

Reactivity of (Vicinal) Carbonyl Compounds with Urea

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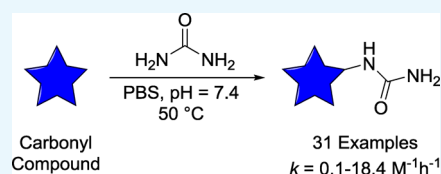
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Supporting Information

ABSTRACT: Urea removal from dialysate is the major obstacle in realization of a miniature dialysis device, based on continuous dialysate regeneration in a closed loop, used for the treatment of patients suffering from end-stage kidney disease. For the development of a polymeric urea sorbent, capable of removing urea from dialysate with high binding capacities and fast reaction kinetics, a systematic kinetic study was performed on the reactivity of urea with a library of low-molecular-weight carbonyl compounds in phosphate-buffered saline (pH 7.4) at 323 K. It was found that dialdehydes do not react with urea under these conditions but need to be activated under acidic conditions and require aldehyde groups in close proximity to each other to allow a reaction with urea. Among the 31 (hydrated) carbonyl compounds tested, triformylmethane, ninhydrin, and phenylglyoxaldehyde were the most reactive ones with urea. This is attributed to the low dehydration energies of these compounds, as calculated by Gibbs free energy differences between the hydrated and dehydrated carbonyl compounds, which are favorable for the reaction with urea. Therefore, future urea sorbents should contain such functional groups at the highest possible density.



INTRODUCTION

In the human body, urea is a metabolite of nitrogen-containing nutrients such as amino acids. These nutrients are metabolized in the liver into the toxic base ammonia, which is rapidly converted into urea by enzymatic processes. Urea is then transported through the bloodstream to the kidneys, where it is excreted into the urine.¹ The production of urea in the human body is 230–400 mmol/day.^{1,2} The majority of patients suffering from end-stage kidney disease undergo single pass hemodialysis in a center of three to four times per week to remove urea and other soluble waste products such as potassium, phosphate, and organic compounds from their body. However, the frequent and long visits to a dialysis center compromise their participation in economic and social life. Home hemodialysis therefore offers more flexibility and autonomy for dialysis patients. However, a large complex dialysis machine is required at home in combination with a large supply of dialysis fluids (≥ 30 L/treatment) or a bulky immobile water purification system, which discourages many patients from selecting home hemodialysis. A miniature dialysis device based on continuous regeneration and reuse of dialysate in a closed loop system, free from a fixed water supply, that can easily be used at home and during travel would greatly enhance patient's freedom and autonomy.^{3,4} It would also facilitate more frequent and longer hemodialysis enhancing blood purification, which is expected to improve clinical outcomes and quality of life.⁵

In current miniaturized dialysis devices that are under development, ion exchangers are used to remove excess ions such as phosphate and potassium, while most organic waste solutes such as creatinine and uric acid can effectively be removed by activated carbon.⁶ However, efficient urea removal is still a major obstacle in the development of a lightweight dialysis device.⁷ In principle, three strategies to remove urea from blood dialysate are available: enzymatic conversion, electrochemical degradation, and adsorption by a sorbent.⁸ Enzymatic conversion by immobilized urease, which has been used in the first commercial portable dialysis device that was marketed from 1973 to 1993 (REDY), results in formation of ammonium (NH_4^+) and carbon dioxide (CO_2).⁹ The ammonium is removed by a large quantity of a zirconium phosphate ion exchanger, which also captures calcium and magnesium ions that need to be replenished from a reservoir, which limits miniaturization of the device.^{3,10} Electrochemical degradation of urea into nitrogen (N_2), hydrogen (H_2), and carbon dioxide (CO_2) does allow miniaturization of hemodialysis. However, this urea removal method also produces harmful side products such as active chlorine species and reaction products thereof such as chloroamines.⁸ Therefore, ideally, a sorbent is applied that efficiently removes urea from dialysate by physi- or chemisorption. These sorbent

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materials are required to be stable that should not leach harmful compounds into the dialysate and should adsorb high quantities of urea from aqueous solution at a sufficient rate. Physisorption in general is a very fast, reversible, and concentration-dependent process. However, all urea sorbents based on physisorption known so far, such as activated carbon and zeolites, have relatively low affinity for urea, and desorption will occur as the urea concentration in the dialysate decreases in time.^{11,12} On the other hand, several polymeric chemisorbents that can bind urea covalently through carbonyl groups have been reported.^{13–18} Despite the weak nucleophilicity of urea nitrogen, it still can react with electrophilic carbonyl groups present in these sorbents such as glyoxaldehydes,^{13,15} ninhydrin groups,¹⁴ α -keto-esters,¹⁶ and oxidized alcohols.^{17,18} Covalent urea sorbents with binding capacities up to 2.0 mmol/g have been reported. However, these materials suffer from slow urea binding kinetics at 37 °C because only a part (1.2 mmol/g at most) of the urea-reactive groups has reacted after 8 h, still requiring more than 330 g of this material to remove the daily urea production (400 mmol).¹⁹

In the search for a more efficient polymeric urea sorbent suitable for incorporation in a miniature dialysis device, with high binding capacity and fast reaction kinetics, we performed a kinetic study on the covalent binding of urea with a library of low-molecular-weight carbonyl compounds. In the literature, uncatalyzed covalent binding of urea (derivatives) in water has been reported with compounds such as ninhydrin,^{20–22} aromatic glyoxaldehydes,²³ and other aldehydes.²⁴ Based on these reports, we systematically analyzed the kinetics of the reaction of several aromatic carbonyl compounds and non-aromatic analogues thereof with urea in aqueous solution representative for dialysate and determined the second-order rate constant. It is expected that the kinetics of the reaction of carbonyl compounds with urea decreases upon incorporation of the carbonyl groups in a polymeric sorbent due to substituent effects²² and decreased accessibility of polymeric groups as compared with molecules in solution. Still, this fundamental knowledge can be used for the future design of urea sorbents with fast kinetics and high binding capacities.

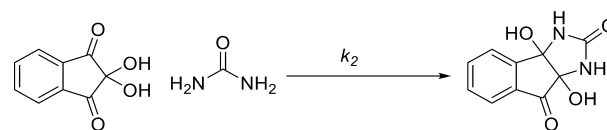
RESULTS AND DISCUSSION

A carbonyl compound can react with urea in a 1:1 molar ratio. Therefore, first-order dependence in both urea and the carbonyl compound is expected according to eq 1. Indeed, in our previous work, we analyzed the kinetics of urea with ninhydrin substituted with electron-donating groups (EDGs) and electron-withdrawing groups (EWGs) and concluded that the reaction of these ninhydrin derivatives with urea is first order in both urea and ninhydrin and thus second order overall.²² Also, the reaction of phenylglyoxaldehyde (PGA) with nucleophilic groups in enzymes such as NH₂ groups of arginine residues has been shown to be first order in PGA.^{25,26}

$$\frac{d[\text{urea}]}{dt} = -k_2[\text{urea}][\text{carbonyl compound}] \quad (1)$$

Quantitative ¹³C NMR was previously used as a robust method to determine the k_2 value of the reaction of urea with ninhydrin (Scheme 1).²² Using this method, ¹³C-enriched urea and ninhydrin are reacted in stoichiometric amounts, and therefore, their relative concentrations remain equal throughout the reaction (eq 2). Simplification of eq 1 using eq 2 results in eq 3

Scheme 1. Reaction of Ninhydrin, a Vicinal Tricarbonyl Hydrate, with Urea²¹



in which the second-order k_2 value is determined by only measuring urea concentrations in time.

$$[\text{urea}] = [\text{carbonyl compound}] \quad (2)$$

$$\frac{d[\text{urea}]}{dt} = -k_2[\text{urea}]^2 \quad (3)$$

The influence of temperature on the rate of the reaction of ninhydrin with urea was determined in phosphate-buffered aqueous saline (PBS) at temperatures between 20 °C (293 K) and 70 °C (343 K). The second-order k_2 values were determined from the plot of the inverse urea concentration in time (Section 2.1 of the Supporting Information), and the $\ln(k_2)$ values were plotted against the inverse of the absolute temperature in the Arrhenius plot (Figure 1). Linear regression

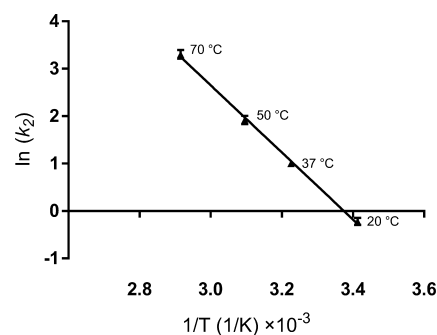


Figure 1. Arrhenius plot of the reaction of ninhydrin with urea in PBS. Linear regression gives $\ln(k_2) = -7073 (1/K) + 23.87$. Arrhenius equation: $\ln(k_2) = (-E_a/RT) + \ln(A)$; thus, $-E_a/R = -7073$ K and $\ln(A) = 23.87$.

and the use of Arrhenius equation resulted in a preexponential constant (A) of $(23.3 \pm 0.6) \times 10^9$ and an activation energy (E_a) of 14.0 ± 0.4 kcal/mol. To have reasonably fast kinetics, the reaction of the carbonyl compounds with urea was carried out at 323 K unless stated otherwise.

First, hydrates of a variety of similar vicinal tri- and tetraketones were reacted with urea in PBS at 323 K. The structures of the tested carbonyl compounds and the corresponding k_2 values are listed in Table 1. 2,2-Dihydroxy-4,4,5,5-tetramethylcyclopentane-1,3-dione (entry 2) is an aliphatic analogue of ninhydrin and was tested to investigate the influence of the aromatic ring on the reactivity of the vicinal triketone with urea. 2,2,3,3-Tetrahydroxy-2,3-dihydro-naphthalene-1,4-dione (entry 3) is an analogue of ninhydrin that consists of four vicinal ketones of which two form hydrates. To increase the number of urea-reactive groups per gram of material, also, 2,2,6,6-tetrahydroxy-*s*-indacene-1,3,5,7-(2*H*,6*H*)-tetraone (entry 4), which contains two triketone groups, was tested, together with its nonaromatic analogue 2,2,5,5-tetrahydroxy-3a,6a-dimethyl-6-methylenetetrahydropentalene-1,3,4(2*H*)-trione (entry 5).

Only the aromatic derivatives of the tested vicinal ketone hydrates in Table 1 (entry 1, 3, and 4) reacted with urea. The

Table 1. Structural Formulas of Tested Vicinal Ketone Hydrates and the Determined k_2 Values of Their Reaction with Urea in PBS at 323 K

Entry	Structure	k_2 ($M^{-1}h^{-1}$)
1		6.8 ± 0.6
2		<0.1
3		0.55 ± 0.07^a
4		0.32 ± 0.09^b
5		<0.1

^aThe compound is probably converted into ninhydrin during the reaction with urea.²⁷ ^bTested in 1:1 (v/v) DMSO/PBS at 343 K due to its low aqueous solubility.

urea concentration in the reaction mixture containing oxolin (entry 3) did decrease in time; however, the shift of the urea carbon peak in the ^{13}C NMR of the product corresponds to the shift of the same carbon in the ninhydrin–urea adduct (Section 1.1 of the Supporting Information). This suggests that the decrease in the urea concentration is not because of the reaction of oxolin with urea but rather due to the conversion of oxolin into ninhydrin²⁷ that subsequently reacts with urea forming the ninhydrin–urea product. In this case, eqs 1–3 are invalid, and the rate observed cannot be used to calculate the k value of oxolin with urea. Interestingly, the k_2 value of the reaction of entry 4 with urea substantially decreased by a factor of ~ 20 when compared to ninhydrin (entry 1), presumably due to the presence of two EWGs on the opposite site of the aromatic ring, that decreases the rate of dehydration of the hydrate, which is a required step to allow subsequent reaction with urea.²²

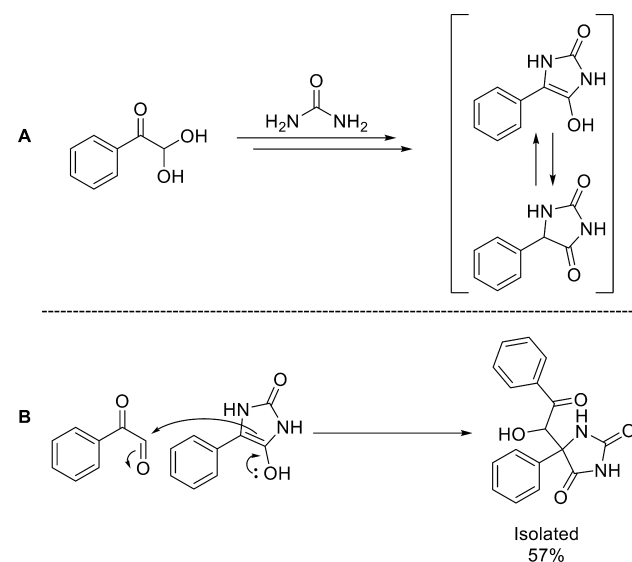
Because phenylglyoxaldehyde (PGA) has been reported to react with urea derivatives,²³ this compound (Table 2, entry 6) and an aliphatic analogue (entry 7) were tested for the reaction with urea. To increase the amount of urea-reactive groups per gram of material, also, 2,2'-(1,3-phenylene)bis(2-oxoacetaldehyde), a compound with an aromatic ring with two glyoxaldehyde functionalities, was tested (entry 8). For the reaction of phenylglyoxaldehyde with urea, the inverse of the urea concentration in time showed a nonlinear correlation (Section 2.2 and Figure S5, Supporting Information). Identification of the isolated reaction product after 24 h at 323 K in PBS by 1H and ^{13}C NMR spectroscopies showed that the initially formed urea–PGA product reacted with another equivalent of PGA. This means that the urea concentration did not remain equal to the PGA concentration during the reaction, and eqs 2 and 3 are therefore invalid, explaining the

Table 2. Structural Formulas of Tested Glyoxaldehydes and the Corresponding k_2 of Their Reaction with Urea in PBS at 323 K

Entry	Structure	k_2 ($M^{-1}h^{-1}$)
6		3.7 ± 0.1
7		1.2 ± 0.0
8		0.78 ± 0.07

nonlinear correlation observed in the plot of the inverse urea concentration in time. Scheme 2 shows (A) the reaction of

Scheme 2. Reaction and Proposed Mechanism^a



^a(A) Reaction of hydrated PGA with urea and (B) proposed mechanism of the reaction of the 1:1 PGA–urea adduct with a second PGA molecule; conditions: 0.5 mmol of urea and 0.5 mmol of PGA in 16.7 mL of PBS (30 mM urea and PGA) at 50 °C.

PGA with urea and (B) the proposed mechanism for the formation of the isolated product. Since the final product was isolated in a 57% yield with respect to PGA and is therefore the major formed product, the reaction of the 1:1 adduct with a second molecule of urea must be faster than the initial reaction of urea with PGA.

The k_2 values of the reaction of glyoxaldehydes (Table 2, entries 6–8) with urea were determined from the initial slope of the plot of the inverse urea concentration in time using second-order polynomial regression (Figure S5, Supporting Information). The k_2 values for the different glyoxaldehydes tested are listed in Table 2. To validate whether the slopes of the lines of Figure S5 can indeed be used for the calculation of the second-order reaction rate constants, PGA was reacted with an excess (3–50 equiv) urea to decrease the probability of the reaction of the 1:1 adduct with another PGA molecule. Thus, pseudo-first-order (PFO) conditions were obtained, and the rate constants (k_{PFO}) were measured. In a control

experiment, ninhydrin (30 mM) was reacted with 5 equiv of urea (150 mM) in PBS, and the k_{PFO} value was determined to be $1.0 \pm 0.1 \text{ h}^{-1}$ (Section 2.3 of the Supporting Information), which corresponds to a k_2 of $6.7 \text{ M}^{-1} \text{ h}^{-1}$ ($1.0 \text{ h}^{-1}/0.15 \text{ M} = 6.7 \text{ M}^{-1} \text{ h}^{-1}$), which is similar to the k_2 for ninhydrin (Table 1, entry 1). This validates that the k_{PFO} can indeed be used to determine the k_2 for the reaction of the carbonyl compound with urea.

The k_{PFO} values for the reaction of PGA with excess urea were plotted against the [urea] (Section 2.4 and Figure S8, Supporting Information), and this figure shows that k_2 (as defined of the derivative of $dk_{\text{PFO}}/d[\text{urea}]$) is not proportional with [urea] but decreases with increasing urea concentration. This is likely because of the dielectric constant of the medium changes, which influences the reactivity of PGA greatly.^{25,26,28} Determination of k_2 for PGA by extrapolating k_{PFO} to 30 mM using polynomial regression resulted in a k_2 of $3.7 \text{ M}^{-1} \text{ h}^{-1}$ (Figure S8 and Table S3, Supporting Information), which is similar to k_2 determined by the initial slope of the plot of the inverse urea concentration in time (Table 2, entry 6), thus validating that this method is an accurate method to determine the k_2 of glyoxaldehydes.

As for the vicinal ketone hydrates, the aromatic glyoxaldehyde (entry 6) showed a faster reaction with urea than the aliphatic (entry 7) and doubly substituted analogues (entry 8), the latter presumably also due to the decreased rate of dehydration of the hydrate in the presence of the meta-positioned EWG.

The enol form of triformylmethane (TFM) (entry 9 of Table 3) is a very electrophilic aldehyde, which has been reported to react with urea to yield an enamine (Scheme 3).²⁴ To further investigate the reaction of urea with other aldehydes than TFM, one monoaldehyde and several dialdehydes were selected, and their reaction kinetics with urea was tested (entries 10–13). In addition, electron-poor perfluorated ketones and aldehydes were also tested (entries 14 and 15), as well as a strained four-membered ring diketone (entry 16) and a dialdehyde that potentially forms an aromatic system upon reaction with urea (entry 17 and Section 3 of the Supporting Information). The structures of the selected aldehydes and ketones and the corresponding k_2 values are listed in Table 3.

In the reaction mixture of urea with TFM (entry 9), the urea concentration dropped very rapidly and remained constant after the first measurement at 30 min and 323 K. Therefore, besides reacting with urea, we assume that TFM was completely consumed by reacting with itself within this timeframe (as reported by others; see Section 4 of the Supporting Information).²⁹ To determine the k_2 value with sufficient accuracy, the reaction temperature was lowered to room temperature (293 K), and the initial slope of the plot of the inverse TFM concentration versus time (Figure S9, Supporting Information) showed a k_2 value of $18.4 \text{ M}^{-1} \text{ h}^{-1}$, which is 23 times faster than the reaction of urea with ninhydrin at the same temperature (Figure 1 and Table S2, Supporting Information) making TFM a very interesting compound to build in a urea sorbent.

In contrast to entry 9, the monoaldehyde (entry 10) and the 1,1-, 1,2-, and 1,3-dialdehydes (Table 3, entries 11–13) did not react with urea within 24 h at 323 K and pH 7.4. To allow the initially formed hemiaminal to dehydrate and form a more stable imine, the pH was decreased to pH 2 (at this pH, urea is not protonated, $\text{p}K_a$ of 0.1).³⁰ Under these conditions, the

Table 3. Structural Formulas of Tested Aldehydes and Ketones and the Determined k_2 Values of Their Reaction with Urea in PBS at 323 K

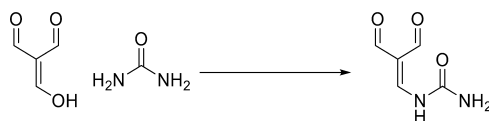
Entry	Structure	k_2 ($\text{M}^{-1}\text{h}^{-1}$) in PBS pH 7.4	k_2 ($\text{M}^{-1}\text{h}^{-1}$) in PBS pH 2.0
9		18.4 ± 5.2^a	
10		<0.1	<0.1
11 ^b		<0.1	18.1 ± 2.3
12 ^c		<0.1	10.8 ± 0.8
13		<0.1	1.0 ± 0.1
14		<0.1 ^d	
15		<0.1 ^d	
16		<0.1 ^d	
17		<0.1 ^d	

^aMeasured at 293 K instead of 323 K to decrease the reaction rate.

^b1,1,3,3-Tetramethoxypropane was stirred in PBS at pH 2 for 30 min at RT after which the pH was adjusted (to pH 2 or 7.4), and urea was added to the reaction mixture. ^c2,5-Dimethoxytetrahydrofuran was stirred in PBS at pH 2 for 30 min at RT after which the pH was adjusted (to pH 2 or 7.4), and urea was added to the reaction mixture.

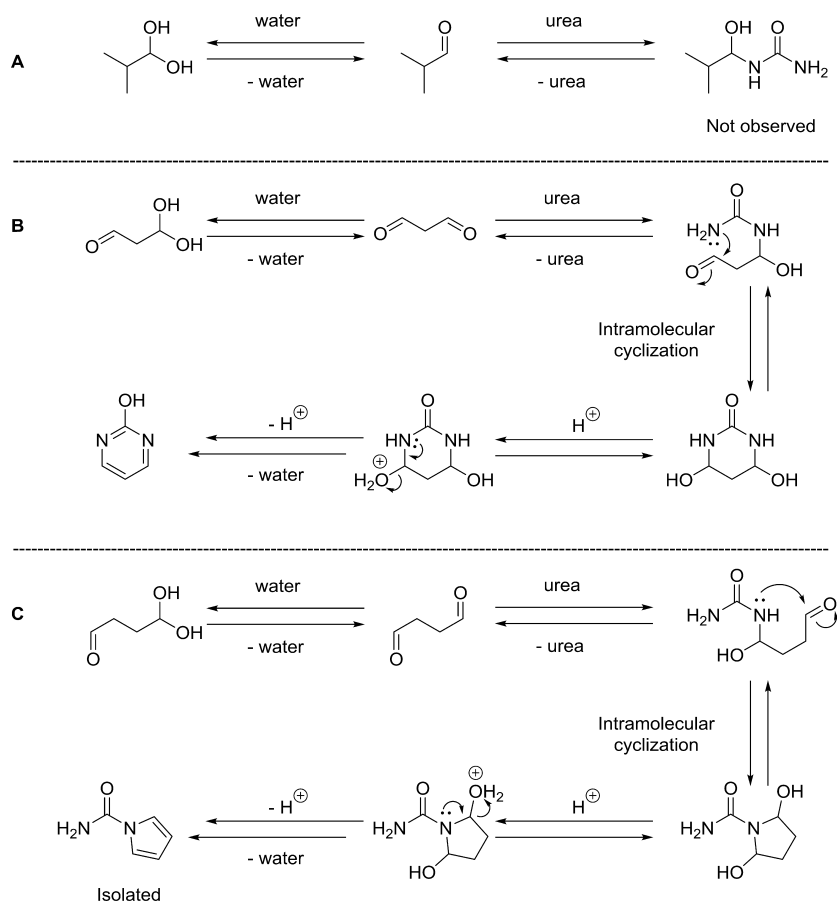
^dMeasured both in PBS and 1:1 PBS/DMSO as solvent at 323 K.

Scheme 3. Reaction of Triformylmethane with Urea²⁴



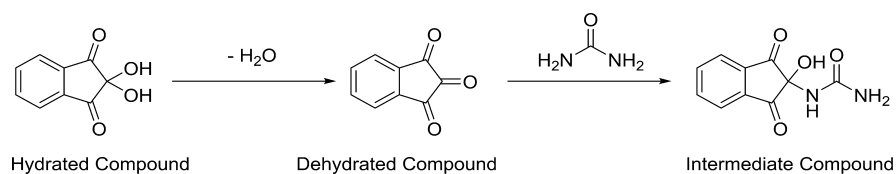
monoaldehyde still did not react with urea, whereas the dialdehydes did react with urea (Table 3, entry 11–13). The formed hemiaminal will be in equilibrium with the (hydrated) aldehyde. However, the presence of a second aldehyde group makes a second intramolecular nucleophilic attack of urea possible, thus driving the equilibrium toward the urea-bound aromatic product (Scheme 4). The results presented in Table 3 show that the kinetics of the reaction with urea decreased with increasing number of carbon atoms between the two aldehyde groups, presumably due to the decreasing mutual inductive electron-withdrawing effect of the aldehydes. Also, intramolecular ring closure likely becomes less favorable when the distance between the urea nitrogen and the aldehyde increases.

Based on these results, we conclude that to drive the equilibrium with water toward the urea adduct, at least two carbonyl groups in close proximity are required, allowing urea to react with both its nitrogens, forming a five- or six-

Scheme 4. Proposed Mechanisms^a

^a(A) Proposed mechanism of the equilibrium between a monoaldehyde and urea in H₂O; (B, C) proposed mechanism of the acid-catalyzed terminating reaction (B) between a 1,3-dialdehyde and urea in H₂O³¹ and (C) between a 1,4-dialdehyde and urea in H₂O.

Scheme 5. Dehydration of Ninhydrin and Subsequent Reaction with Urea

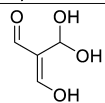
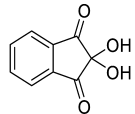
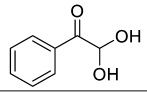
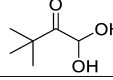
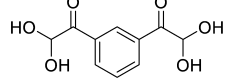
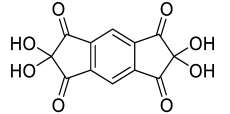
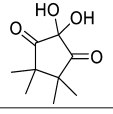
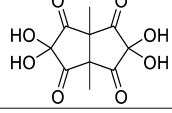
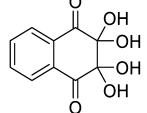
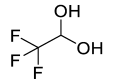
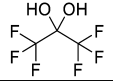


membered ring. TFM did not undergo an intramolecular ring closure toward a cyclic product at pH 7.4 because it can easily form a stable enamine at this pH (Scheme 3). Maltsev et al. have shown that under acidic conditions, TFM, like the dialdehydes tested (entries 11–13), undergoes an intramolecular ring closure into a six-membered ring.³²

The electron-deficient fluorinated aldehyde and ketone hydrates (entries 14 and 15) did not react with urea under the applied experimental conditions. Also, cyclobutanedione (entry 16) did not react with urea, despite that the reaction with urea could cause release of the ring strain because the hybridization of the carbonyl carbon would change from sp² to sp³ upon reaction with urea (Table 3, entry 16). The one with the possibility to form an aromatic system upon reaction with urea (Table 3, entry 17) as the possible driving forces did not react with urea, like many other carbonyl compounds that were expected to react with urea but did, in fact, not (listed in Section 1 of the Supporting Information).

To understand why certain carbonyl compounds and hydrates such as ninhydrin and PGA react with urea, whereas carbonyl compounds and hydrates such as hexafluoroacetone (entry 15) and 2,2-dihydroxy-4,4,5,5-tetramethylcyclopentane-1,3-dione (entry 2) do not, we calculated the changes in Gibbs free energy (ΔG) for these reactions using a computational model that we used in our previous study on the reaction of ninhydrin derivatives with urea.²² Most of the carbonyl compounds that we tested are predominately present in their hydrated forms. Consequently, the first step in the reaction mechanism is the dehydration of the hydrate and thus the formation of the carbonyl (as shown in Scheme 5 for ninhydrin).^{22,33} The ΔG of the dehydration of the carbonyl compound was calculated ($\Delta G_{\text{dehydration}}$) as well as the subsequent reaction of the dehydrated carbonyl compound with urea to form the intermediate adduct ($\Delta G_{\text{intermediate}}$). The structural formulas of the calculated hydrated, dehydrated, and intermediate compounds are listed in Section 5.2 of the Supporting Information.

Table 4. Structural Formulas, Rate Constants (323 K in PBS, pH 7.4), $\Delta G_{\text{dehydration}}$, and $\Delta G_{\text{intermediate}}$ of Selected Aldehydes and Ketones

Entry	Hydrated Compound	k_2 -value ($\text{M}^{-1}\text{h}^{-1}$)	$\Delta G_{\text{dehydration}}$ (kcal/mol)	$\Delta G_{\text{intermediate}}$ (kcal/mol)
A		18.4±5.2 ^a	-10.7	14.1
B ^b		6.8±0.6	-0.8	5.4
C		3.7±0.1	-6.1	7.6
D		1.2±0.0	-5.7	8.8
E		0.78±0.07	-3.5	9.2
F		0.32±0.09 ^c	0.4	4.3
G		<0.1	1.4	3.9
H		<0.1	0.0	4.3
I		<0.1 ^d	2.3	8.3
J		<0.1	1.3	2.9
K		<0.1	5.1	2.5

^aMeasured at 293 K in PBS pH 7.4. ^bData also reported in our earlier work.²² ^cMeasured at 343 K PBS/DMSO 1:1 (v/v). ^dOxolin is converted into ninhydrin.

By using a continuum solvent model as an approximation in the calculations, the solvation effects might not be accurately represented because the model does not include interactions such as hydrogen bonding. Additionally, the contribution of translational entropy to the Gibbs free energy in the gas phase may be inaccurately treated in solution. For these reasons, the obtained ΔG values should not be considered absolute values, but they can be compared relatively to each other as all reactions are dehydration or condensation. In fact, actual $\Delta G_{\text{dehydration}}$ should be positive in most cases because the compounds listed in Table 4 are present as the hydrate in water (with exception of TFM, entry A), and the free energy of the hydrate should be lower than the free energy of the corresponding carbonyl compound. NMR analysis showed that TFM is indeed mainly present in its nonhydrated form, which corresponds to the lowest calculated $\Delta G_{\text{dehydration}}$ for this compound. Interestingly, the compounds that react with urea

at 323 K (entries A–E) show a low ΔG value for dehydration (–10.7 to –3.5 kcal/mol), whereas the hydrates that show a relatively high value for dehydration (0.0 to 5.1 kcal/mol) react slowly (entry F) or not (entries G–K) with urea at 343 K. This shows that the Gibbs free energy for dehydration of carbonyl compounds gives a good indication for the reaction rate of the carbonyl compound with urea and can be used to screen promising carbonyl compounds.

TFM shows a rapid reaction with urea (Table 4, entry A) despite the high $\Delta G_{\text{intermediate}}$ because it is primarily dehydrated in solution. Ninhydrin (entry B) has a less favorable dehydration but a relatively low $\Delta G_{\text{intermediate}}$ and shows relatively fast kinetics for the reaction with urea as compared to the other hydrated carbonyl compounds. A clear trend that lower values of $\Delta G_{\text{dehydration}}$ and $\Delta G_{\text{intermediate}}$ coincide with a fast reaction rate with urea is observed for entries C–E. Carbonyl compounds F–K all have a relatively high

$\Delta G_{\text{dehydration}}$; therefore, the concentration of the urea-reactive species will be low, resulting in a rate constant of $<0.1 \text{ M}^{-1} \text{ h}^{-1}$.

CONCLUSIONS

In our aim to design an efficient polymeric urea sorbent, the kinetics of the reaction of urea with carbonyl compounds was systematically analyzed. TFM, ninhydrin, and PGA showed the fastest kinetics of their respective class of compounds among the ones tested. Importantly, to react with urea, at least two carbonyl groups are needed in close proximity to allow both nitrogens of urea to react, resulting in the formation of a stable five- or six-membered ring and effectively driving the equilibrium with water toward the urea adduct. This trend is shown for vicinal triketones, glyoxaldehydes, and dialdehydes. TFM, a trialdehyde that is predominately present in its dehydrated form, showed the highest reaction rate with urea and did not form a cyclic product but a stable enamine instead. Computational studies suggest that for carbonyl compounds with a relatively high $\Delta G_{\text{dehydration}}$, the dehydration of the hydrate is unfavorable and therefore the concentration of the urea-reactive species is low resulting in a very slow, often not measurable, reaction with urea.

To design urea sorbents with fast reaction kinetics at physiological conditions, suitable for incorporation in a miniature dialysis device, the focus should be on introducing a high density of functional groups such as TFM, ninhydrin, or PGA, for example, a polymer matrix.

MATERIAL AND METHODS

General Procedure. All chemicals were obtained from Sigma-Aldrich (Zwijndrecht, The Netherlands) and used as received unless stated otherwise. Phosphate-buffered saline (PBS, pH = 7.4; ion composition: Na^+ , 163.9 mM; Cl^- , 140.3 mM; HPO_4^{2-} , 8.7 mM; H_2PO_4^- , 1.8 mM) was obtained from B. Braun (Melsungen AG, Germany). 1,1,3,3-Tetramethoxypropane and 2,6-dimethylhepta-2,5-dien-4-one (phorone) were obtained from Combi Blocks (CA, USA). 4-Bromo-2-hydroxybenzaldehyde and 1,1'-(1,3-phenylene)bis(ethan-1-one) were obtained from AK Scientific (CA, USA). 2-Phenylmalonaldehyde was obtained from Apollo Scientific (Manchester, U.K.). 2,2,3,3-Tetrahydroxy-2,3-dihydronaphthalene-1,4-dione and [1,1'-biphenyl]-2,2'-dicarbaldehyde were obtained from Santa Cruz Biotechnology (CA, USA). 1,1-Dihydroxy-3,3-dimethylbutan-2-one was obtained from Enamine (NJ, USA). 3-Isothiocyanatoprop-1-ene was obtained from FluoroChem (Hadfield, U.K.). NMR spectra of the different compounds were recorded on an Agilent 400-MR DD2 equipped with a OneNMR probe or a Bruker 600 MHz with a BBI probe at RT. The quantitative 1D ^{13}C NMR experiments were carried out on the Agilent 400-MR DD2 using the standard s2pul.c pulse system with a spectral width of 24038.5 Hz and an acquisition time of 1.363 s. Residual solvent signals were used as internal standard (^1H : δ 2.50 ppm, ^{13}C (^1H): δ 39.52 ppm for $\text{DMSO}-d_6$; ^1H : δ 7.26 ppm, ^{13}C (^1H): δ 77.16 ppm for CDCl_3 ; ^1H : δ 5.32 ppm, ^{13}C (^1H): δ 53.84 ppm for CD_2Cl_2). Chemical shifts (δ) are given in ppm, and coupling constants (J) are given in hertz (Hz). Resonances are described as s (singlet), d (doublet), t (triplet), q (quartet), bs (broad singlet), and m (multiplet), or combinations thereof. UV absorption spectra and intensities were recorded with a BMG LABTECH SpectroStar Nano platereader using UV-Star Microplate 96 well obtained from Greiner Bio-One (Alphen

aan de Rijn, The Netherlands). Flash chromatography was performed over silica gel (particle size of 40–63 μm , VWR Chemicals, Leuven, Belgium) using the indicated eluent. Thin-layer chromatography (TLC) was performed using TLC plates from Merck Darmstadt, Germany, (SiO_2 , Kieselgel 60 F254 neutral, on aluminum with a fluorescence indicator), and the compounds were visualized by UV detection (254 nm).

General Procedure for the Kinetic Analysis of the Carbonyl Compounds with Urea. The carbonyl compound of interest (0.5 mmol, 1.0 equiv) was dissolved in 16.7 mL of PBS or a 1:1 (v/v) mixture of DMSO and PBS. ^{13}C -labeled urea (31.5 mg, 0.5 mmol, 1.0 equiv) was added to the solution, and the mixture was stirred magnetically at the indicated temperature for 2–96 h. Samples from the reaction were taken at different time points, and the urea concentrations in the different samples were determined by quantitative ^{13}C NMR (see below). The kinetic experiment was continued until more than 50% conversion was reached or when the urea concentration remained constant. The k_2 values for the carbonyl compounds were determined from the plots of $1/[\text{urea}]$ versus time.

Determination of Urea Concentration by ^{13}C NMR.²²

For the calibration curve, a stock solution of ^{13}C -labeled urea in PBS (50 mM) was prepared and diluted to 5, 10, 15, 20, 25, 30, and 35 mM. A solution of Me_2SO_2 in D_2O (1.50 mmol in 1000 μL , 1.34 M) was prepared and added to the analytes in a 10:1 ratio (1000 μL of analyte and 100 μL of internal standard). The mixture was shaken, and 600 μL was transferred into an NMR tube. For reaction mixtures in 1:1 PBS/DMSO, 200 μL of internal standard solution was added to 1000 μL of reaction mixture and transferred into an NMR tube. During the 1D ^{13}C NMR experiments (101 MHz), proton signals were decoupled. FIDs were Fourier transformed and automatic phase correction, and the third-order polynomial baseline correction and automatic integration were applied with MestReNova version 10.0.2-15465. Intra- and interassay coefficients of variation were within 10%. The ratios of integrals for urea at 162.08 ppm and internal standard at 42.15 ppm were compared with those of the calibration curve for calculating urea concentration in the samples.

Identification of Urea-Dialdehyde Adduct under

Acidic Conditions. 2,5-Dimethoxytetrahydrofuran (mixture of cis and trans) (0.5 mmol, 1.0 equiv) was dissolved in PBS, and the pH was adjusted with 1 M HCl to pH 2 and stirred for 30 min at RT. Urea (300.3 mg, 5 mmol, 10 equiv) was added, and the reaction mixture was heated to 50 $^\circ\text{C}$ and stirred for 5 h. The mixture was allowed to cool to RT and freeze-dried. The crude mixture was purified over silica (EtOAc/MeOH 19:1), and the ^1H NMR spectra of the isolated compound corresponded with 1H-pyrrole-1-carboxamide.³⁴

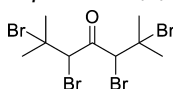
Determination of Carbonyl Compound Concentration by UV.

A carbonyl compound stock solution (30 mM) in PBS was diluted 30 times (4350 μL of PBS + 150 μL of 30 mM stock solution) to a 1 mM stock solution. A calibration curve was prepared by diluting the 1 mM stock solution to 0.005, 0.010, 0.020, 0.040, 0.060, 0.080, 0.100, and 0.120 mM with PBS. Samples of the solutions (200 μL) were transferred into a UV-Star 96-well plate in triplo. The concentration of the carbonyl compound in the reaction mixture was determined by taking a sample and directly diluting it 300 times at RT (30 times dilution; 150 μL of reaction mixture + 4350 μL of PBS) followed by 10 times dilution (150 μL of sample + 1350 μL of PBS) after which 200 μL was transferred into the UV-Star 96-

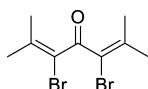
well plate in triplo. The TFM, ninhydrin, or PGA concentration in the samples was determined with the average UV absorption of 3 wells at 248, 232, and 250 nm, respectively.

Computational Studies. DFT calculations were performed using the Gaussian 09 software package, using the B3LYP (Becke, three-parameter, Lee–Yang–Parr) functional with 6-31 g(d,p) as the basis set on all atoms.³⁵ Structure optimizations were carried out with water as a solvent without any symmetry restraints in water. Frequency analyses were performed on all calculations. Input: #B3LYP/6-31G(d,p) opt = tight freq scf = tight int. = ultrafine pop = regular SCRF = (solvent = water).

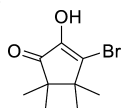
Syntheses and Spectroscopic Data. Synthesis of 4,4,3,3-Tetramethylcyclopentane-1,2,3-trione Hydrate.^{36,37}



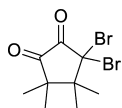
Bromine (10.3 mL, 145 mmol, 2.0 equiv) was dissolved in CH_2Cl_2 (50 mL). The solution was cooled to 0 °C, and 2,6-dimethylhepta-2,5-dien-4-one (10 g, 72.5 mmol, 1.0 equiv) was dissolved CH_2Cl_2 (50 mL) and added dropwise. Next, the reaction mixture was stirred for 1 h at 0 °C. The solvent was evaporated, and the crude mixture was recrystallized from EtOH, yielding 2,3,5,6-tetrabromo-2,6-dimethylheptan-4-one as a white solid (27.8 g, 60.9 mmol, 84%). ^1H NMR (CDCl_3 , 400 MHz, δ) main isomer: 5.15 (s, 2H), 2.02 (s, 12H). ^{13}C NMR (CDCl_3 , δ) main isomer: 193.88 (C_q), 62.83 (C_q), 59.62 (CH), 33.90 (CH_3), 29.29 (CH_3).



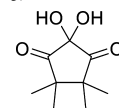
2,3,5,6-Tetrabromo-2,6-dimethylheptan-4-one (50 mmol, 1.0 equiv) was dissolved in cold pyridine (30 mL) and kept for 24 h at 0 °C without stirring. The reaction mixture was transferred into a separation funnel with CH_2Cl_2 , washed with 30% H_2SO_4 (twice), H_2O , and brine, and dried over Na_2SO_4 . After evaporation of the solvent, 3,5-dibromo-2,6-dimethylhepta-2,5-dien-4-one was isolated as a brown oil, which started crystallizing when dried under vacuum and at RT (5.35 g, 49 mmol, 98%). ^1H NMR (CDCl_3 , 400 MHz, δ): 2.04 (s, 6H), 1.99 (s, 6H). ^{13}C NMR (CDCl_3 , 101 MHz, δ): 190.25 (C_q), 144.98 (C_q), 115.64 (C_q), 25.99 (CH_3), 22.55 (CH_3).



3,5-Dibromo-2,6-dimethylhepta-2,5-dien-4-one (4.158 g, 16 mmol, 1.0 equiv) was dissolved in cold H_2SO_4 (2.56 mL, 48 mmol, 3.0 equiv) and stirred at 0 °C for 24 h. The mixture was poured over crushed ice during which 3-bromo-2-hydroxy-4,4,5,5-tetramethylcyclopent-2-en-1-one precipitated. The resulting suspension was filtered, and the obtained solid material was washed with water and dried under vacuum at RT, yielding 3-bromo-2-hydroxy-4,4,5,5-tetramethylcyclopent-2-en-1-one in a 76% yield (2.8 g, 12 mmol). ^1H NMR (CDCl_3 , 400 MHz, δ): 5.82 (bs, OH), 1.13 (s, 6H), 1.11 (s, 6H). ^{13}C NMR (CDCl_3 , 101 MHz, δ): 204.73 (C_q), 147.49 (C_q), 138.72 (C_q), 51.18 (C_q), 46.60 (C_q), 24.99 (CH_3), 22.64 (CH_3).

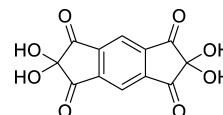


3-Bromo-2-hydroxy-4,4,5,5-tetramethylcyclopent-2-en-1-one (3.5 mmol, 1.0 equiv) was dissolved in cold acetic acid (4 mL) to which bromine (178 μL , 3.5 mmol, 1.0 equiv) in acetic acid (4 mL) was slowly added at 0 °C. The reaction mixture was allowed to heat to RT and stirred for 2 h. The mixture was poured over ice and rested for 10 min. The suspension was filtrated, and the obtained solid material was washed with water and dried under vacuum, yielding 3,3-dibromo-4,4,5,5-tetramethylcyclopentane-1,2-dione as a brown solid in a 38% yield (419 mg, 1.34 mmol). ^1H NMR (CDCl_3 , 400 MHz, δ): 1.39 (s, 6H), 1.34 (s, 6H). ^{13}C NMR (CDCl_3 , 101 MHz, δ): 214.13 (C_q), 204.74 (C_q), 73.32 (C_q), 50.68 (C_q), 46.92 (C_q), 25.45 (CH_3), 24.10 (CH_3).



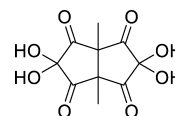
The procedure as described by Gill et al. and Shoppee^{36,37} did not yield 2,2-dihydroxy-4,4,5,5-tetramethylcyclopentane-1,3-dione in our hands. Therefore, we modified the procedure inspired by the work of Liang et al.³⁸ 3,3-Dibromo-4,4,5,5-tetramethylcyclopentane-1,2-dione (157 mg, 0.5 mmol, 1.0 equiv) was dissolved in DMSO (1 mL), and iodine (0.1 mmol, 0.2 equiv) and water (0.5 mmol, 1.0 equiv) were added. The reaction mixture was stirred at 50 °C for 24 h after which the resulting mixture was freeze-dried to remove DMSO. Purification of the crude mixture over silica (hexane/EtOAc 1:1) yielded 2,2-dihydroxy-4,4,5,5-tetramethylcyclopentane-1,3-dione as a yellow solid in a 90% yield. ^1H NMR (CDCl_3 , 400 MHz, δ): 6.02 (bs, 2H), 1.14 (s, 6H), 1.12 (s, 6H). These peak shifts differ from those reported by Gill et al. and Shoppee (δ : 1.20 (s, 6H) and 1.08 (s, 6H)). ^{13}C NMR (CDCl_3 , 101 MHz, δ): 214.21 (C_q), 137.78 (C_q), 51.10 (C_q), 24.97 (CH_3), 22.61 (CH_3).

Synthesis of *s*-Indacene-1,2,3,5,6,7-hexaone Hydrate.



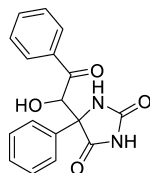
s-Indacene-1,3,5,7(2*H*,6*H*)-tetraone was synthesized according to the procedure of Krief et al.³⁹ (^1H NMR (CD_2Cl_2 , 400 MHz, δ): 8.48 (s, 2H), 3.42 (s, 4H)) and oxidized into *s*-indacene-1,2,3,5,6,7-hexaone hydrate according to the procedure reported by Marminon et al.⁴⁰ *s*-Indacene-1,3,5,7(2*H*,6*H*)-tetraone (171 mg, 0.8 mmol, 1.0 equiv) was dissolved in a 10:1 mixture of dioxane (3 mL) and H_2O (0.3 mL) in a microwave tube equipped with a magnetic stirrer. Selenium dioxide (275 mg, 2.5 mmol, 3.1 equiv) was added, and the tube was sealed. The mixture was shaken vigorously until selenium dioxide was dissolved, and the tube was placed in the microwave where it was heated for 5 min at 180 °C. The crude reaction mixture was impregnated on silica and purified over silica (8:1 $\text{CH}_2\text{Cl}_2/\text{MeOH}$), obtaining the product in a 30% yield (66 mg, 0.24 mmol). ^1H NMR (DMSO, 600 MHz, δ): 8.45 (s, 2H), 7.82 (bs, 4H). ^{13}C NMR (DMSO, 125 MHz, δ): 195.66 (C_q), 143.66 (C_q), 119.68 (CH), 88.31 (C_q).

Synthesis of 3*a*,6*a*-Dimethyldihydropentalene-1,2,3,4,5,6-hexaone.



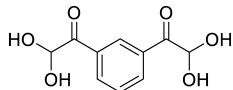
3a,6a-Dimethyltetrahydropentalene-2,5(1H,3H)-dione (499 mg, 3.0 mmol, 1.0 equiv) was dissolved in DMSO, and 48% HBr (4.1 mL, 36 mmol, 12 equiv), 57% HI (171 μ L, 1.5 mmol, 0.5 equiv), and I₂ (380 mg, 1.5 mmol, 0.5 equiv) were added. The mixture was stirred magnetically at 80 °C for 4 h. Afterward, the mixture was allowed to cool to RT and was extracted with CH₂Cl₂ five times. The organic layers were combined, dried over Na₂SO₄, and concentrated during which red crystals precipitated from the solution. After filtration, the crystals were collected in a 36% yield (279 mg, 1.1 mmol). ¹H NMR (DMSO, 400 MHz, δ): 10.68 (bs, 4H), 1.41 (s, 3H), 1.32 (s, 3H). ¹³C NMR (DMSO, 101 MHz, δ): 191.81 (C_q), 148.24 (C_q), 133.28 (C_q), 65.11 (C_q), 52.39 (C_q), 19.47 (CH₃), 17.08 (CH₃).

Synthesis of 5-(1-Hydroxy-2-oxo-2-phenylethyl)-5-phenylimidazolidine-2,4-dione.



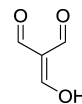
Phenylglyoxaldehyde (456 mg, 3.0 mmol, 1.0 equiv) was dissolved under continuous stirring in PBS (100 mL), and urea (180 mg, 3.0 mmol, 1.0 equiv) was added. The reaction mixture was heated to 50 °C and stirred for 24 h. Afterward, the mixture was allowed to cool to RT, and the solvent was removed by freeze-drying. The crude reaction mixture was impregnated on silica and purified over silica (hexanes/EtOAc = 5:1), obtaining 5-(1-hydroxy-2-oxo-2-phenylethyl)-5-phenylimidazolidine-2,4-dione as a mixture of diastereomers (266 mg, 0.86 mmol) as a white solid in a 57% yield with respect to phenylglyoxaldehyde. Major diastereomer: ¹H NMR (DMSO, 400 MHz, δ): 10.77 (bs, 1H), 8.23 (bs, 1H), 7.82 (d, *J* = 8.3 Hz, 2H), 7.59 (d, *J* = 8.2 Hz, 2H), 7.53 (m, 1H), 7.40 (t, *J* = 7.6 Hz, 2H), 7.25 (t, *J* = 7.6 Hz, 2H), 7.18 (m, 1H), 6.65 (d, *J* = 8.7 Hz, 1H), 5.71 (d, *J* = 8.5 Hz, 1H). ¹³C NMR (DMSO, 101 MHz, δ): 196.77 (C_q), 174.88 (C_q), 157.11 (C_q), 136.02 (C_q), 135.48 (C_q), 133.33 (CH), 128.69 (CH), 128.43 (CH), 128.09 (CH), 127.68 (CH), 125.78 (CH), 73.64 (CH), 70.60 (C_q).

Synthesis of 1,1'-(1,3-Phenylene)bis(2,2-dihydroxyethan-1-one).



1,1'-(1,3-Phenylene)bis(ethan-1-one) (162 mg, 1.0 mmol, 1.0 equiv) was dissolved in a 10:1 mixture of dioxane (3 mL) and H₂O (0.3 mL) in a microwave tube equipped with a magnetic stirrer. Selenium dioxide (344 mg, 3.1 mmol, 3.1 equiv) was added, and the tube was sealed. The mixture was shaken vigorously until selenium dioxide was completely dissolved, and the tube was placed in the microwave oven in which it was heated for 5 min at 180 °C. The crude reaction mixture was impregnated on silica and purified over silica (EtOAc), obtaining an orange oil in a 41% yield (93 mg, 0.41 mmol). ¹H NMR (DMSO, 600 MHz, δ): 8.71 (t, *J* = 2.0 Hz, 1H), 8.32 (m, 2H), 7.66 (t, *J* = 7.6 Hz, 1H), 6.90 (bs, 4OH), 5.68 (s, 2H). ¹³C NMR (DMSO, 125 MHz, δ): 188.67 (C_q), 133.92 (CH), 133.84 (C_q), 130.07 (CH), 128.64 (CH), 89.64 (CH).

Synthesis of Triformylmethane.



Triformylmethane was synthesized according to the procedure of Buděšínský et al.⁴¹ ¹H NMR (CDCl₃, 400 MHz, δ): 9.50 (s, 1H), 9.02 (s, 2H). ¹³C NMR (CDCl₃, 101 MHz, δ): 186.60 (CH), 118.42 (C_q). One ¹³C signal is not observed, possibly due to tautomerization.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acsomega.9b01177.

Structures of carbonyl compounds tested that are not listed in main manuscript; ¹³C NMR spectra of the reaction products of oxolin and ninhydrin with ¹³C-labeled urea; plots of the inverse urea concentration in time for the reaction of urea with carbonyl compounds; PFO experiments of PGA and ninhydrin with excess urea; scheme of the reaction of entry 17 with urea; overview of side reactions of TFM in water; details of the computational data; structural formulas, calculated *G* values and optimized coordinates of all hydrated, dehydrated, and intermediate compounds; ¹H and ¹³C NMR spectra of synthesized compounds (PDF)

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Notes

The authors declare no competing financial interest.

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