raman spectroscopy to nanoscale systems is limited due to the poor detection limit, typically around 1–10 wt % of the probed area.⁴ Surface-enhanced Raman spectroscopy (SERS) solves that problem by employing gold and silver nanoparticles (NPs).^{4–6} Au and Ag NPs of sufficient size (starting at >10 nm; optimum around 100 nm)⁷ possess surface plasmons which, when excited with light of the resonant frequency, can create an electromagnetic field that enhances the Raman scattering of species on or within several nm of the surface. Using this principle, in ideal cases, SERS has been reported to enhance Raman signals with factors of up to 10¹⁴, corresponding to even single molecule

levels.^{8–11} For catalysis this means that SERS has the potential to not only elucidate structure–performance relationships through studying adsorption of species of interest in metal surfaces but also through the detection of short-lived intermediates or even transition states.¹² However, the surface plasmons and their resonance SERS is based on can influence chemical reactions to be studied by injecting “hot electrons”, causing photocatalytic side-reactions.^{4,13} Furthermore, low stability of the Ag and Au NPs limits reaction conditions to be studied. These challenges can be overcome by using shell-isolated nanoparticle-enhanced Raman spectroscopy (SHINERS).¹⁴ For SHINERS, the SERS-active NPs are coated with a thin layer of dielectric oxide, for example SiO₂. This thin layer not only increases the thermal and chemical stability of the SERS-active core, several studies have showed that the SiO₂ layer also confines hot electrons to the core.^{13,15} They demonstrated this by studying the reduction of 4-nitrothiophenol (4-NTP) to 4-aminothiophenol (4-ATP) over a

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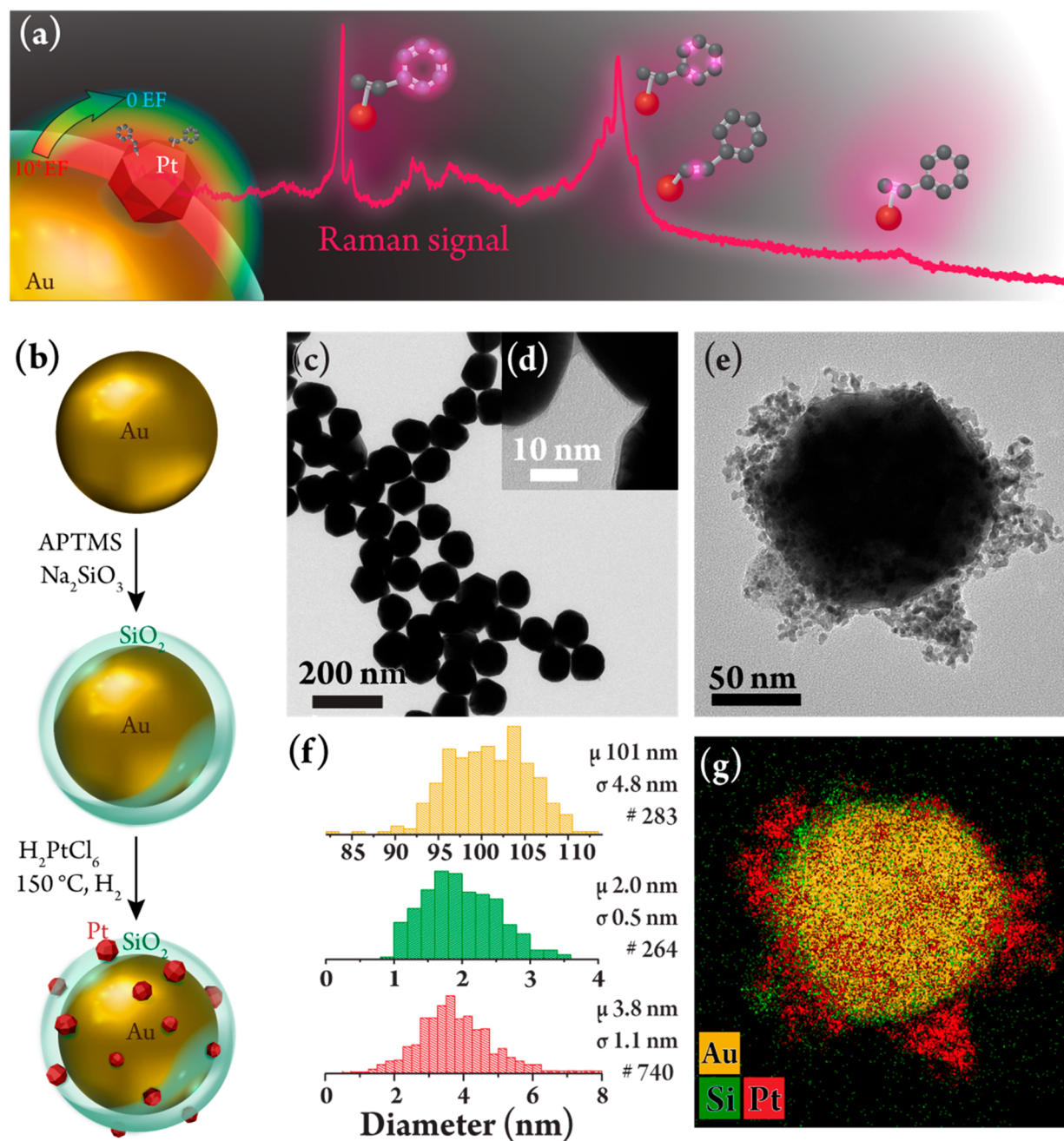


Figure 1. (a) Schematic illustrating the goal of this study, which is to observe the conversion of phenylacetylene to styrene and ethylbenzene over a Pt hydrogenation catalyst placed inside the enhanced electromagnetic field of a SERS-active Au@SiO₂ shell-isolated nanoparticle (SHIN). (b) Schematic visualization of the Pt/Au@SiO₂ catalyst/SHIN preparation. (c) Transmission electron microscopy (TEM) images of the Au core nanoparticles, (d) Au@SiO₂ SHINs, and (e) of a Pt/Au@SiO₂ catalyst/SHIN. (f) Size distributions based on TEM measurements. The Au core (top, yellow) NPs have a mean diameter of 101 nm and are coated with a 2 nm layer of SiO₂ (middle, green). Pt NPs (bottom, red) are ~3.8 nm in diameter. (g) Elemental map of (e) confirming the presence of Pt (red) on top of the Au@SiO₂ (yellow, green, respectively) SHINs. Note that the Pt/Au@SiO₂ TEM samples were prepared after *in situ* SHINERS experiments by redispersion from the Si wafer and transfer to a TEM grid. Therefore, the aggregation state and dispersion of Pt NPs over the Au@SiO₂ SHINs is not a representative image of those during *in situ* experiments.

Au catalyst. Hot electrons are known to influence this reaction by catalyzing dimerization to 4,4'-dimercaptoazobenzene (DMAB).^{16–18} Using a catalyst comprised of SERS-inactive colloidal Au NPs of <5 nm supported on Au@SiO₂ shell-isolated nanoparticles (SHINs) prevented the formation of DMAB, thus proving that hot electrons could not participate in the reaction. A similar principle where colloidal nanoparticle catalysts were deposited on SHINs was used by the group of

Tian to study hydrogenation of 4-NTP¹⁹ and CO adsorption and oxidation²⁰ on transition metal/Au@SiO₂ SHINs. Furthermore, our group has recently reported that catalyst/SHIN systems can also be prepared by impregnation of SHINs with a catalyst precursor followed by *in situ* reduction.^{21,22} Subsequently, support effects were studied by investigating CO hydrogenation over Ru and Rh catalysts supported on Au@SiO₂ and Au@TiO₂ SHINs.

In this study, we will extend the applicability of SHINERS to reactions beyond those involving 4-NTP and simple gases, such as CO. For the first time, we report an *in situ* SHINERS study of sequential hydrogenation reactions of organic molecules, more specifically phenylacetylene to styrene and finally ethylbenzene, on the surface of a Pt/Au@SiO₂ catalyst, as shown in Figure 1a.

Normally, this chemical reaction is carried out in the liquid phase, at slightly elevated temperatures and pressures.²³ Furthermore, often a Lindlar catalyst composed of Pd poisoned with S and Pb is used to selectively produce alkenes and prevent further hydrogenation to alkanes.²⁴ However, we will prepare Au@SiO₂ SHINs and deposit catalytically active Pt catalysts on the surface by wet impregnation and *in situ* reduction as reported by Hartman et al.²¹ We will show that SHINERS is a suitable technique to study hydrogenation reactions of alkynes and alkenes and that aromatic side groups can be used as spectroscopic tags to monitor the reaction. Furthermore, by using phenylacetylene vapor saturated in N₂, we show that the application of SHINERS to heterogeneous catalysis is not limited to intrinsically gaseous compounds.

■ EXPERIMENTAL SECTION

Preparation of Au@SiO₂ SHINs. The preparation of Au@SiO₂ SHINs is described in detail elsewhere.²⁵ In short, gold seeds were synthesized using the Turkevich method²⁶ to obtain Au seeds of approximately 17 nm in Millipore water (MQ; resistivity 18.2 MΩ·cm) (HAuCl₄ × 3H₂O 99.99% obtained from Alfa Aesar; sodium tricitrate dihydrate, 99%, Sigma-Aldrich). NPs were grown to a size of approximately 100 nm using NH₂OH × HCl (99.995%, Sigma-Aldrich) as a mild reducing agent.²⁷ Subsequent SiO₂ growth was realized through the APTMS (99%, Sigma-Aldrich)/sodium silicate (27%, Sigma-Aldrich) method to yield Au@SiO₂ SHINs.²⁵ Successful synthesis of these particles was confirmed as per the characterization data presented in Figure 1c, d, and f and the Supporting Information Figures S1–3.

Preparation of Pt/Au@SiO₂ SHINs. A 50 μL portion of a 10 mM H₂PtCl₆ (99.995%, Sigma-Aldrich) solution was mixed with 0.5 mL Au@SiO₂ SHINs. After mixing and equilibrating for ~30 min, 15 μL of the mixture was dropcast on an Si wafer and was vacuum-dried. The sample was cleaned in an Ossila E511 UV-Ozone Cleaner for 15 min before *in situ* SHINERS experiments. The sample was placed in a Linkam Cell and Pt was reduced *in situ* at 150 °C under 50% H₂ flow in Ar for 1 h, while monitoring the process with SHINERS on a Renishaw inVia confocal Raman microscope using a 20× objective (NA = 0.4), 1200 l/mm grating, 0.12–1.96 mW laser intensity with 785 nm laser excitation.

Characterization. Au and Au@SiO₂ NP size and LSPR were investigated with UV–vis absorption spectroscopy on a Varian Cary 50 UV–visible spectrophotometer in quartz cuvettes with an optical path length of 10.00 mm. Au and Au@SiO₂ NP size and layer thickness were investigated on a FEI Tecnai 12 and FEI Tecnai 20 transmission electron microscope (TEM), operating at 120 and 200 keV, respectively. Samples were prepared by drying 15 μL of the dispersions onto Formvar/carbon-coated 300 mesh copper TEM grids (Van Loenen Instruments). The Pt/Au@SiO₂ catalyst/SHINs were investigated with elemental analysis using high-angle-annular dark field scanning transmission electron microscopy (HAADF-STEM) and energy dispersive X-ray spectroscopy (EDX) on a FEI Talos F200 electron microscope operating at

200 keV. Samples were prepared by redispersing *in situ* reduced samples in isopropanol and subsequent transfer to Formvar/carbon-coated 300 mesh copper TEM grids.

Raman Spectroscopy of Reference Compounds.

Raman spectra of pure phenylacetylene, styrene, and ethylbenzene were measured in a quartz cuvette with an optical path length of 10.00 mm using an Avantes Raman probe in combination with a 532 nm Cobalt laser set to 51 mW connected to an Andor EMCCD detector equipped with a 1200 l/mm grating.

CO Adsorption Experiments. After *in situ* reduction, 10 mL/min CO and 40 mL/min Ar were introduced into the Linkam Cell at 150 °C. Adsorption of CO was monitored with *in situ* SHINERS on a Renishaw inVia confocal Raman microscope using a 20× objective (NA = 0.4), 1200 l/mm grating, 0.12–1.96 mW laser intensity with 785 nm laser excitation.

Phenylacetylene Adsorption Experiments. N₂ saturated with phenylacetylene was introduced into the Linkam Cell. Adsorption was studied at 25, 40, 60, and 80 °C for 30 min each and was monitored with *in situ* SHINERS on a Renishaw inVia confocal Raman microscope using a 20× objective (NA = 0.4), 1200 l/mm grating, 0.12–1.96 mW laser intensity with 785 nm laser excitation.

Phenylacetylene Hydrogenation Experiments. Phenylacetylene hydrogenation was studied by flowing H₂ over the catalyst at 60 °C. The feed consisted of a 1:1 mixture of H₂ with either N₂ saturated with phenylacetylene, or with Ar. Dehydrogenation conditions were achieved by flowing only Ar. The total flow was always 50 mL/min. Hydrogenation was monitored using *in situ* SHINERS on a Renishaw inVia confocal Raman microscope using a 20× objective, 1200 l/mm grating, 0.12–1.96 mW laser intensity with 785 nm laser excitation.

■ RESULTS

Characterization of Pt/Au@SiO₂ Catalyst/SHINs. Figure 1b shows a schematic representation of the preparation of Pt/Au@SiO₂. The transmission electron microscopy (TEM) images of the corresponding Au NPs at the top, the Au@SiO₂ shell-isolated nanoparticles (SHINs) in the middle, and the Pt/Au@SiO₂ catalyst/SHINs at the bottom are shown in Figure 1c–e, respectively, along with the accompanying particle size distributions in Figure 1f. The 101 nm diameter of the Au core lies close to the reported optimum for enhanced Raman scattering,⁷ whereas the ultrathin SiO₂ coating of 2.0 ± 0.5 nm is a good trade-off between minimizing the loss in signal intensity and obtaining a pinhole-free coating.^{19,25,28} Small nanostructures can be observed covering the Au@SiO₂ SHINs, which were confirmed to consist of Pt by the energy dispersive X-ray spectroscopy (EDX) map in Figure 1g. Particle size analysis of 740 Pt NPs gives an average size of 3.8 nm and a standard deviation of 1.1 nm. Taking into account the thickness of the SiO₂ layer, this means that on average the outer surface of the Pt NPs is about 5.5 nm away from the Au core. As the SERS effect is a very local effect with most intense enhancement factors arising within ~5 nm of the Au NP surface,⁴ this potentially means that we cannot observe processes occurring on the entire Pt NP surface. In order to test this, we performed CO adsorption experiments on the Pt/Au@SiO₂ samples. First, samples were cleaned with UV/ozone and then reduced *in situ* in a Linkam Cell at 150 °C for 1 h in a 1:1 mixture of H₂ and Ar. The Raman spectra were recorded

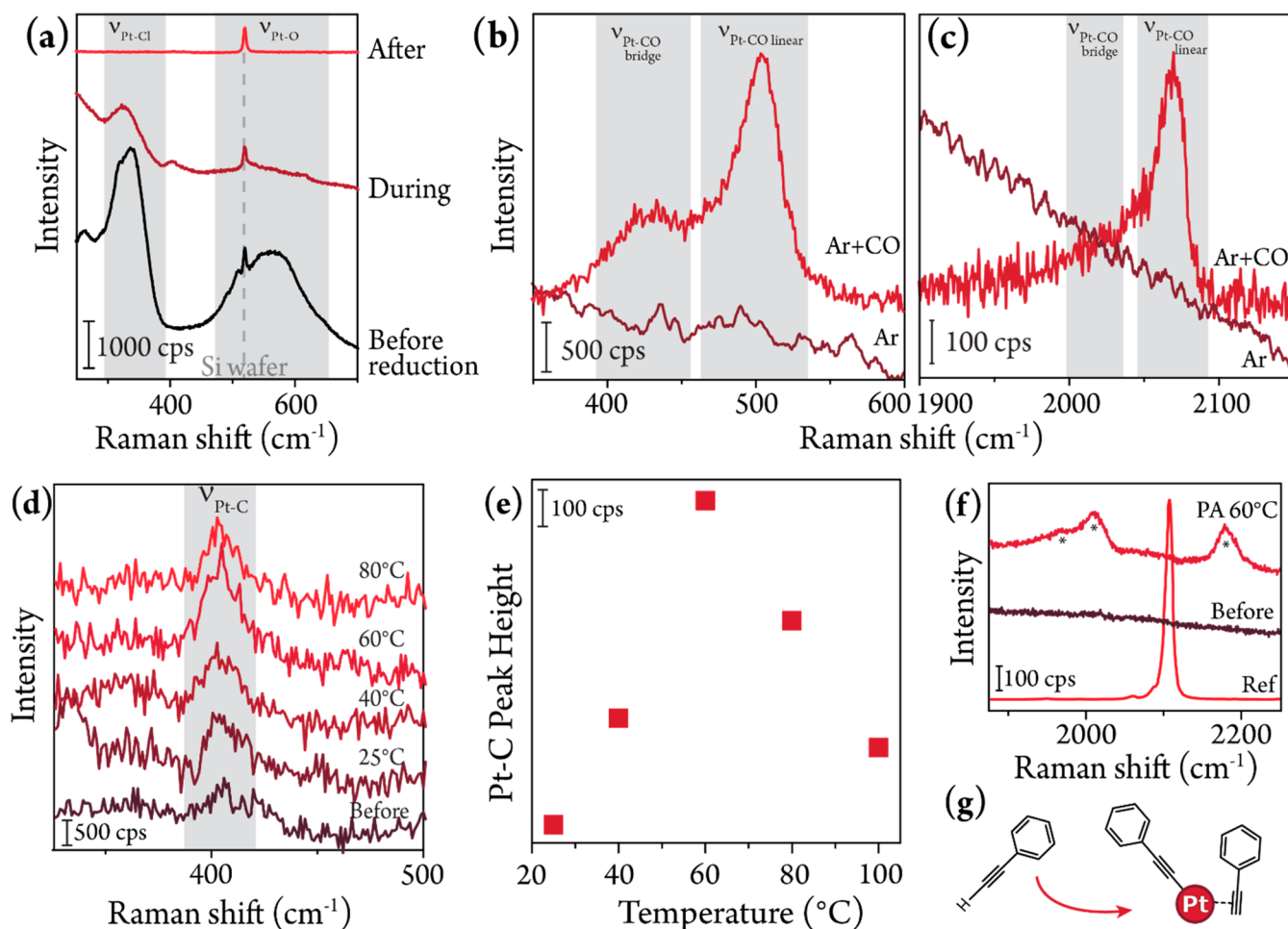


Figure 2. (a) *In situ* SHINER spectra of reduction of Pt_2Cl_6 (330 cm^{-1}) and PtO_2 (560 cm^{-1}) to metallic Pt. Note the sharp peak at 520 cm^{-1} originating from the Si wafer used to support the Pt/Au@SiO₂ catalyst/SHINs. (b, c) *In situ* SHINER spectra of CO adsorption at $150\text{ }^\circ\text{C}$. Both peaks assigned to CO in a bridge (b, 430 cm^{-1} ; c, 2010 cm^{-1}) and linear (b, 505 cm^{-1} ; c, 2070 cm^{-1}) configuration on Pt were observed. (d–g) Phenylacetylene adsorption on activated Pt/Au@SiO₂ catalyst/SHINs. (d) Metal–adsorbate stretching region where a peak assigned to Pt–C stretching can be observed at different temperatures. (e) Peak height for the Pt–C stretching vibration showing an optimum at $60\text{ }^\circ\text{C}$. (f) $\text{C}\equiv\text{C}$ stretching region. The reference spectrum of phenylacetylene exhibits an intense peak originating from free $\text{C}\equiv\text{C}-\text{H}$ at 2105 cm^{-1} , whereas phenylacetylene chemisorbed on Pt at $60\text{ }^\circ\text{C}$ shows three peaks at 1960 , 2010 , and 2190 cm^{-1} , indicated with asterisks. (g) Schematic of phenylacetylene chemisorption on Pt NPs through σ - and π -interactions.

using a 785 nm excitation source, which is known to give limited potential plasmonic side reactions in combination with Au nanoparticles (NPs) under the applied conditions.^{12,22}

Figure 2a shows the SHINER spectra during reduction. At the start of the experiment, two broad bands can be observed, namely one around 330 cm^{-1} for Pt–Cl stretching vibrations from the platinum precursor, and one around 560 cm^{-1} , which does not originate from Pt–Cl vibrations.²⁹ During UV/ozone treatment, the platinum chloride precursor becomes partially oxidized, resulting in the appearance of Pt–O stretching vibrations of amorphous platinum oxide.^{30–32} Upon introducing reducing conditions, the bands start to decrease in intensity and after an hour they are completely gone, indicating the formation of (Raman-inactive) metallic Pt.³³ Subsequently, the *in situ* cell was flushed with Ar and CO was introduced at $150\text{ }^\circ\text{C}$. Figure 2b and c shows the low and high wavenumber region of the SHINER spectra obtained during *in situ* CO adsorption. The low wavenumber region is known to show bands arising from metal–adsorbate stretching vibrations. Indeed, when we compare our CO adsorption spectra to those in the literature, we can see that we observe similar peaks

corresponding to CO adsorption on a Pt surface.^{20,21,34} In particular, we observe two different vibrations, one due to CO adsorbed on Pt in a bridge configuration at 430 cm^{-1} and one corresponding to CO adsorbed in a linear fashion at 505 cm^{-1} . Characteristic peaks for these two configurations can also be found in the high wavenumber region, looking at the C–O stretching vibration. Again, we observe a vibration associated with bridged CO as a shoulder at 2010 cm^{-1} , while at 2070 cm^{-1} we can see C–O stretching of CO adsorbed in a linear configuration. These experiments show that (1) the Pt NPs are within the range for signal enhancement from the SERS effect; (2) we can observe interactions of adsorbates with multiple sites of the Pt catalyst; and (3) we have sufficient accessible Pt surface that can be used to study heterogeneous catalytic reactions. Finally, as no shift in the stretching vibrations or bands corresponding to CO on Au were observed, we can state that our SiO₂ layer is pinhole-free and our SHINs are stable (for additional tests, see Figure S3–S5).^{14,35}

Phenylacetylene Adsorption. Adsorption of phenylacetylene on Pt/Au@SiO₂ catalyst/SHINs was investigated at various temperatures to determine the optimum conditions

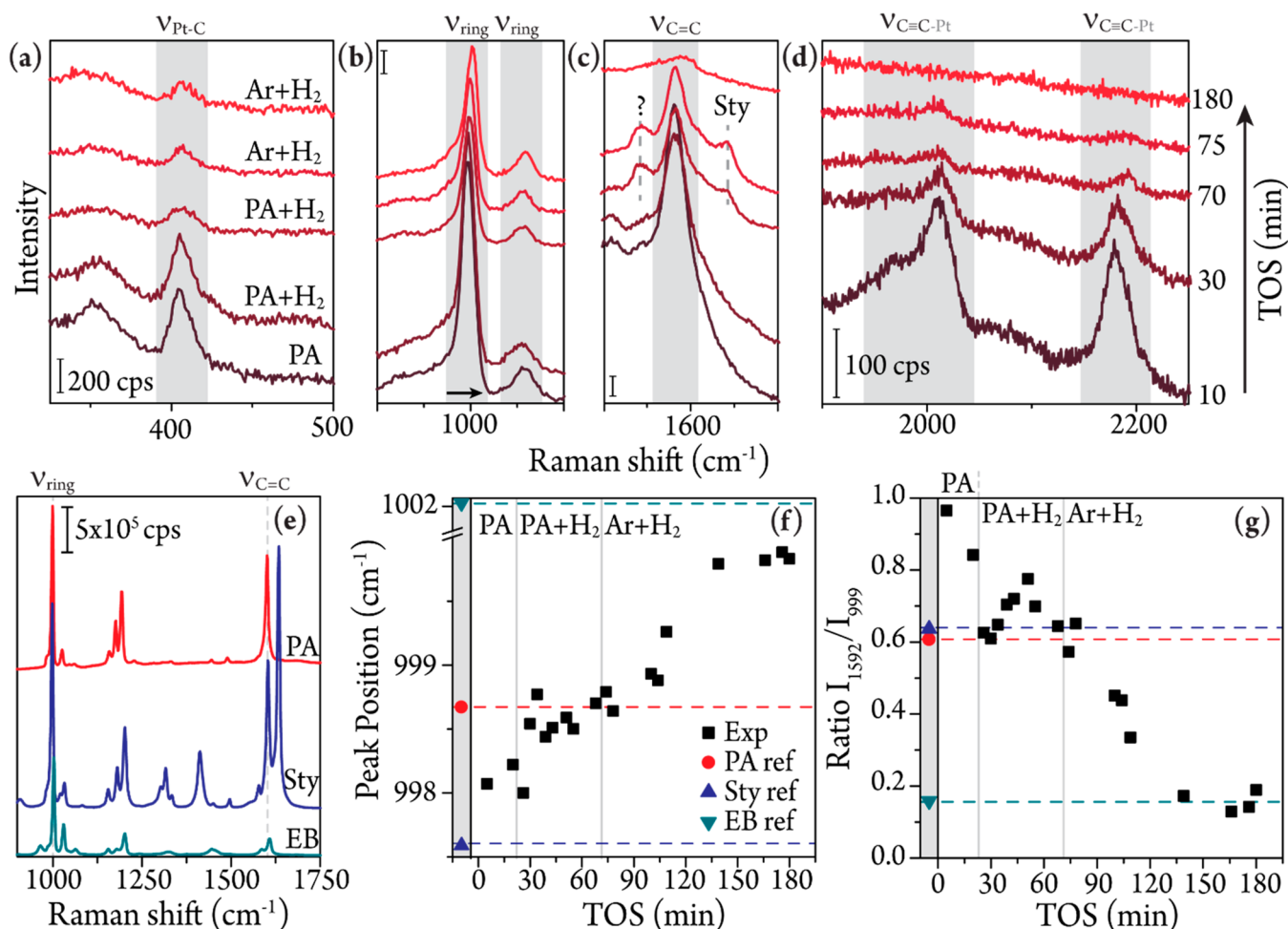


Figure 3. (a–d) Hydrogenation of phenylacetylene over Pt/Au@SiO₂ catalyst/SHINs at 60 °C with time-on-stream (TOS). (a) Pt–C stretching region. Upon hydrogenation of phenylacetylene, the peak attributed to Pt–C stretching decreases slightly and then remains constant. (b) Spectral region showing ring vibrations. Upon hydrogenation a shift to higher frequencies is observed, indicating hydrogenation occurs. (c) C=C stretching region indicating C=C stretching from the phenyl group. Two additional peaks attributed to C=C stretching can be observed during hydrogenation; at higher frequencies due to C=C stretching of the ethylene group present in styrene, and at lower frequencies due to π -bound ethylene from styrene. (d) C≡C stretching region which shows complete disappearance of alkyne vibrations over the course of hydrogenation. (e) Reference spectra of phenylacetylene (PA, red), styrene (Sty, blue) and ethylbenzene (EB, green) in which ring breathing vibrations and the aromatic C=C stretching vibrations used for further analysis are indicated. (f) Tracking the peak position of the ring breathing vibration with TOS shows a clear shift toward the reference position of ethylbenzene (green dashed line). Gas feed composition is indicated at the top of the graph. (g) Ratio I_{1592}/I_{999} with TOS compared to reference values of phenylacetylene (red), styrene (blue), and ethylbenzene (green) indicates the formation of ethylbenzene.

for hydrogenation reactions. N₂ was flowed through a bubbler filled with phenylacetylene to obtain a saturated gas feed that was led over the catalyst placed in a Linkam Cell (See Figure S6 for the experimental setup). Figure 2d and e report on the Pt–C vibration observed at 405 cm⁻¹ as we regard it as a measure of the amount of phenylacetylene present on the Pt surface.^{36,37} Therefore, phenylacetylene hydrogenation will be studied at the temperature for which we find an optimum in adsorption. Although no internal standard is present, by examining and focusing on the same spot on the sample, peak intensities can be reliably compared to one another within one experiment. Fitting and integration of the peak originating from the Pt–C stretching vibration in Figure 2d shows that an optimum in peak height exists at 60 °C, as displayed in Figure 2e. Figure 2f shows the C≡C stretching region where for pure phenylacetylene we expect to see one sharp peak at 2105 cm⁻¹. However, we observe that upon adsorption on the Pt catalyst, it splits into two distinct peaks at 2190 and 2010 cm⁻¹ with the

latter possessing a shoulder around 1960 cm⁻¹. This indicates that adsorption only takes place through the acetylene group and not through the aromatic ring. These observations are in line with adsorption reported of phenylacetylene and other alkynes on different transition metals^{38–46} and are attributed to chemisorption; σ -bonding and $\sigma\pi$ -complexation of a weak nature with different sites on the Pt surface, based on the relatively small shift in the C≡C vibrations compared to free (phenyl)acetylene.^{43,45} Furthermore, surface science experiments have shown that the aromatic ring in substituted benzenes orients itself away from the surface and is therefore not involved in adsorption.^{43,46}

Phenylacetylene adsorption SHINERS experiments were also performed on Au NPs and Au@SiO₂ SHINs for comparison (Figure S7). A discussion of these spectra shown in Figure S7 can be found in the Supporting Information.

Phenylacetylene Hydrogenation. Figure 3 shows SHINER spectra from hydrogenation of phenylacetylene to

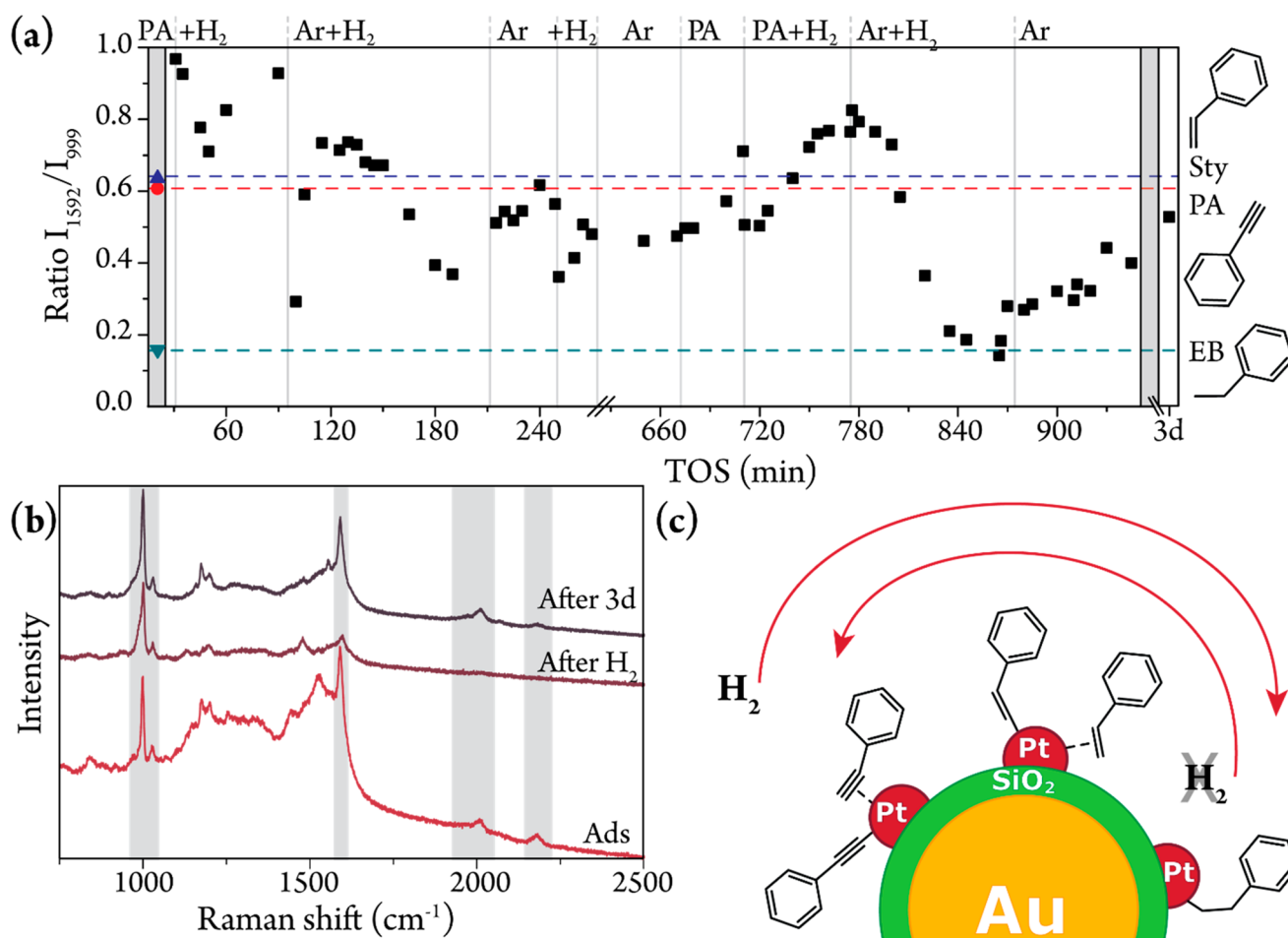


Figure 4. (a) Monitoring the ratio I_{1592}/I_{999} with TOS in hydrogen-rich and hydrogen-poor gas feeds compared to the reference compounds phenylacetylene (PA, red), styrene (Sty, blue), and ethylbenzene (EB, green). The gas feed composition is indicated above the figure. In hydrogen-rich environments, a clear shift toward the reference value for ethylbenzene is observed, whereas the reverse shift occurs in the absence of hydrogen. (b) Spectra of phenylacetylene adsorption at the start of the experiment (Ads), after hydrogenation (TOS = 860 min), and after 3 days under pure Ar flow. (c) Scheme of phenylacetylene hydrogenation and dehydrogenation including all observed species. Note that upon changing the gas feed composition, and especially when introducing H₂, the ratio I_{1592}/I_{999} tends to drop. We attribute this to the high heat capacity of H₂ and the fluctuations in conditions this brings about before equilibration.

styrene and ultimately ethylbenzene at 60 °C with time-on-stream (TOS) under different conditions. First, N₂ saturated with phenylacetylene and Ar were led over the catalyst in a 1:1 ratio. As expected based on the adsorption experiments, we observe a peak attributed to Pt–C stretching at 405 cm⁻¹ in Figure 3a and three peaks in the C≡C region in Figure 3d. Furthermore, Figure 3b shows peaks originating from ring breathing vibrations of a monosubstituted benzene around 1000 and 1030 cm⁻¹, and in Figure 3c we observe a peak from aromatic C=C stretching at 1592 cm⁻¹. All of the peaks originating from aromatic ring vibrations appear unperturbed compared to the reference spectra in Figure 3f and the literature.³⁶ This indicates a lack of interaction between the phenyl ring and the Pt surface, and further supports adsorption solely through C≡C as previously discussed.

After 30 min, the Ar flow was switched to H₂ and immediately a decrease in C≡C stretching was observed, whereas a similar decrease was not observed in the other spectral regions. Only after ~60 min of simultaneous phenylacetylene adsorption and hydrogenation did we observe a decrease, alongside an even further decrease in the C≡C stretching peak intensities. Additionally, we observe the

appearance of two new peaks in the C=C region. Comparison to reference spectra shown in Figures 3e and S8 shows that the peak at ~1634 cm⁻¹ can be attributed to C=C stretching originating from the ethylene group in styrene, whereas the peak at ~1571 cm⁻¹ does not match any of the peaks in the reference spectra. However, Mrozek and Weaver have reported similar peaks for π -bound ethylene on Pt electrodes studied with SERS.³¹ Monitoring the presence of this band over the experiment as displayed in Figure S9 shows that it is only present alongside the 1634 cm⁻¹ peak from the ethylene group in styrene, thereby indicating that the peak at 1571 cm⁻¹ can be attributed to π -bound styrene on Pt. Subsequently, we turned off the phenylacetylene flow and continued the experiments with a gas feed composed of H₂ and Ar to ensure full hydrogenation of all adsorbed species. A further decrease in peaks originating from C≡C stretching vibrations is observed, alongside the disappearance of those corresponding to ethylene. The aromatic C=C stretching vibrations at 1592 cm⁻¹ remain present, although with very little intensity. Furthermore, both Pt–C vibrations and ring breathing vibrations remain present, but whereas the Pt–C stretching vibrations remains at 405 cm⁻¹, the ring breathing vibrations

exhibit a shift to higher frequencies. Based on the reference spectra in Figure 3e and S8, this shift indicates a decrease in conjugation. Tracking the position of the ring breathing vibration at 999 cm^{-1} by deconvolution of the spectra (See Figure S10)⁴⁷ and subsequent comparison with reference spectra as displayed in Figure 3f show that the peak position moves in the direction of the reference value for ethylbenzene under hydrogenation conditions. Further inspection of the reference spectra also shows that the ratio of the peak intensity of the aromatic C=C stretching vibration (I_{1592}) to the peak intensity of the ring breathing vibration at 999 cm^{-1} (I_{999}) differs among phenylacetylene, styrene, and ethylbenzene. Upon fitting of the experimental spectra,⁴⁷ the ratio I_{1592}/I_{999} was compared to those of the reference compounds and again a trend was observed that indicates the formation of ethylbenzene, as illustrated in Figure 3g. Interestingly, switching back to pure Ar flow leads to an upward shift of the ratio I_{1592}/I_{999} , back to the values of styrene and phenylacetylene (Figure S11).

To determine whether we are witnessing dehydrogenation or simply readsorption of residual phenylacetylene, cycles of hydrogenation–dehydrogenation experiments were performed, as reported in Figure 4. Figure 4a shows the ratio I_{1592}/I_{999} with TOS under different gas feed compositions: pure phenylacetylene in N_2 (PA), PA and H_2 , Ar and H_2 , and pure Ar as indicated above the graph. Upon introduction of hydrogen, the ratio I_{1592}/I_{999} decreases to indicate ethylbenzene formation. However, as soon as H_2 is removed from the gas feed and we flow pure Ar, the ratio starts to increase again, away from the ethylbenzene reference value. This trend can be observed throughout the entire experiment and is also reflected in the ring breathing peak position (Figure S12). After leaving the cell under continuous Ar flow for 3 days, the ratio I_{1592}/I_{999} reaches the reference values for styrene and phenylacetylene. Over the course of this experiment, the intensity of the Pt–C stretching vibration was monitored for changes in the amount of adsorbates, presented in Figure S13. Inspection of this intensity shows that after equilibration the Pt–C peak height remains at constant values, excluding the possibility for readsorption of residual phenylacetylene which would be reflected in an increase in peak height. Figure 4b shows spectra at the start of the experiment, when phenylacetylene was adsorbed, at the end of the hydrogenation experiments at TOS = 860 min and after keeping the sample under constant Ar flow for 3 days. Ring breathing, C=C and C≡C stretching vibrations indicated in the figure show, first, that the aromatic ring stays intact, second, that no peaks ascribed to styrene are observed (see also Figure S9), and, third, C≡C stretching vibrations disappear upon hydrogenation but return in the absence of hydrogen, showing that the ethylbenzene that remains on the catalyst surface after hydrogenation is dehydrogenated back to phenylacetylene.

Based on the spectral observations we made, we propose that in our experiments phenylacetylene hydrogenation occurs as follows: Phenylacetylene is adsorbed on the Pt surface, giving rise to a Pt–C stretch at 405 cm^{-1} and a splitting of the C≡C band into three bands at 1960, 2010, and 2190 cm^{-1} due to chemisorption (σ -bonding and weak $\sigma\pi$ -complexation) and physisorption. The absence of free $\text{C}_2\equiv\text{C}_1\text{--H}$ stretching vibrations and lack of distortion in ring vibrations confirm adsorption takes place solely through the triple bond. Upon hydrogenation, we observe styrene both in a linear, σ -bound fashion at 1621 cm^{-1} and π -bound through its ethylene group

at 1571 cm^{-1} , depending on whether a hydrogen is inserted first in C_1 or C_2 , respectively. Upon complete disappearance of both the C≡C stretching vibrations and vibrations ascribed to styrene, ethylbenzene is observed. The presence of the Pt–C stretching peak allows us to assign this peak to linear σ -bonding of the terminal alkyl C_1 atom to the Pt surface, resulting in the scheme presented in Figure 4c. This is in line with the observation of a lack of π -bonding styrene species upon dehydrogenation, as logically linearly bound styrene should be formed. Furthermore, after dehydrogenation the peak ratios of the different C≡C stretching vibrations differ from those observed during adsorption. The peaks at 2010 and 1960 cm^{-1} we attribute therefore to chemisorbed species: The most intense peak at 2010 cm^{-1} we assign to $\text{C}_2\equiv\text{C}_1\text{--Pt}$ σ -bonding, and its shoulder at 1960 cm^{-1} to the weak $\sigma\pi$ -complexation of the acetylene group with Pt, as π -interaction of conjugated bonds parallel to the surface is often reported to result in a redshift in wavenumber.^{31,48,49} Finally, we propose the peak at 2190 cm^{-1} arises from adsorption at a less stable configuration or site,⁴⁰ based on the low intensity of the peak observed upon dehydrogenation.

CONCLUSIONS

Herein, we showed that *in situ* SHINERS allows for identification of catalyst-adsorbate interactions on the nano-scale, using the sequential gas-phase hydrogenation of phenylacetylene to styrene and ultimately ethylbenzene over Pt/Au@SiO₂ catalyst/SHINs as a model reaction. This work demonstrates that *in situ* SHINERS can be used to monitor hydrogenation reactions while they occur on the surface of the catalyst, proving its value as a generally applicable characterization method in the field of heterogeneous catalysis. By employing a setup and method that allows the reverse reaction to be studied, we were able to give more detailed Raman peak assignment and gain understanding of the reaction mechanism. Furthermore, we show that spectator or unreacting side groups can be used as an internal reference, opening up the possibility to study reactions of chemical groups with an inherent low Raman cross-section and reactions of complicated molecules over heterogeneous catalysts using *in situ* SHINERS.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acscatal.9b03010.

UV–vis spectroscopy of Au seeds, Au NPs, and Au@SiO₂ SHIN dispersions; TEM analysis data of Au seeds, Au NPs, Au@SiO₂ SHINs, and additional images of Pt/Au@SiO₂ catalyst/SHINs after reaction; pinhole and SERS-activity test for Au@SiO₂ SHINs; experimental setup for *in situ* SHINERS experiments; additional SHINER spectra; full reference spectra of phenylacetylene, styrene, and ethylbenzene; monitoring presence of peaks related to styrene; fitting of spectra; monitoring Pt–C stretching peak intensity; tracking ring breathing peak positions with TOS (PDF)

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Notes

The authors declare no competing financial interest.

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